

# Bench Scale Performance of Partitioning Electron Donors for TCE DNAPL Bioremediation

by

Jeffery D. Roberts

A thesis  
presented to the University of Waterloo  
in fulfillment of the  
thesis requirement for the degree of  
Master of Science  
in  
Earth Sciences

Waterloo, Ontario, Canada, 2008

©Jeffery D. Roberts, 2008

## **AUTHOR'S DECLARATION**

I hereby declare that I am the sole author of this thesis. This is a true copy of the thesis, including any required final revisions, as accepted by my examiners. I understand that my thesis may be made electronically available to the public.

## Abstract

Prior to the implementation of an enhanced bioremediation pilot study for a trichloroethene (TCE) source area at an industrial site in the United Kingdom (the Site), laboratory microcosm and column studies were performed. The purpose of this column study was to determine if TCE removal rates could be increased with the addition of partitioning electron donors and bioaugmentation with KB-1<sup>®</sup> culture. Three 1-meter continuous flow columns were constructed using aquifer solids from the Site and artificial groundwater. A TCE dense non-aqueous phase liquid (DNAPL) zone was emplaced in each column. SRS<sup>™</sup>, a commercially available emulsified vegetable oil (EVO) product, and n-butyl acetate (nBA) were evaluated as partitioning electron donors, while the third column acted as an unamended control. Both nBA and SRS<sup>™</sup> were successfully used in previous microcosm studies with high concentrations of TCE (400 and 800 mg/L) to successfully promote the reductive dechlorination of TCE to ethene.

Dechlorination of TCE to *cis*-1,2-dichloroethene (*cis*-DCE) with trace amounts of vinyl chloride (VC) and ethene, as well as sulfate reduction, were observed in the SRS<sup>™</sup> column effluent while DNAPL was present. A dissolution enhancement factor of 2.1 was calculated. The TCE source zone was depleted after approximately 300 days of column operation. Following depletion of the TCE DNAPL, high concentration (~400 mg/L) of TCE amended artificial groundwater was pumped through the column to simulate high TCE concentrations in a plume down gradient from a source zone. Dechlorination of TCE via *cis*-DCE and VC to ethene was observed in the column effluent along with increases in *Dehalococcoides* (Dhc) counts. Sulfate concentrations increased during the plume phase while dechlorination to ethene still occurred indicating that complete dechlorination to ethene was possible in the presence of sulfate.

Dechlorination of TCE to *cis*-DCE was observed, but neither VC nor ethene was detected in the nBA Amended column. The nBA was observed to degrade in the column to butyl alcohol and acetate, neither of which partition as strongly as nBA, and were not retained in

the column. A continuous addition of nBA promoted the highest amount of *cis*-DCE production and sulfate reduction was also observed. Once the continuous addition was stopped, dechlorination and sulfate reduction halted indicating that electron donor retention in the column was not achieved. *Dehalococcoides* (Dhc) concentrations did not increase in the effluent of this column. A dissolution enhancement factor of 1.2 was calculated for the nBA column.

## Acknowledgements

I would like to thank my supervisor Dr. Jim Barker for all of his guidance and support and my co-supervisors Dr. Colin Mayfield and Dr. David Major for their valuable comments.

Thanks to Dr. David Major and Evan Cox at Geosyntec for inspiring me to pursue this degree and giving me the opportunity while continuing to work full time. Also, to Sandra Dworatzek and Phil Dennis at SiREM for all of your help and for allowing me to take the time needed away from the lab to complete this degree. A special thanks to Eric Hood at Geosyntec for your advice and thoughtful comments along the way.

To the SiREM lab staff - Rita, Ian, Kitson, Ryan, Jen, Jen, Jenny, Ximena, Diane and Liz. I really appreciate all of the support on this project; from packing the columns to measuring flow rates, making artificial groundwater, taking samples, analyzing samples, and compiling data - I was able to make it through this with your help. Special thanks to Rita, Kitson and Ian for checking the columns on weekends so that I didn't have to make the drive from Waterloo. Also to all for allowing me to take over the new fume hood for over two years!

Thanks to Shirley Chatten and Marianne VanderGriendt in the Organic Geochemistry laboratory for analytical support and for allowing me to borrow equipment. Thanks to Rendnil Manufacturing for constructing the effluent collection vial flow through devices.

Thank you to all on the Sabre team, especially the laboratory workgroup. Our monthly conference calls provided me with invaluable advice and ensured I was always moving forward in a sensible direction. Thanks to Amalia Kokkinaki for the modeling work and helpful discussions. I feel lucky to be part of such a wonderful group and look forward to working with all of you in the future.

To my parents – Thank you for your love and support, in all aspects of my life, and for always believing in me. Thanks to my brothers and sister-in-laws, Mike and Joanne, Chris and Laurie, and Jamie and Kate for always being there. I'd also like to thank Brian and Ethel Wood for your encouragement and motivation and especially my wife Pam for your ongoing love, support, and patience; I couldn't have done this without you.

## Dedication

*For Pam*

## Table of Contents

AUTHOR'S DECLARATION .....	ii
Abstract .....	iii
Acknowledgements .....	v
Dedication .....	vi
Table of Contents .....	vii
List of Figures .....	x
List of Tables .....	xii
Chapter 1 Introduction.....	1
1.1 Introduction .....	1
1.2 Research Objectives .....	3
Chapter 2 Background Information.....	4
2.1 Trichloroethene and DNAPLS .....	4
2.2 Bioremediation .....	6
2.2.1 Reductive Dechlorination.....	7
2.2.2 KB-1 <sup>®</sup> Culture .....	8
2.2.3 Electron Donors.....	9
2.2.4 Important Side Reactions .....	12
2.2.5 Bioremediation of DNAPLS .....	14
Chapter 3 Materials and Methods.....	17
3.1 Equipment Used .....	17
3.2 Analytical Procedures.....	18
3.2.1 Volatile Organic Compound and Dissolved Hydrocarbon Gas Analysis.....	18
3.2.2 Anions and Volatile Fatty Acids .....	20
3.2.3 pH and ORP.....	21
3.2.4 Molecular Methods.....	21
Chapter 4 Batch Tests.....	24
4.1 Summary of Batch Tests .....	24
4.1.1 Summary of Sabre Microcosm Study.....	24
4.1.2 Partitioning Experiment.....	26
4.1.3 nBA Toxicity Experiment .....	29
4.1.4 SRS <sup>™</sup> Loading Percentage Experiment.....	32

4.2 Conclusions from Batch Tests .....	33
Chapter 5 Column Study Experimental Design .....	34
5.1 Column Set-up and Construction .....	34
5.1.1 Soil Collection and Artificial Groundwater .....	34
5.1.2 Column Packing .....	34
5.1.3 DNAPL Emplacement .....	35
5.2 Electron Donor and Nutrient Addition and Bioaugmentation .....	37
5.2.1 nBA Column Amendments .....	37
5.2.2 SRS™ electron donor .....	40
5.2.3 Nutrient Addition, Bioaugmentation, and Buffering .....	40
Chapter 6 Column Study Results .....	42
6.1 Column Study Results .....	42
6.1.1 Control Column Results .....	43
6.1.2 nBA column results .....	45
6.1.3 SRS™ Column Results .....	51
Chapter 7 Column Study Discussion .....	58
7.1 TCE DNAPL Distributions .....	58
7.2 Effect of TCE Concentrations on Indigenous Microorganisms .....	60
7.3 Electron donor retention in columns .....	60
7.3.1 SRS™ retention in column .....	60
7.3.2 nBA Retention in Column and Mass Balance .....	63
7.4 Effects of Biostimulation .....	65
7.5 Effects of Bioaugmentation .....	67
7.5.1 nBA column .....	67
7.5.2 SRS™ Column .....	68
7.5.3 Importance of Sulfate Reduction and Dechlorination in the SRS™ Column .....	71
7.6 Mass balances .....	72
7.6.1 Chlorinated Ethene and Ethene Mass Balances .....	72
7.6.2 Chloride Mass Balances .....	73
7.7 Evidence of Enhanced Dissolution .....	76
Chapter 8 Conclusions and Future Work .....	80
8.1 Conclusions .....	80



8.2 Future Work .....	81
Appendix A Procedure For Making Artificial Ground Water.....	87
Appendix B Electron Donor Demand Calculations .....	89
Appendix C Bromide Tracer Test .....	94
Appendix D Mass Balance Calculations .....	97
Appendix E Raw Data Sheets.....	100

## List of Figures

Figure 2.1: Structures of chlorinated ethenes.....	4
Figure 2.2: DNAPL Transport Processes (U.S. EPA, 2000) .....	6
Figure 2.3: Pathway for the reductive dechlorination of chlorinated ethenes.....	7
Figure 3.1: Schematic of Column Study Setup.....	17
Figure: 3.2 Effluent collection vial .....	18
Figure 4.1: Chlorinated Ethenes, Ethene, and Sulfate Concentration trends in SRS™ amended high TCE plus nutrients plus KB-1® Microcosms .....	25
Figure 4.2: Chlorinated Ethenes, Ethene, and Sulfate Concentration trends in nBA amended low TCE plus nutrients plus KB-1® Microcosms.....	25
Figure 4.3: VOC and ethene concentration trends in nBA toxicity test microcosms .....	31
Figure 4.4: SRS Partitioning Experiment Reactor Bottles.....	32
Figure 4.5: TCE Concentration Profile of the 4.5 % SRS™ Amended Reactor.....	33
Figure 5.1: Sabre test cell aquifer solids .....	35
Figure 5.2: Schematic of DNAPL Addition and Drainage .....	36
Figure 6.1: VOC and ethene concentration trends in the control column.....	44
Figure 6.2: Anion concentration trends in the control column .....	44
Figure 6.3: pH and ORP trends in the control column.....	45
Figure 6.4: VOC and ethene concentration trends in the nBA Amended column .....	48
Figure 6.5: Anion concentration trends in the nBA Amended column.....	49
Figure 6.6: VFA concentration trends in the nBA Amended column.....	49
Figure 6.7: nBA and butanol concentration trends in the nBA Amended column .....	50
Figure 6.8: pH and ORP trends in the nBA amended column .....	50
Figure 6.9: VOC, Ethene and Methane concentration trends in the SRS amended column .....	54
Figure 6.10: TCE concentration trends in the influent artificial groundwater bag .....	55
Figure 6.11: Anion concentration trends in the SRS amended column .....	55
Figure 6.12: VFA concentration trends in the SRS amended column .....	56
Figure 6.13: pH and ORP trends in the SRS amended column .....	56
Figure 6.14: Dhc and vcrA concentration trends in the SRS amended column.....	57

Figure 7.1: Experimental and Simulation Results for the TCE Breakthrough curves in the Control and SRS <sup>TM</sup> Columns.....	59
Figure 7.2: Dyed TCE distribution in the Control Column.....	59
Figure 7.3: Effluent concentration trends of nBA, butanol, butyrate, and acetate .....	64
Figure 7.4: Dhc, vcrA, VC, and Ethene concentration trends in the SRS amended column.....	69
Figure 7.5: Cumulative Mass Recovery in the SRS and Control Columns.....	78
Figure 7.6: Cumulative Mass Recovery in the nBA and Control Columns .....	78
Figure 7.7: Normalized Enhanced Dissolution in the SRS and Control Columns .....	79
Figure 7.8: Normalized Enhanced Dissolution in the nBA and Control Columns.....	79
Figure C1: Bromide Breakthrough curves for Control, nBA, and SRS <sup>TM</sup> Columns .....	96

## List of Tables

Table 2.1: Physical and Chemical Properties of chlorinated ethenes and ethene .....	5
Table 3.1 – HP 7694 Parameters.....	19
Table 4.1: Summary of nBA Partitioning Experiment Results.....	28
Table 5.1: Summary of DNAPL Additions .....	37
Table 5.2: Summary of nBA Amendments.....	38
Table 5.3: Concentrations of nBA, acetate, and butyrate in the influent bag during the 2,000 mg/L nBA amendment phase .....	39
Table 6.1 Summary of Column Properties.....	42
Table 7.1: Summary of TOC Retention in the GE Control and DNAPL Columns .....	61
Table 7.2: Summary of SRS <sup>TM</sup> mass balance results .....	63
Table 7.3: Summary of nBA Additions and Mass Recoveries .....	64
Table 7.4: Summary of Chlorinated Ethene and Ethene Mass Balance Results .....	74
Table 7.5: Summary of chloride mass balance Results.....	75
Table B1: Summary of nBA demand calculations.....	90
Table B2: Summary of SRS <sup>TM</sup> demand calculations .....	92
Table C1: Bromide tracer test data .....	95
Table D1: Summary of hydrogen consumption calculations.....	99

## List of Abbreviations

AGW	artificial groundwater
cm	centimeter
cm/day	centimeter per day
CO <sub>2</sub>	carbon dioxide
CHC	chlorinated hydrocarbon
<i>cis</i> -DCE	<i>cis</i> -1,2-dichloroethene
COC	contaminant of concern
Dhc	<i>Dehalococcoides</i>
°C	degrees Celsius
°C/min	degrees Celsius per minute
DNAPL	dense non aqueous phase liquid
DAP	diammonium phosphate
DHG	dissolved hydrocarbon gases
EVO	emulsified vegetable oil
GC	gas chromatograph
gene copies/L	gene copies per liter
g	grams
g/cm <sup>3</sup>	grams per centimeter cubed
g/L	grams per liter
g/mol	grams per mole
H <sub>2</sub>	hydrogen
IC	ion chromatograph
Kg	kilogram
LCFA	long chain fatty acid
MCL	maximum contaminant level
MCLG	maximum contaminant level goal
µg/L	micrograms per liter
µM	micromolar
mL	milliliter
mL/min	milliliters per minute

mg/L	milligram per liter
mm	millimeter
mm Hg	millimeters of mercury
mM	millimolar
mV	millivolt
min	minutes
nBA	n-butyl acetate
nM	nanomolar
PED	partitioning electron donor
pmol	picomoles
PCR	polymerase chain reaction
RPM	revolutions per minute
RPD	relative percent difference
rRNA	ribosomal RNA
PCE	tetrachloroethene
TOC	total organic carbon
TCE	trichloroethene
VC	vinyl chloride
<i>vcrA</i>	vinyl chloride reductase A gene
VFA	volatile fatty acid
VOC	volatile organic compound

# Chapter 1

## Introduction

### 1.1 Introduction

Chlorinated solvents, notably trichloroethene (TCE) and perchloroethene (PCE, also referred to as tetrachloroethene) are among the most common contaminants in soil and groundwater (Ferguson and Pietari, 2000; Cope and Hughes, 2001). The limited aqueous solubility of TCE and PCE often leads to TCE and PCE contaminated sites characterized by the presence of dense non-aqueous phase liquid (DNAPL) source zones. These source zones provide a long term source of groundwater contamination as dissolution of the source zone occurs for up to hundreds of years under natural conditions (Johnson and Pankow, 1992).

Most of the contaminant mass at a DNAPL site is found in the source zone and much attention has been given to explore ways to contain and/or remove the source zone. Typical remediation technologies used include pump and treat, vapor stripping, solvent extraction, thermal technologies, and chemical oxidation (Aulenta et al., 2007). These remedial options typically have high operating costs, prolonged timelines and are often ineffective at treating the DNAPL source zone. More recently there has been an interest in the use of enhanced anaerobic biodegradation to treat DNAPL source zones. This technology has the potential to be both less expensive and faster than other technologies in treating chlorinated solvent source zones (ITRC, 2005).

Bioremediation has been shown in many studies to be an effective remedial option for treating down gradient plume concentrations from a source zone, and more recently appears to be a promising remedial approach for source zones. Biostimulation involves the addition of electron donors and/or nutrients to the subsurface in order to stimulate indigenous anaerobic bacteria present at a contaminated site to transform the target compounds (ITRC, 2005), while bioaugmentation is the introduction of microorganisms capable of transforming a target compound in contaminated media (Adamson et al., 2003). Da Silva et al. (2006) concluded that bioaugmentation enhanced PCE mass removal 1.6 times over biostimulation alone in experimental controlled release system tanks, Yang and McCarty (2000) measured a 5-fold increase in PCE DNAPL dissolution in biological column

experiments and Cope and Hughes (2001) showed that biological activity could enhance the rate of PCE dissolution by up to 16 fold. These laboratory studies were all performed with PCE, which has a lower solubility value than TCE (180 mg/L versus 1,100 mg/L respectively). Although far fewer DNAPL studies have been performed with TCE, Neilson and Keasling (1999) previously demonstrated that an uncharacterized microbial consortium was capable of dechlorination of saturation levels of TCE in microcosm studies and Harkness et al., 1999 showed that TCE concentrations up to 170 mg/L were dechlorinated in bioaugmented columns.

Biodegradation does not act directly on the free phase DNAPL, but depletes the concentrations from groundwater near the DNAPL:water interface. This increases the concentration gradient, which enhances diffusion – driven mass transfer from the DNAPL to the aqueous phase where the degradation occurs (ABA, 2005). Reductive dechlorination of the parent compounds creates the more soluble daughter products *cis*-1,2-dichloroethene (*cis*-DCE) and vinyl chloride (VC). Although it would be ideal to have complete dechlorination to ethene enhanced dissolution of TCE and the production of *cis*-DCE and VC have advantages because they increase the effectiveness of down gradient treatment systems (pump and treat) and the daughter products are biodegradable through aerobic processes, such as cometabolism (Adamson et al., 2003 and Yang and McCarty 2000). The primary goal of bioenhanced DNAPL dissolution and source treatment is not typically to reduce contaminant concentration to regulatory levels but rather to achieve and maintain a high flux of contaminant from the DNAPL and sorbed phases to the aqueous phase (Aulenta et al., 2006).

With the use of biodegradation to treat DNAPL source zones being recognized, attention has been focused on the use of electron donors that partition into the DNAPL phase and provide a source of electron donor at the DNAPL:water interface. If these electron donors can be introduced into the subsurface and come in contact with the DNAPL, they have the potential to partition into the DNAPL phase and then slowly dissolve back into the water phase providing optimum conditions for dechlorination to occur at the DNAPL:water interface and enhance dissolution. These electron donors can also be cheaper to apply because they do not require continuous or batch additions in order to maintain their effectiveness (Harkness, 2000).

SABRE (source Area BioREmediation) is a public/private consortium of twelve companies, two government agencies, and three research institutions whose charter is to determine if enhanced



anaerobic bioremediation can result in effective and quantifiable treatment of chlorinated solvent DNAPL source areas. The focus of the 4-year, \$5.7 million dollar research and development project is a field site in the United Kingdom containing a DNAPL source area with groundwater concentrations exceeding several hundred mg/L of TCE. Unique features of the SABRE study include a systematic attempt to quantify the effectiveness of bioremediation to treat DNAPL in the field, treatment of TCE (as opposed to PCE) DNAPL, and the consideration of slow release and partitioning electron donors such as emulsified vegetable oil (EVO), hexanol, and n butyl acetate (nBA). Prior to field application laboratory microcosm and column experiments were performed. Four laboratories (GE, Dupont, Terra systems, and SiREM) performed microcosm studies, while GE, Dupont, and SiREM also conducted column studies.

## **1.2 Research Objectives**

The focus of this report is on a column study that was performed at the SiREM lab. Important information from the GE column study will be used as required to support observations from this study. The objectives of this study were to:

1. Determine if reductive dechlorination is inhibited by the presence of pure phase TCE when partitioning substrates are used as electron donors.
2. Compare 2 partitioning electron donors for (1) the extent of reductive dechlorination in the presence of TCE DNAPL (2) retention in column, and (3) ability to sustain activity over time.
3. Estimate the extent of DNAPL dissolution enhancement created by the bioremediation process.

This thesis is organized into 8 chapters.

Chapter 2 presents background information

Chapter 3 describes the materials and methods

Chapter 4 describes the batch studies that were performed

Chapter 5 presents the column study experimental design

Chapter 6 presents the results from the column study

Chapter 7 provides a discussion of the column study results

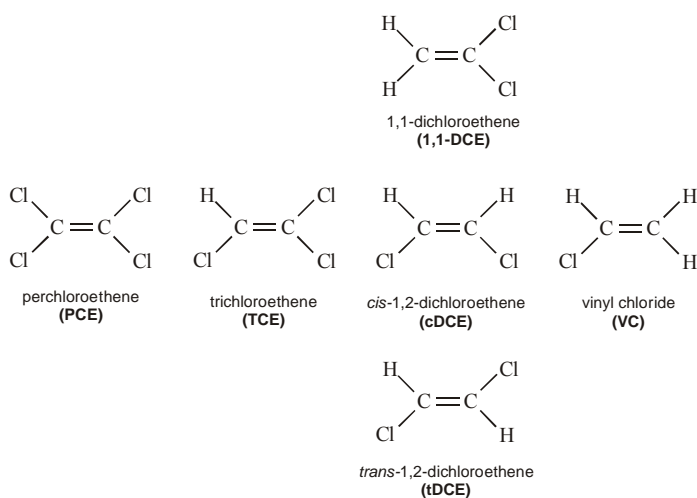
Chapter 8 states the conclusions from this study and planned future work

## Chapter 2

### Background Information

#### 2.1 Trichloroethene and DNAPLS

TCE is a chlorinated hydrocarbon compound (CHC). The structures of TCE and other chlorinated ethenes are shown in Figure 2.1. Since the 1950s TCE has been used in many industrial processes as a degreasing agent for metal parts and textiles and many other general solvent purposes such as in paints, paint strippers, and adhesives (Agency for toxic Substance and Disease Registry, 1997).



**Figure 2.1** Structures of chlorinated ethenes

TCE contaminated soil and groundwater has resulted due to past disposal directly onto land, accidental spills, and leaking storage tanks. TCE is a contaminant of concern at over 500 National Priority List sites in the United States (EPA, 2004). The maximum contaminant level goal (MCLG) for TCE is set at zero, while the maximum contaminant level (MCL) is 0.005 mg/L. Exposure to high concentrations of TCE can cause headaches, drowsiness, and eye nose, or skin irritation and long-term exposure to TCE concentrations in drinking water can cause damage to the liver, kidney, nervous, and immune system (Prager, 1996). There is some evidence to suggest that exposure to TCE over many years may be linked to several types of cancer (Prager, 1996). The physical properties of interest of TCE and other chlorinated ethenes are shown in Table 2.1.

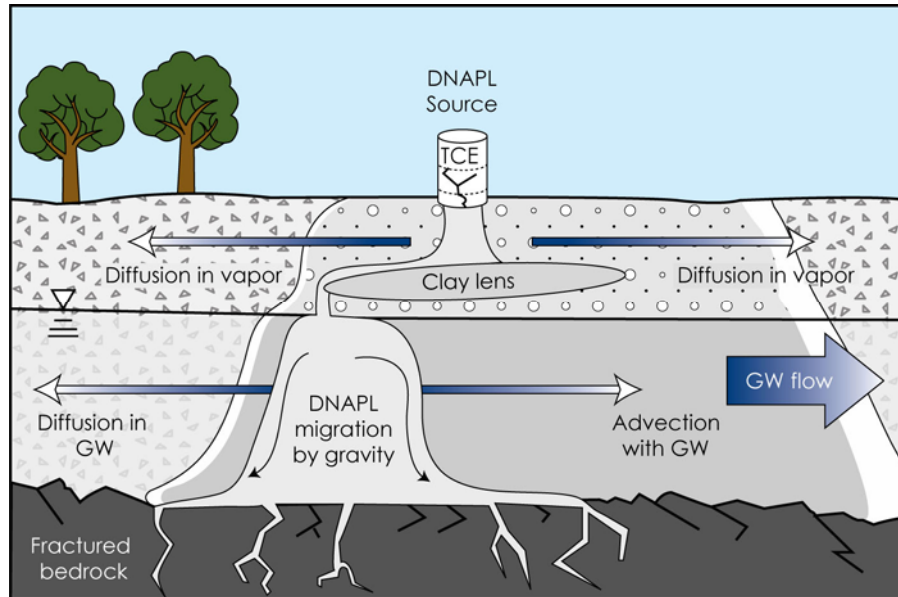
**Table 2.1** Physical and Chemical Properties of chlorinated ethenes and ethene (From Montgomery, 2000 and Mackay *et al.*, 1993)

	Molecular Weight (g/mol)	Aqueous solubility (mg/L)	Density (g/cm <sup>3</sup> )	Vapor Pressure (mmHg)	Log K <sub>OC</sub>
TCE	131.4	1,100	1.464	74	1.81
Cis-DCE	96.94	3,500	1.284	200	1.50
VC	62.5	2,700	0.9106	2,531	0.39
Ethene	28.5	131	--	22,800	--

The physical properties of TCE determine the fate of TCE in the environment. The solubility of TCE is relatively low, but because it is much higher than the MCL, large volumes of contaminated groundwater can result from small volumes of TCE. Also, the solubility of TCE is several times higher than the drinking water limit of 0.005 mg/L set by the US EPA (Norris *et al.*, 1994). With a relatively low partitioning coefficient into the organic phase of soil (K<sub>OC</sub>), TCE is not strongly retarded relative to the rate of groundwater flow allowing plumes to travel long distances (Russel *et al.*, 1992). TCE is denser and less viscous than water resulting in rapid rates of subsurface migration and eventually in the formation of DNAPLS, which can be extremely difficult to remediate due to their inaccessibility and insolubility, potentially leading to long term contamination of aquifers (Isalou and Sleep 1998).

When TCE is released it can move through the unsaturated zone, below the water table and migrate with groundwater flow. TCE will not migrate downwards uniformly, but will migrate along multiple pathways both vertically and laterally. Hysteretic capillary forces cause retention of a portion of the TCE within the pores as disconnected blobs and ganglia, which is referred to as residual DNAPL (Environment Agency, 2003). The blobs and ganglia are held in place by capillary forces that arise because the interface between the DNAPL and water is in a state of tension (Environment Agency, 2003). Residual DNAPL saturations have been observed in a sand aquifer below the water table at 1-15% (Kueper *et al.*, 1993). TCE can also form pools, which contain TCE that is continuous between adjacent pores and have saturations up to approximately 70 percent of the pore space (Environment Agency, 2003). TCE pools tend to form above finer grained horizons such as clay or silt units.

Residual DNAPL and pooled DNAPL form what is commonly referred to as the source zone. DNAPL mass tends to dissolve slowly into flowing ground water, serving as a long term source of groundwater contamination (Aulenta et al., 2006). The main processes affecting the fate and transport of DNAPLs in groundwater are shown in Figure 2.2.



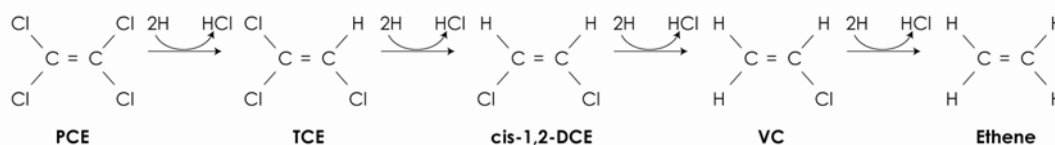
**Figure 2.2** DNAPL Transport Processes (U.S. EPA, 2000)

## 2.2 Bioremediation

Bioremediation is the use of microorganisms to degrade contaminants in soil and groundwater. When the organisms capable of degrading the contaminants of concern (COCs) are present at a site biostimulation may be used to enhance degradation. Biostimulation involves the addition of electron donors and/or nutrients to the subsurface in order to stimulate indigenous anaerobic bacteria present at a contaminated site to transform the target compounds (ITRC, 2005). If the organisms required to carry out the degradation are not present at a contaminated site then bioaugmentation may be required. Bioaugmentation is the introduction of microorganisms capable of transforming a target compound in contaminated media (Adamson et al., 2003).

## 2.2.1 Reductive Dechlorination

Freedman and Gossett (1989) were the first to discover that PCE and TCE can be readily biodegraded anaerobically by reductive dechlorination to the non hazardous end product ethene (Figure 2.3)



**Figure 2.3** Pathway for the reductive dechlorination of chlorinated ethenes

It should be noted that TCE and dechlorination products can also be mineralized to carbon dioxide (CO<sub>2</sub>) cometabolically by several organisms under aerobic conditions following a different pathway than reductive dechlorination (Duhamel, 2005). However, this will not be discussed in detail because several studies have shown that TCE degradation is most efficient following the reductive dechlorination pathway under anaerobic conditions.

In reductive dechlorination the chlorinated compounds serve as electron acceptors for microorganisms and hydrogen serves as the electron donor. In the reduction-oxidation (Redox) reaction the electron donor is oxidized and the electron acceptor is reduced. As seen in Figure 2.3 reductive dechlorination involves the stepwise replacement of chlorine atoms with hydrogen atoms thus releasing chloride at each step. The predominant intermediate formed is *cis*-DCE, although small quantities of *trans*-1,2-dichloroethene may be produced cometabolically (McCarty, 1994). Each step of the dechlorination sequence yields energy; some organisms are able to use this energy for growth, while others perform this function fortuitously (Simmonds, 2007). It is preferable to exploit growth-related pathways for bioremediation because this ensures the most efficient use of added nutrients (Duhamel, 2005). The microorganisms capable of obtaining growth from the reduction of chlorinated compounds are known as dehalorespiring bacteria and include members from a large number of genera, such as *Anaeromyxobacter*, *Desulfitobacterium*, *Trichlorobacter*, and *Dehalobacter* (Smidt and De Vos, 2004).

Several types of organisms have been shown to dechlorinate PCE and TCE to *cis*-DCE, including *Sulfurospirillum multivorans*, *Dehalobacter restrictus*, *Desulfuromonas michiganensis*, *Desulfitobacterium* sp. PCE-1 (Da Silva et al., 2006), and Geobacter strain SZ (Sung et al., 2006). The dechlorination of PCE and TCE to *cis*-DCE, which accumulates, is observed at many chlorinated solvent contaminated sites and is often referred to as “*cis*-DCE stall”. *Dehalococcoides* spp. (Dhc) is the only group of organisms known to completely dechlorinate PCE and TCE to ethene (Da Silva et al., 2006 and Maymo-Gatell et al., 1999).

There have been five *Dehalococcoides* strains isolated to date, strain 195 (Maymo-Gatell et al., 1997), strain CBDB1 (Adrian et al., 2000), strain BAV1 (He et al., 2005), strain FL-2 (He et al., 2003), and strain GT (Sung et al., 2006). These Dhc strains use hydrogen (H<sub>2</sub>) as the electron donor and PCE and TCE as electron acceptor to produce VC. Of these five strains only two (BAV1 and GT) have been shown to be capable of growth on VC, while the others cometabolically dechlorinate VC to ethene (Da Silva et al., 2006). Different strains of Dhc have been shown to be specialized to different dechlorination steps (Amos et al., 2007) and since only two strains are known to metabolically reduce VC to ethene it is important to know what type of Dhc is present at a field site. If Dhc are not present, or if the wrong strain of Dhc is present, a contaminated site may require bioaugmentation to remove the contaminants. The vinyl chloride reductase A gene (*vcrA*) produces the corresponding enzyme which is capable of converting VC to ethene (Muller et al., 2004). *vcrA* or closely related genes with similar activity are found in many (but not all) Dhc bacteria. Hendrickson et al. (2002) tested samples from 24 chlorinated ethene contaminated sites for the presence of Dhc populations and observed a positive correlation between ethene formation and the presence of Dhc. Lu et al. (2006) collected samples from eight field sites and determined that if Dhc DNA is detected at a density of 1 X 10<sup>7</sup> Dhc gene copies per liter (gene copies/L) of groundwater then natural attenuation is likely occurring at a generally useful rate.

### **2.2.2 KB-1<sup>®</sup> Culture**

KB-1<sup>®</sup> is a commercial dechlorinating culture produced and distributed by SiREM located in Guelph, Ontario. The culture originated from a former manufacturing site in South Western Ontario in 1996. The site was co-contaminated with TCE and methanol. High concentrations of ethene were also detected in the Site groundwater, suggesting that organisms capable of complete dechlorination to ethene were active in situ. A microcosm study was conducted which led to the development of

enrichment cultures maintained in Professor Elizabeth Edwards laboratory at the University of Toronto as described in Duhamel et al. 2002. In 2003 the culture was transferred to SiREM lab where the culture was scaled up to larger volumes for use as a bioaugmentation culture at contaminated sites. Over 100 sites have been bioaugmented with KB-1<sup>®</sup> to date. KB-1<sup>®</sup> is a Dhc dominated culture with approximately  $10^{11}$  Dhc gene copies/L. The culture contains the following categories of organisms:

- Bacteria that reductively dechlorinate chlorinated ethenes (dechlorinators);
- Bacteria that ferment organic compounds;
- Methanogenic Archaea (methanogens);
- Sulfate reducing bacteria; and
- Relatively low proportion of facultative aerobic bacteria typically found in soil;

The microbial composition of KB-1<sup>®</sup> is relatively complex (i.e., it is not a pure culture) due to its origin as an enrichment from a diverse natural microbial community. Anaerobic dechlorinating consortia rely on a synergistic web of activities of different groups of microorganisms to ultimately achieve the reductive dechlorination of chlorinated ethenes.

### **2.2.3 Electron Donors**

The presence of an electron donor is required for reductive dechlorination to occur. The most utilized electron donor for reductive dechlorination is H<sub>2</sub> and Dhc are dependent on H<sub>2</sub> as an electron donor (Yang and McCarty, 1998). Several studies have attempted to determine the thermodynamic threshold concentration of H<sub>2</sub> required for dechlorination to occur. Fennel and Gossett (1998) suggested it was at least as low as 1.5 nanomolar (nM); Yang and McCarty (1998) observed a threshold of 2.2 +/- 0.9 nM; and Smatlak et al. (1996) indicated a value of less than 2 nM. The H<sub>2</sub> needed for reductive dechlorination is often produced from fermentation of other substrates. Many electron donors suitable of supplying hydrogen for promoting reductive dechlorination of chlorinated solvents by anaerobic bacteria are available. Sugars (e.g., molasses), organic acids (e.g., lactate, formate, butyrate, propionate, and benzoate), alcohols (e.g., methanol, ethanol), and yeast extract have been most widely used in enhanced biodegradation applications to date. Fermentation of these organic compounds leads to the production of the necessary H<sub>2</sub> for dechlorination to occur. These

substrates are soluble in water and are highly biodegradable and therefore may need to be added periodically to maintain supply (Harkness 2000).

Insoluble electron donors have seen increasing application in enhanced biodegradation projects. These carbon sources biodegrade slowly over time and include substances like lactic acid polymers, soybean oil, chitin, and wood chips. These electron donors slowly hydrolyze, biotically or abiotically, into more soluble compounds (e.g. vegetable oils are hydrolyzed into soluble long chain fatty acids [LCFAs]) providing a slow-fermenting source of H<sub>2</sub> (Aulenta et al., 2006). They do not require continuous or batch additions in order to maintain their effectiveness and therefore can be cheaper to apply due to reduced substrate addition and system maintenance costs (Harkness, 2000).

### **2.2.3.1 Partitioning Electron Donors**

Yang and McCarty (2002) compared different electron donors to bioenhance DNAPL dissolution in column studies. They found that olive oil exhibited the greatest conversion of PCE to ethene and was the most efficient substrate compared to pentanol and oleate. This led to more focus on partitioning electron donors (PEDS) for use in treating DNAPL zones. If PEDs can be effectively introduced into the subsurface near the DNAPL:water interface, these substrates will partition into the DNAPL phase and then slowly dissolve into the water phase with the DNAPL, providing a long-term source of electron donor to support the reductive dechlorination of the dissolving solvent. This will result in much higher and sustained concentrations of electron donor at the DNAPL:water interface than could be achieved with existing electron donor delivery methods and will promote the growth of dechlorinating biomass close to the DNAPL and enhance DNAPL dissolution rates. The two PEDs tested in this study are n-butyl acetate (nBA) and slow release substrate (SRS<sup>TM</sup>).

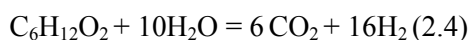
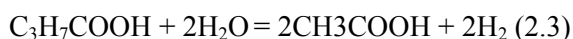
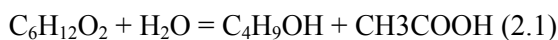
#### n-butyl acetate

The compound nBA (also known as butyl ethanolate) is readily available and costs 0.60 to 0.64 dollars per pound (Chemical Market Reporter, 2001). The solubility of nBA is reported to be 8,145 milligrams per liter (mg/L) (Montgomery, 2000), but was measured to be 6,000 mg/L by SiREM for this study, and has a log K<sub>OW</sub> value of 1.82 (Montgomery, 2000). A specific TCE partitioning coefficient of 458 was calculated for this study and is discussed in more detail in Section 4.1.2.



nBA is a clear, volatile, flammable organic solvent with a sweet odour (David et al., 2001). It is an important solvent in the chemical industry primarily in the manufacturing of paint and coatings and in the lacquer industry. Because of its lower impact on the environment, nBA is often used to replace the toxic and teratogenic ethoxy ethyl acetate (Steinigeweg and Gmehling 2002). It has been approved for use as a food additive by the FDA.

nBA is known to rapidly hydrolyze to n-butanol and acetic acid (David et al., 2001). Each mole of nBA produces one mole of n-butanol and one mole of acetic acid. The n-butanol produced can then be further broken down to 2moles of butyric acid, 2 additional moles of acetic acid, and 2 moles of H<sub>2</sub>. The butyric acid produced can further be broken down to 2 moles acetic acid and 2 additional moles of H<sub>2</sub>. If the acetic acid produced during these reactions is fully oxidized then 16 moles of H<sub>2</sub> is produced per mole of nBA. The reactions of interest are as follows:



The use of nBA as an electron donor for bioremediation could not be found in the literature.

However, nBA was used in the microcosm phase of the Sabre project with successful results as is discussed in Chapter 4.

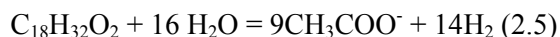
### SRS

SRS<sup>TM</sup> is a commercially available EVO product manufactured by Terra Systems, Inc., Wilmington, Delaware. SRS<sup>TM</sup> is designed to release bio-available hydrogen over long periods of time (3 to 5 years). SRS contains the following: soybean oil (60%), sodium lactate (4%), food grade emulsifier package (6%), proprietary nutrient package (<1%), and water (30%).

Edible oils are relatively inexpensive, innocuous, food-grade substrates that have been used to stimulate in situ anaerobic bioremediation processes in both barriers and source area treatments (Zawtocky et al., 2004). As discussed above vegetable oils are relatively immobile and they do not require continuous or batch additions in order to maintain their effectiveness and therefore can be cheaper to apply (Harkness, 2000). Edible oils can be injected as pure oil, or as oil in water emulsions. Achieving effective distribution of pure oil is difficult because of the limited spread of the

oil and large amount of chase water needed to displace the oil (Zenker et al., 2000). Also, the oil tends to migrate upward due to buoyancy effects (Lee et al., 2005), and results in high residual saturations and high permeability loss (Long and Borden, 2006). To overcome these problems oil in water emulsions have been developed to (1) be stable for extended time periods; (2) have small, uniform droplets to allow transport in most aquifers; and (3) have a negative surface charge to reduce droplet capture by the solid surfaces (Long and Borden 2006).

Soybean oil initially hydrolyses releasing glycerol and LCFAs (Hanaki et al., 1981). It has been hypothesized that LCFAs initially sorb to sediment surfaces, and then are slowly fermented via beta oxidation to acetate and H<sub>2</sub> (Long and Borden, 2006). The main fatty acid found in soybean oil is linoleic acid (May, 2008). Therefore linoleic acid is often used to represent soybean oil in stoichiometric calculations for bioremediation studies. Linoleic acid is fermented to 9 moles of acetate and 14 moles of H<sub>2</sub>. If acetate is completely oxidized to CO<sub>2</sub> an additional 28 moles of H<sub>2</sub> is produced. The reactions for linoleic acid are as follows:



SRS has been used at many manufacturing sites throughout the USA and was used in the microcosm phase of the Sabre project with successful results as is discussed in Chapter 4.

It should be noted that although these equations show acetate being oxidized and releasing H<sub>2</sub>, it has been shown that acetate is not likely an electron donor for dechlorination activity (Freedman and Gossett, 1989). However, acetate can be used directly by sulfate reducers (Widdel, 1988) and acetoclastic methanogens (Yang and McCarty 2002). The complete oxidation equations are shown to represent the maximum amount of H<sub>2</sub> available from these compounds.

## 2.2.4 Important Side Reactions

There are several other reactions that occur in the subsurface that are important to understand when biostimulation and/or bioaugmentation are chosen as a remedial technology. The addition of electron donor may stimulate the activity of dechlorinating microorganisms along with competing populations such as methanogens, acetogens, sulfate reducers, and nitrate reducers (Aulenta et al., 2007), the most important of these, in most situations, being the methanogens and sulfate reducers. The

competition for the fermentation products ( $H_2$  or acetate) of the supplied electron donor is an important factor in determining the rate and extent of dechlorination, the amount of electron donor necessary for dechlorination to occur, and the success of the process (Aulenta et al., 2002).

Methane gas produced by methanogens is a problem in bioremediation systems because of the wastage of electron donor and clogging caused by significant biomass and methane gas production (Yang and McCarty 2002). Competition from methanogens for  $H_2$  can potentially reduce the effectiveness of electron donor's added (Isalou and Sleep 1998), but are not actually detrimental to the dechlorinators (Adamson et al., 2003). Methane is produced either by hydrogen-utilizing (autotrophic) methanogens, or acetate-utilizing (acetoclastic) methanogens (Yang and McCarty, 2000). The  $H_2$  threshold for methanogens is reported to be in the range of 5-95 nM (Aulenta et al., 2006), which is considerably higher than the  $H_2$  threshold for dechlorination discussed in section 2.2.3 suggesting that dechlorinators have the potential to out compete methanogens when  $H_2$  is present at low concentrations.

The role of sulfate is controversial primarily because both dechlorinators and sulfate reducers are able to thrive at very low and similar  $H_2$  levels (Aulenta et al., 2007). Heimann et al. (2005) concluded that sulfate can negatively affect dechlorination at low  $H_2$  concentrations, while Aulenta et al. (2007a) concluded that dechlorination can take place under sulfate reducing conditions and that dechlorinators may benefit from  $H_2$  produced by sulfate reducing bacteria. Other studies have shown that sulfate had no inhibition effect on dechlorination and even that sulfate reducers were out competed by dechlorinators at low  $H_2$  concentrations (Hoelen and Reinhard, 2004), partial inhibition (Cabirol et al., 1998), or complete inhibition (Nelson et al., 2002). Fennell and Gossett (2003) suggested that, when observed, the lack of dechlorination under sulfate-reducing conditions may be the result of (1) direct competition for supplied electron donor by sulfate-reducing microorganisms, (2) inhibition by sulfate enzymes involved in dechlorination, (2) preferential use of sulfate as terminal electron acceptor, instead of chlorinated compounds with in the same organism (e.g. *Desulfitobacterium* spp.), or (4) larger predominance and faster growth kinetics of sulfate-reducing bacteria compared with dechlorinators in sulfate-rich environments. Heimann et al. (2005) also summarized that apart from just outcompeting dechlorinators for reducing equivalents there are other mechanisms for sulfate impairing dechlorination including, dehalogenating enzyme repression, sulfide accumulation, and inhibition by potential intermediates of sulfate reduction such as sulfite and thiosulfate.

To date the typical approach for enhancing dechlorination in sulfate rich environments is to dose the system with excess electron donor to first deplete the sulfate present. This approach may result in several side effects due to the excessive consumption of electron donor including bioclogging of the aquifer, explosive hazards through excess methane production, and the accumulation of secondary contaminants (Fe(II) and Mn(II)) from the reduction of Fe(III) and Mn(IV)) with resulting deterioration of groundwater quality (Aulenta et al., 2007). Therefore, it would be beneficial if this excessive dosing of electron donor was not necessary.

### **2.2.5 Bioremediation of DNAPLs**

DNAPLs dissolve slowly into groundwater and serve as long-term sources of groundwater contamination (Aulenta et al., 2006). In the past bioremediation methods focused only on plume treatment because concentrations associated with DNAPL source zones were believed to be toxic to micro-organisms. For example PCE dehalogenation by *Dehalospirillum multivorans* was inhibited when PCE concentration was higher than 0.3 millimolar (mM) (Yang and McCarty, 2000), which is well below the saturation concentration of about 0.9 mM for PCE. Yang and McCarty (2000) reported that only trace amount of *cis*-DCE was observed when TCE DNAPL was present, yielding a solution TCE concentration of 8.4 mM. Amos et al. (2007) reported that concentrations of TCE as high as 80 mg/L completely inhibited the activity of the tested dechlorinating isolates and Adamson et al. (2003) observed that PCE saturation concentrations severely inhibited dechlorination past *cis*-DCE in a mixed culture study.

In contrast, other studies have shown that reductive dechlorination of DNAPLs or at least of high concentrations of chlorinated solvents existing near DNAPL source zones can occur. Isalou and Sleep (1998) showed that ethene was produced in the lower portion of a column, where PCE and TCE were present. Adamson et al. (2003) showed strong evidence that inoculation successfully resulted in the growth of dechlorinating organisms in close proximity to the PCE-DNAPL, although dechlorination past *cis*-DCE was not observed. Nielsen and Keasling (1999) reported that a microbial culture enriched from a TCE contaminated groundwater aquifer could actually produce VC and ethene at a faster rate under saturation conditions than in a subsaturated system. In the study that Yang and McCarty (2000) reported a five times increase in PCE DNAPL dissolution rate and *cis*-DCE was the major dechlorination product, but significant amounts of VC and ethene were also formed. Also TCE concentrations up to 170 mg/L (a concentration suggesting the presence of

DNAPL) were dechlorinated after bioaugmentation in Dover soil columns (Harkness et al., 1999). In the microcosm phase of the Sabre project TCE concentrations as high as 800 mg/L were dechlorinated to *cis*-DCE and VC, and in a few cases completely to ethene.

Yang and McCarty (2000) list three benefits of the bioremediation of DNAPLs; (1) Enhanced dissolution of TCE and high concentrations of *cis*-DCE that results can reduced overall time for cleanup as well as increase the effectiveness of down gradient treatment systems, such as pump and treat, (2) dehalogenation products from PCE are biodegradable through aerobic processes, such as cometabolism, and (3) the problem of competitive utilization of added electron donor substrate is greatly alleviated with saturated solution PCE dehalogenation. The third benefit is because methanogens and other H<sub>2</sub> utilizers are often inhibited by high concentrations of PCE and TCE (Yang and McCarty 1998 and 2000, Nielsen and Keasling, 1999). This allows for more efficient utilization of H<sub>2</sub> produced from the biodegradation of organic substrates by dechlorinating organisms.

Although methanogens have been shown to be inhibited by high concentrations of PCE and TCE there are cases when methane production is observed in the presence of DNAPL. Methane generation was still occurring in a column in the presence of 600 micromolar (μM) PCE (Isalou and Sleep 1998). It was believed that methanogenesis was occurring in microenvironments not exposed to high concentrations of PCE. Yang and McCarty (2002) had a similar hypothesis for the observed methane production in columns with PCE DNAPL present. They believed that the heterogeneity of aquifer material in the columns and different spatial distribution of different organisms and substrates might play an important role in the observed reduced toxicity. This is further supported by the fact that a PCE gradient is likely to exist from the bottom of the column to the top of the column, and thus the PCE concentration in some parts of the column is likely to be below the saturation, or toxicity limit.

A similar hypothesis could be put forward for observed dechlorination in DNAPL zones in porous media. The amount of residual DNAPL retained by a typical porous medium is typically between 5 and 20 percent of the pore space in the particular lenses and laminations invaded by the DNAPL (Environment Agency, 2003), leaving much of the pore spaces not occupied by DNAPL and the opportunity for microenvironments to exist. Denis et al. (2003) observed high concentrations of *cis*-DCE and VC in a PCE DNAPL zone in columns, and suggested that some members of the microbial

community were able to withstand very high levels of PCE near the source zone, possibly in microenvironments exposed to lower concentrations of the compound.

There are complicating factors for the bioremediation of DNAPLs including (1) DNAPL exists as a separate phase and often as non uniform ganglia so donor substrate and nutrients must be delivered to near the DNAPL surface to enhance dissolution, (2) microbial growth near the DNAPL can cause a marked reduction in hydraulic conductivity over time, and (3) gas production and entrapment as a result of microbial activity can results in clogging and flow diversion around DNAPL (Yang and McCarty 2002). PEDs can be used to overcome many of these factors. PEDs will readily partition into the DNAPL phase and provide a source of electron donor near the DNAPL water interface. The conditions near the DNAPL water interface are not favorable for methanogens so gas production and clogging will be reduced as observed by Yang and McCarty (2002) when olive oil was mixed with PCE DNAPL. Other potential advantages of PEDs are (1) they are easily transported in the source area, which aids in their mixing throughout the source zone and contacting DNAPLs, (2) they are inexpensive on a per electron equivalent basis, and less expensive than other commercial electron donor products (eg., HRC<sup>TM</sup>), and (3) they are slowly metabolized; therefore, they can be transported without significant loss and efficient distribution throughout the source zone becomes feasible.

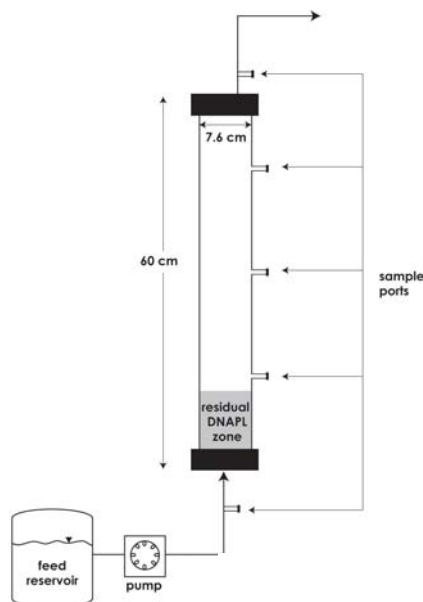
## Chapter 3

### Materials and Methods

#### 3.1 Equipment Used

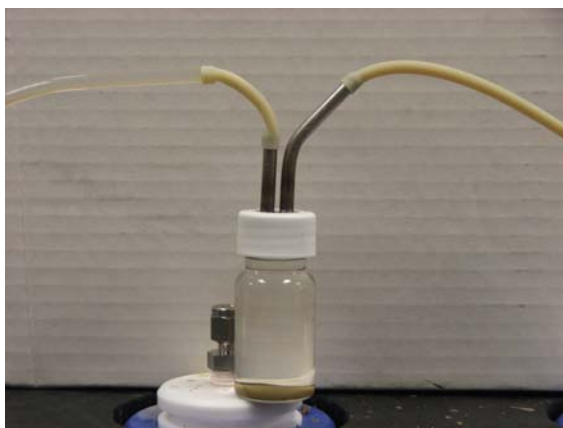
Columns were custom made by Namdar Custom Glassblowing (Mississauga, Ontario) of borosilicate glass. The columns were 60 centimeters (cm) long with a 7.5 cm inside diameter and were fitted with Teflon end caps. The bottom end cap included a coarse fritted glass filter disk to ensure uniform distribution of the influent fluid across the cross-sectional area of the column. Three sampling ports were included spaced every 15 cm along length of column. However, it should be noted here that these ports were unreliable and collecting samples from these ports was not possible throughout the column study.

The pump used to regulate flow was a Master Flex L/S precision standard drive (1-100 rotations per minute [rpm]) equipped with a Master Flex L/S multi channel head capable of producing the desired flow rate. Influent water was pumped into the columns from glass and plastic bottles or tedlar bags depending on the phase of the study. Figure 3.1 shows a schematic of the column set-up.



**Figure 3.1** Schematic of Column Study Setup

All components were connected using 1/8<sup>th</sup> inch Teflon, viton, or stainless steel tubing to minimize sorption losses. Effluent from the columns was directed to a 20 milliliter (mL) VOA vial fitted with a stainless steel flow through apparatus to allow for collection of samples (Figure 3.2). This allowed the effluent to continually flow through the VOA vial and a sample could be collected at any time with minimal losses due to volatilization.



**Figure: 3.2** Effluent collection vial

## **3.2 Analytical Procedures**

### **3.2.1 Volatile Organic Compound and Dissolved Hydrocarbon Gas Analysis**

This section describes the methods to quantify the chlorinated ethenes as well as methane, ethene and ethane (dissolved hydrocarbon gases [DHG]), nBA, and n-butanol. The Method Quantitation Limit (MQL) for the chlorinated ethenes, DHG's, and nBA ranged between 100 micrograms per liter ( $\mu\text{g/L}$ ) and 1000  $\mu\text{g/L}$  depending on the phase of the study and the sample volume analyzed.

Chlorinated ethenes, DHGs, and nBA concentrations were measured using a Hewlett-Packard (Hewlett Packard 5890 series II Plus) gas chromatograph (GC) equipped with a head-space auto sampler (Hewlett Packard 7684) programmed to heat each sample to 75 degrees Celsius ( $^{\circ}\text{C}$ ) for 45 minutes (min) prior to injection into a GSQ Plot column (0.53 millimeters x 30 meters, J&W) and a flame ionization detector. The injector temperature used is 200 $^{\circ}\text{C}$ , and the detector temperature is 250 $^{\circ}\text{C}$ . The specific headspace parameters are shown in Table 3.1. The oven temperature is programmed as follows: 35 $^{\circ}\text{C}$  for 2 min, increase to 100 $^{\circ}\text{C}$  at 50 degrees Celsius per minute ( $^{\circ}\text{C}/\text{min}$ ), then increased to 185 $^{\circ}\text{C}$  at 25 $^{\circ}\text{C}/\text{min}$  and held at 185 $^{\circ}\text{C}$  for 6.80 min, followed by an



increase to 225°C at 25°C/min and held for 0 min. The carrier gas was helium at a flow rate of 11 milliliters per minute (mL/min).

**Table 3.1** HP 7694 Parameters

Oven temperature	75 <sup>o</sup> C
Transfer line temperature	80 <sup>o</sup> C
Loop temperature	90 <sup>o</sup> C
Pressurization time	0.0 min
Loop fill time	0.2 min
Injection time	3.0 min
Vial equilibration time	45 min
GC Cycle time	21 min
Loop equilibration time	0.0 min

Chlorinated ethene, DHG, and nBA concentrations were measured by withdrawing a liquid sample from the effluent collection vial for the column effluent samples, or from the appropriate vial for batch tests, and injecting the sample into a 10-mL headspace vial containing acidified deionized water. The pH of the sample after acidification was approximately 2 to inhibit dechlorination and fermentation activity. The vial was sealed with an inert Teflon-coated septum and aluminum crimp-cap for automated injection onto the GC. Five-point calibrations were performed. Chlorinated ethene calibrations were performed using external standards purchased as standard solutions (Sigma St Louis, Missouri). The nBA standards were prepared gravimetrically in methanol using 99.5 % purity nBA (Fisher Scientific, Whitby, Ontario). DHG calibration was performed using standards that were prepared by injecting known volumes of ethene, ethane, and methane gases of greater than 99% purity (Scotty II, Alltech Associates Inc., Deerfield, IL) into 250 mL bottles containing 150 mL of water. The concentrations of DHGs were calculated using dimensionless Henry's law constants (27.20 for methane, 20.42 for ethane, and 8.76 for ethene). Standards were prepared along with samples and analyzed with each run to ensure recovery of the target analytes were within 15% of expected. Blanks were run every 20 samples to ensure there was no carry over of target analytes or instrument contamination. Data was integrated using Peak Simple Chromatography Data System Software (SRI Inc., Torrance, California).

n-butanol was analyzed by the Organic laboratory at the University of Waterloo using a 7673A Hewlett Packard Auto sampler and a Hewlett Packard 5890 gas chromatograph equipped with

a flame ionization detector and a packed column. Peaks were measured with a Hewlett Packard 3395 integrator. The method was calibrated using an external calibration mode. The column was 10 feet in length by 0.125 inches in diameter and packed with 3 % SPI500 on CaropackB (80/100) mesh. The carrier gas was helium at a flow rate of 20 ml/min. The injector temperature was 115°C, the detector was 230°C and the oven was isothermal at 145°C. The MQL for n-butanol was 0.075 mg/L.

### **3.2.2 Anions and Volatile Fatty Acids**

This section describes the methods used to quantify anions and volatile fatty acids (VFAs).

#### Anions

Anion analysis was performed on a Dionex DX-600 ion chromatograph (IC) equipped with a Dionex AS-40 auto-sampler and an AS18 column. The sample loop volume was 25 µL. An isocratic separation was performed using 33 mM sodium hydroxide eluent for 13 min. Calibration was performed using external standards of known concentrations. Five-point calibrations were performed. Standards were prepared gravimetrically using purchased salts of the highest purity from Bioshop Canada Inc. (Burlington, Ontario). Standards were analyzed along with samples to ensure recovery of the target analytes is within 15% of expected. Blanks were run every 20 samples to ensure there was no carry over of target analytes or instrument contamination. Data were integrated using Dionex's Peaknet chromatography software. The MQLs were as follows: 0.03 mg/L chloride, 0.03 mg/L sulfate, and 0.39 mg/L bromide.

Anion concentrations were measured by withdrawing a liquid sample from the effluent collection vial for the column effluent samples, or from the appropriate vial for the batch tests, and injecting the sample into a 1.5 mL micro-centrifuge tube. Samples were centrifuged for five minutes to settle out solids. The supernatant was removed, diluted 10-fold in deionized water and placed in a Dionex auto-sampler vial with a filter cap for automated injection onto the IC.

#### VFAs

VFA analysis was performed on the same Dionex DX-600 IC equipped with a Dionex AS-40 auto-sampler. For VFA analysis an AS11 high capacity column was used. The sample loop volume was 25 µL. A gradient separation using sodium hydroxide was performed as follows; 0.5 mM from 0-2.5

minutes, increase to 5 mM from 2.5 – 6.0 minutes, hold at 5 mM from 6.0-10.5 minutes, increase to 33 mM from 10.5-11.1 minutes, hold at 33 mM from 11.1-25 minutes. Calibration was performed using external standards of known concentrations. Five-point calibrations were performed. Standards were prepared gravimetrically using purchased salts of the highest purity from Bioshop. Standards were prepared along with samples and analyzed with each run to ensure recovery of the target analytes was within 15% of expected. Blanks were run every 20 samples to ensure there was no carry over of target analytes or instrument contamination. Data were integrated using Dionex's Peaknet chromatography software. The QLs are as follows: lactate 0.40 mg/L, acetate 0.54 mg/L, propionate 0.31 mg/L, formate 0.23 mg/L, butyrate 0.41 mg/L, and pyruvate 0.69 mg/L.

VFA concentrations were measured by withdrawing a liquid sample from the effluent collection vial for the column effluent samples, or from the appropriate vial for batch tests, and injecting the sample into a 1.5 mL micro-centrifuge tube. Samples were centrifuged for five minutes to settle out solids. The supernatant was removed, diluted 10-fold in deionized water and placed in a Dionex auto-sampler vial with a filter cap for automated injection onto the IC.

### **3.2.3 pH and ORP**

The pH measurements were performed using an Accumet AB15 pH meter with combination pH electrode. The pH meter was calibrated according to the manufacturer's instructions using pH standards. Calibrations of the probe were performed daily and standards were analyzed at each sampling event to ensure they were within the calibration range of the instrument at each sampling event.

The ORP measurements were performed using a Corning 313 meter with double junction ORP electrode (Ag/AgCl reference). A single point calibration of the meter was performed at each sampling event with Zobell ORP calibration solution.

### **3.2.4 Molecular Methods**

Quantitative polymerase chain reaction (PCR) was performed at SiREM using commercially available Quantitative Gene-Trac-Dhc (16S ribosomal RNA [rRNA] gene) and Gene-Trac-VC methods (vcrA gene). These proprietary methods are subject to technology licensing agreements between SiREM and DuPont and Stanford University respectively. 16S rRNA primers are similar to

those described by Hendrickson et al. (2002) and are protected under US patent US6894156B2 (Hendrickson and Ebersole), use of *vcrA* gene (Müller et al, 2004) is subject to US Patent Application USSN 60/598459 (Spormann and McCarty).

### Groundwater Filtration

Groundwater samples were stored at 4°C prior to concentration of biomass by vacuum filtration. Groundwater was decanted into a 0.20 micrometer disposable filter unit (Nalgene) and vacuum filtered. Groundwater volumes ranging from (100 ml to 500 ml) were used depending on the amount of biomass in the groundwater. Filters were removed from the housing, transferred into a sterile 50 mL centrifuge tubes and stored at -20°C until DNA extraction.

### Genomic DNA Extraction

Extraction of DNA was performed using the PowerSoil™ DNA Isolation Kit (MO BIO Laboratories, Inc, Carlsbad, California). Up to 0.25 grams (g) of solids were used per isolation. Filters were placed into Bead Solution and pulverized using a sterile pipette tip. All subsequent steps of the DNA extraction were performed according to the manufacturer's instructions. Cell lyses were performed using a MiniBeadbeater-8™ (Biospec Products, Bartlesville, Oklahoma) at 50% of the maximum setting for 30 seconds. A negative control consisting of sterile water was filtered and extracted with each set of samples to rule out background DNA or cross contamination.

### Quantitative PCR

40 µl PCR reactions consisted of 20 µl of 2X iQ™ SYBR Green Supermix, (Bio-Rad Laboratories Inc., Hercules, California) 1.6 µl of primer mix (10 picomoles (pmol) each primer) and 18.4 µl of template DNA. For quantification of total *Dehalococcoides*, primers targeting the 16S rRNA gene producing a 512-bp amplicon were used. Primers used for the quantification of the *vcrA* gene were rdhA14\_642f GAAAGCTCAGCCGATGACTC and rdhA14\_846r TGGTTGAGGTAGGGTGAAGG (Waller et al., 2005). Thermocycling was performed as follows for 16 S rRNA primers: initial denaturation of 94 °C for 2 min; 40 cycles of 94 °C for 1 min, 60 °C for 1 min and 72 °C for 1 min. with a final extension of 8 minutes, for *vcrA* primers: initial denaturation at 94 °C for 5 min with 36 cycles of 94 °C for 1 min, 58 °C for 1 min and 72 °C for 1 min with a final extension of 10 min at 72 °C.

Data collection and multi-component analysis were performed with the MyiQ™ Single-Color Real-Time PCR Detection System Software (Bio-Rad Laboratories Inc.). Standard curves of threshold fluorescence ( $C_t$ ) versus  $\log_{10}$  of the gene copy number were produced using cloned 16S rRNA and *vcrA* genes and were used to quantify the number of DNA targets in groundwater samples. In order to verify specificity and identity of the PCR products, melt curve analysis between 72 °C and 95 °C was performed.

## Chapter 4

### Batch Tests

#### 4.1 Summary of Batch Tests

Prior to the start of the column study a large multi-laboratory microcosms study was performed. Three batch tests were also conducted prior to and along with the column study. The batch tests performed were to (1) determine a partitioning coefficient of nBA into TCE DNAPL (2) determine if there is a concentration of nBA that is toxic to dechlorinators in the KB-1<sup>®</sup> culture, and (3) determine the SRS loading percentage for columns. Batch tests one and two were performed by SiREM, while batch test three was performed by GE.

##### 4.1.1 Summary of Sabre Microcosm Study

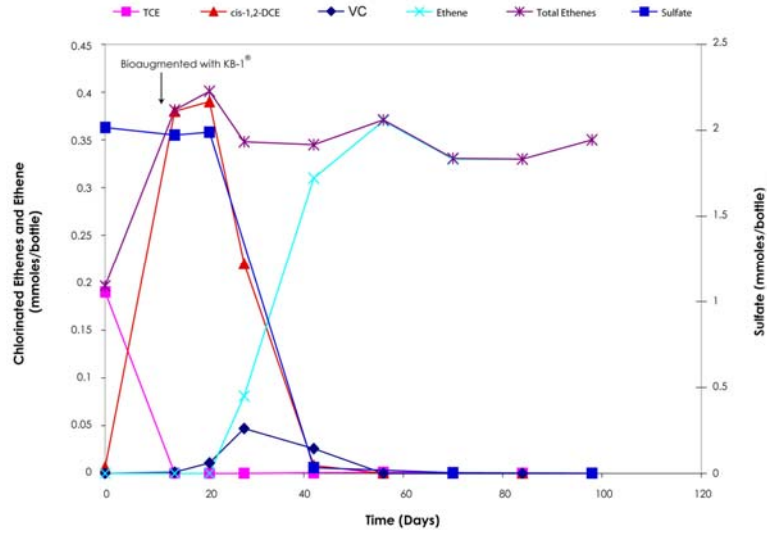
This section provides a brief review of the microcosm study that was performed prior to the column studies. A paper will be submitted for publication by the Sabre lab group of which I will be a co-author.

The objectives of the microcosm phase of the sabre program were to: (1) select the optimal electron donor for the column studies and field implementation; (2) determine if supplemental nutrients (nitrogen, phosphate, and trace minerals) are necessary or beneficial to the dechlorination process; (3) determine if bioaugmentation is necessary or beneficial to the same process; and (4) determine if the high levels of TCE present at the site will have a negative impact on the dechlorination process.

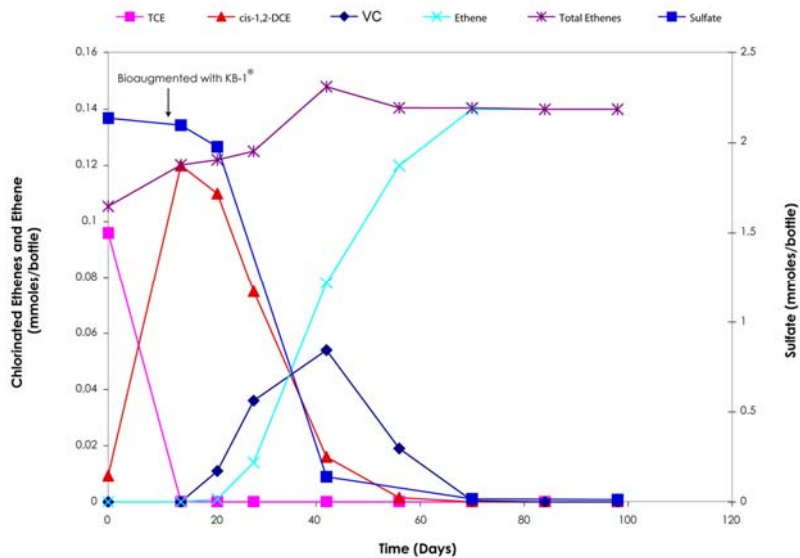
The microcosm study was designed as a fractional factorial experiment. The design consisted of 48 combinations of the levels of the experimental variables. These include four laboratories (DuPont, GE, Terra Systems, SiREM), six electron donors (Lactate, Acetate, Methanol, SRS<sup>™</sup>, Hexanol, and nBA), supplemental nutrients (With and Without), bioaugmentation with KB-1<sup>®</sup> (With and Without), and TCE (low level [100 mg/L] and high level [400 mg/L]).

The results showed that all of the electron donors promoted complete dechlorination of TCE to ethene in some microcosms during the incubation period. The classic reductive dechlorination sequence of TCE to *cis*-DCE to VC to ethene was observed in all cases. Figure 4.1 and 4.2 show examples of the

results obtained for SiREM microcosms, SRS<sup>TM</sup> amended high TCE plus nutrients plus KB-1<sup>®</sup> and nBA amended low TCE plus nutrients plus KB-1<sup>®</sup> respectively. The presence of sulfate did not inhibit TCE to *cis*-DCE dechlorination and was observed to be reduced concurrently with *cis*-DCE to VC. The pH in these microcosms was monitored throughout the incubation period and remained close to neutral indicating that the site material was capable of buffering the acid produced from the dechlorination process in the microcosm bottles.



**Figure 4.1** Chlorinated Ethenes, Ethene, and Sulfate Concentration trends in SRS<sup>TM</sup> amended high TCE plus nutrients plus KB-1<sup>®</sup> Microcosms



**Figure 4.2** Chlorinated Ethenes, Ethene, and Sulfate Concentration trends in nBA amended low TCE plus nutrients plus KB-1<sup>®</sup> Microcosms.

A re-spike experiment was conducted on twenty-nine of the microcosms that went to completion in the main microcosm study. These microcosms were re-spiked with TCE to a concentration of 800 mg/L. The results indicated that more than half of these microcosms supported dechlorination of TCE to *cis*-DCE and VC after 210 days and a few microcosms went completely to ethene. In addition to spiking with TCE selected microcosm were re-spiked with 1,000 mg/L of sulfate. Interestingly the addition of sulfate appeared to enhance the rate of dechlorination in the re-spike experiment. In all cases sulfate was reduced concurrently with *cis*-DCE to VC.

The results of the microcosm study indicate that reductive dechlorination of TCE to ethene is feasible at high TCE concentrations (up to 800 mg/L) that will be encountered in DNAPL source areas. SRS<sup>TM</sup> supported the fastest dechlorination of TCE to ethene and bioaugmentation with KB-1<sup>®</sup> and nutrient addition were shown to benefit the speed of reductive dechlorination. Therefore SRS<sup>TM</sup>, bioaugmentation, and nutrient addition were chosen as the treatment parameters of the column studies. Although nBA was not the best performing electron donor in the microcosm study complete dechlorination to ethene was achieved in many of the microcosms (e.g., Figure 4.2), and therefore nBA was chosen as an electron donor in the SiREM column study.

#### 4.1.2 Partitioning Experiment

A variation of the shake flask method described in the OECD test guideline (107) was used to calculate TCE-DNAPL water partitioning co-efficients (P) for nBA. The partitioning coefficient is defined as the ratio of equilibrium concentration ( $C_i$ ) of a dissolved substance in a two phase system consisting of two largely immiscible solvents. In this case a TCE and water system.

$$P = C_{i-TCE}/C_{(water)} \quad (4.1)$$

Three ratios of TCE to water were tested (1:1, 1:2, and 2:1) in duplicate. Six 40 mL VOA vials were capped with septa caps and weighed. TCE was added gravimetrically to the six vials at the amounts specified in Table 4.1. A 4,000 mg/L nBA stock solution was created in dionized water with a measured concentration of 3,898 mg/L. The nBA solution was used to fill all six vials so that no headspace remained; the vials were capped and weighed again. The vials were placed on a New Brunswick Scientific shaker (Edison, New Jersey) at 100 RPM for 6 days. The vials were removed from the shaker and the aqueous phase was sampled for nBA concentrations. The TCE phase was not sampled, but instead the concentration of nBA in the TCE phase was calculated as the difference in nBA concentrations from the measurements at day six and the original stock solution. The partition



coefficient was calculated as the concentration of nBA in TCE at day 6 divided by the concentration of nBA in water at day 6. An average partitioning co-efficient of 458 was calculated with a standard deviation of 32. Results are summarized in Table 4.1. Since the DNAPL phase was not analyzed directly for nBA duplicate control vials of aqueous nBA without TCE present were constructed to account for losses of nBA due to hydrolysis. The concentration of nBA in the control vials remained relatively stable over the incubation period with a decrease in nBA mass of less than 10% indicating that losses of nBA in the aqueous phase of the test vials containing DNAPL were due to partitioning of the nBA into the TCE DNAPL.

**Table 4.1** Summary of nBA Partitioning Experiment Results

	empty vial (g)	Vial and TCE (g)	Filled with NBA (g)	TCE (g)	TCE (mL)	nBA (g)	nBA (mL)	nBA in water (mg/L) T=0	nBA in water (mg/L) T=6	nBA in water (mg) T=0	nBA in water (mg) T=6	nBA in TCE (mg) T=6	nBA in TCE (mg/L) T=6	Decrease in nBA mass (%)	P
TCE1:1	29.35	60.45	82.55	31.1	21	22.1	22	3989	9.6	88	0.21	88	4140	--	430
TCE1:1	28.70	59.60	81.90	30.9	21	22.3	22	3989	9.8	89	0.22	89	4204	--	430
TCE2:	29.30	71.50	85.90	42.2	29	14.4	14	3989	4.6	57	0.07	57	1990	--	436
TCE2:1	28.90	70.95	85.55	42.1	26	14.6	15	3989	4.4	58	0.06	58	2245	--	507
TCE1:2	28.90	49.55	79.45	20.7	13	29.9	30	3989	19.8	119	0.59	119	9328	--	471
TCE1:2	29.05	49.75	79.15	20.7	13	29.4	29	3989	19.2	117	0.56	117	9151	--	477
nBA Control-1	29.1	29.10	71.60	0.00	0.00	42.5	13	3989	3652	170	155	--	--	8.4	--
nBA Control-2	29.80	29.80	73.25	0.00	0.00	43.5	14	3989	3695	173	161	--	--	7.4	--
average	--	--	--	--	--	--	--	--	--	--	--	--	--	--	458
stdev	--	--	--	--	--	--	--	--	--	--	--	--	--	--	32

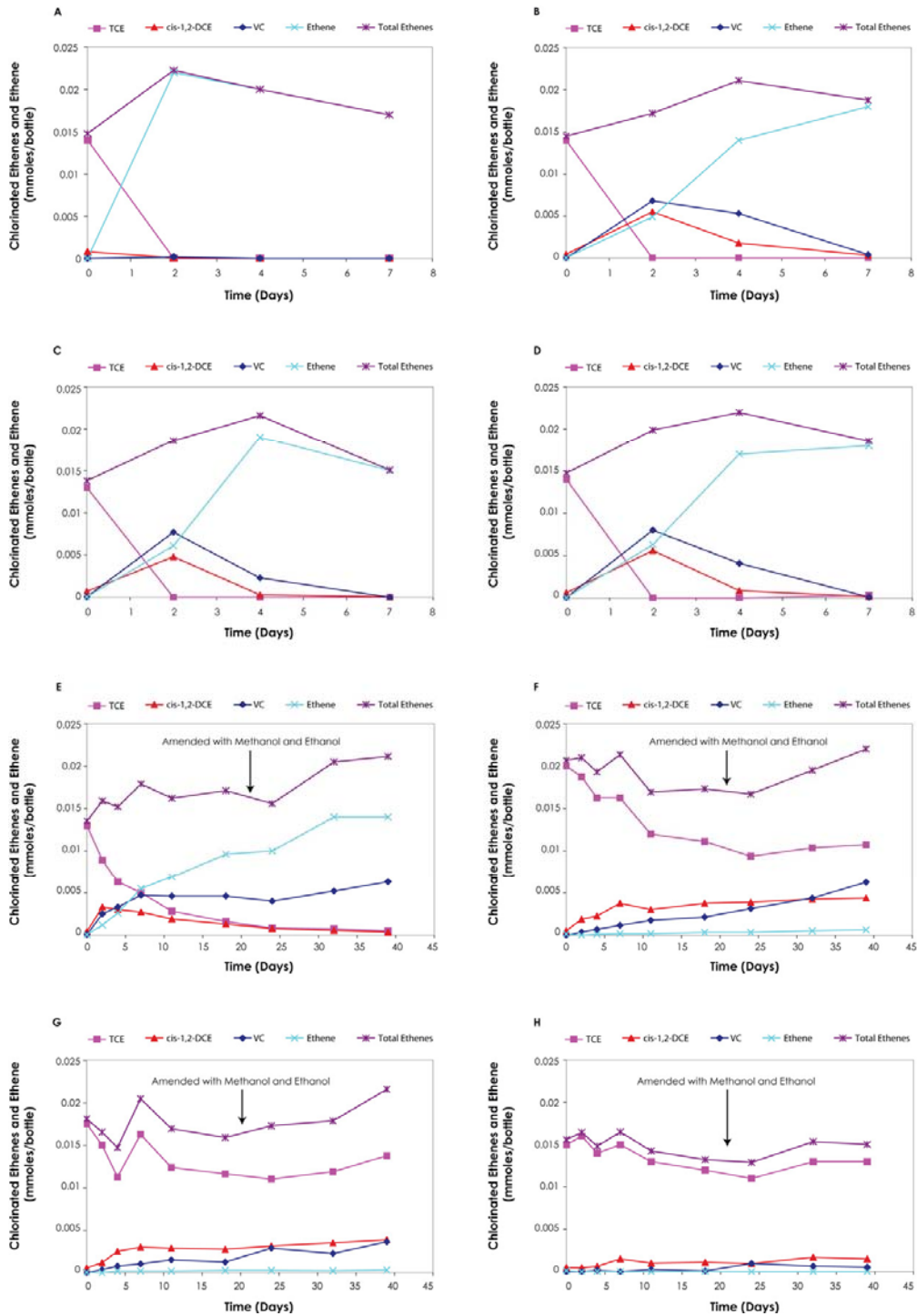
Notes:  $P = (\text{nBA in TCE (mg/L) at T=6} / (\text{nBA in water (mg/L) at T=6 days})$

### 4.1.3 nBA Toxicity Experiment

This batch test was conducted after the results were obtained from the first two nBA additions to the nBA column as discussed in Section 6.1.2. The objective of the nBA toxicity experiment was to determine if there was a concentration threshold of nBA that would inhibit reductive dechlorination of TCE by the KB-1<sup>®</sup> culture. Past work by Duhamel et al., (2002) showed that 1,1,1-trichloroethane and chloroform are potent inhibitors of chlorinated ethene dechlorination by the KB-1<sup>®</sup> culture, so it was decided to check if nBA would have a similar inhibitory effect at some level. The concentration of nBA used in the Sabre microcosm study was 500 mg/L and complete dechlorination of TCE to ethene was achieved, therefore 500 mg/L was used as the lowest concentration of nBA in this batch test. The ideal concentration used to load a DNAPL would be at the solubility limit of nBA (6,000mg/L). This corresponds to the maximum amount of nBA that could be injecting in a given volume of water and would be the most efficient way to load a DNAPL with nBA. Therefore concentrations from 500 to 6,000 mg/L of nBA were tested in this batch test.

Microcosms were constructed by filling sterile 40 mL (nominal volume) VOA vials with 35 mL of KB-1<sup>®</sup> culture. A sample was removed from the KB-1<sup>®</sup> growth vessel using a one liter transfer bottle. The culture was purged for 20 minutes with a nitrogen/carbon dioxide gas mixture to remove volatile organic compounds (VOCs) and DHGs from the culture. The culture was transferred to an anaerobic chamber (Coy Laboratory Products, Grass Lake, MI) filled with approximately 80 % nitrogen, 10 % carbon dioxide, and 10 % hydrogen (BOC Gases, Guelph, Ontario). The vials were capped with Mininert<sup>™</sup> closures to allow repetitive sampling of the vials without VOC loss, and to allow nutrient amendment, as needed throughout the incubation period. Anaerobic conditions were maintained by constructing the microcosms in the anaerobic chamber; anaerobic microcosms were incubated and sampled in the anaerobic chamber. One replicate of each treatment was amended with resazurin to monitor redox conditions. Resazurin is clear under anaerobic conditions but turns pink when exposed to oxygen. Duplicate microcosms were prepared for a positive control (methanol and ethanol amended) and for the nBA amended microcosms. All microcosms were spiked with TCE to a target concentration of 50 mg/L. The nBA concentrations tested were 500 mg/L, 1,000 mg/L, 2,000 mg/L, 3,000 mg/L, 4,000 mg/L, 5,000 mg/L, and 6,000 mg/L.

Microcosms were sampled for chlorinated ethenes and ethene for up to 39 days. Figure 4.3 shows the chlorinated ethene and ethene profiles for the different treatments over the incubation period. Complete dechlorination of TCE to ethene with excellent mass balance was observed by day 2 in the methanol and ethanol control microcosms. Complete dechlorination of TCE to ethene with excellent mass balance was observed by day 7 in the 500, 1000, and 2000 mg/L nBA microcosms. These microcosms were no longer sampled after day 7. Partial dechlorination of TCE to ethene was observed in the 3,000 mg/L nBA amended microcosms by day 7, while dechlorination was limited in the 4,000 mg/L, 5,000 mg/L, and 6,000 mg/L amended microcosms. At day 21 the 3,000 mg/L and higher nBA amended microcosms were amended with methanol and ethanol to determine if the higher concentrations of nBA were actually inhibitory to the dechlorinators or if the dechlorinators were not able to use the nBA as an electron donor at these concentrations. TCE and *cis*-DCE concentrations decreased with corresponding increases in VC and ethene in the 3,000 mg/L amended treatment to day 39. TCE and *cis*-DCE remained relatively stable in the 4,000 mg/L nBA amended microcosms with slight increases in VC. TCE concentrations remained relatively stable in the 5,000 mg/L and 6,000 mg/L microcosms with low detects of *cis*-DCE and VC. Ethene production was not observed. The microcosm study was ended at day 39.



**Figure 4.3** Chlorinated ethenes and ethene concentration trends in nBA toxicity test microcosms (A) Control; (B) 500 mg/L nBA; (C) 1,000 mg/L nBA; (D) 2,000 mg/L nBA; (E) 3,000 mg/L nBA; (F) 4,000 mg/L nBA; (G) 5,000 mg/L nBA; and (H) 6,000 mg/L nBA

#### 4.1.4 SRS<sup>TM</sup> Loading Percentage Experiment.

The objective of the SRS<sup>TM</sup> loading batch test was to quantify the equilibrium partitioning behavior of TCE in the presence of site soil, a residual TCE DNAPL phase, and varying amount of SRS<sup>TM</sup>.

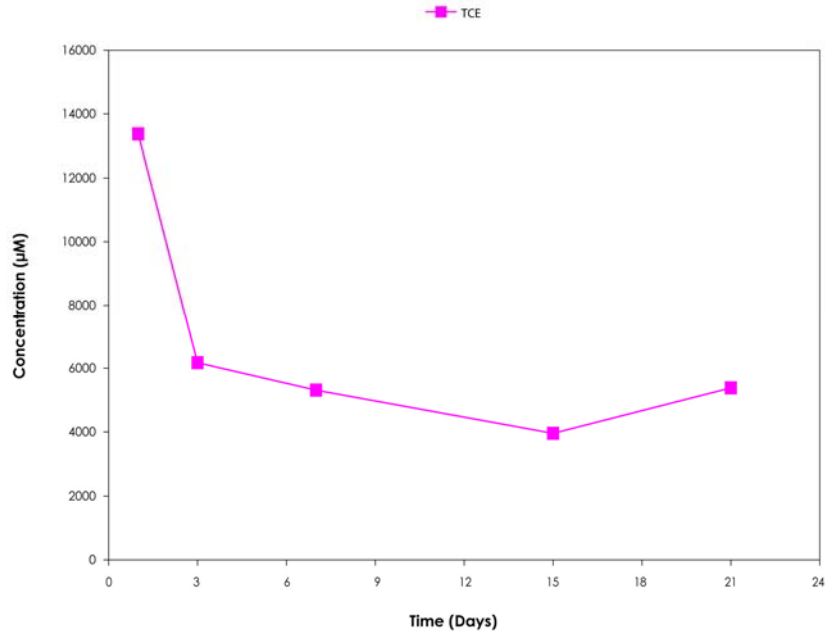
Equilibrium partition coefficients are not available for soybean oil in the neat or emulsified forms. Neat soybean oil has similar properties to octanol and may be modeled as a separate oil phase, but this may not be true of the emulsion phase as the emulsion phase is not stable. Total organic carbon (TOC) measurements were used as a surrogate for direct measurement of the soybean oil in the SRS<sup>TM</sup>.

The experiments were performed in 120 mL glass serum bottles crimp-sealed with Teflon-lined butyl rubber septa. Six bottles were constructed with 50 g of Site soil and 75 mL of dionized water. TCE (1.5 g) was added to each bottle. SRS<sup>TM</sup> was added to each bottle to give different percentages of SRS<sup>TM</sup> ranging from 0 to 6.25% (v/v). The bottles were shaken well and allowed to stand over night to allow TCE to come to equilibrium between the soil and water phases. Figure 4.4 shows the reactor bottles with soil, water, TCE and increasing amounts of SRS<sup>TM</sup>.



**Figure 4.4** SRS<sup>TM</sup> Partitioning Experiment Reactor Bottles

The aqueous phase was measured for TCE on days 1, 3, 7, 15, and 21. The TCE profile for the 4.5 % SRS<sup>TM</sup> reactor is shown in Figure 4.5. At day one TCE concentrations were elevated slightly above the solubility limit of TCE followed by a decrease at Day 3 to concentrations lower than TCE saturation which remained stable through day 21. The elevated concentrations of TCE were only temporary indicating that loading rates of SRS<sup>TM</sup> close to 4.5 % are not likely to mobilize TCE DNAPL. Also, EVO loading rates of 2-10% are commonly used in field studies and therefore a loading rate of 5% was chosen for this study



**Figure 4.5** TCE Concentration Profile in the 4.5 % SRS<sup>TM</sup> Amended Reactors

## 4.2 Conclusions from Batch Tests

The following conclusions from the batch tests were made;

- 1) SRS<sup>TM</sup> will be used in the column study because it supported the fastest dechlorination of TCE to ethene in the Sabre microcosm study;
- 2) Bioaugmentation with KB-1<sup>®</sup> and nutrient addition will be used in the column study because they were shown to enhance the rate of reductive dechlorination in the Sabre microcosm study;
- 3) nBA supported complete dechlorination of TCE to ethene in the Sabre microcosm study and will be used as a partitioning electron donor in the SiREM column study;
- 4) nBA partitions readily into TCE DNAPL with a partitioning coefficient of 458+/-32;
- 5) Concentrations of nBA above 2,000mg/L appear to inhibit complete dechlorination of TCE to ethene with the KB-1<sup>®</sup> culture;
- 6) SRS<sup>TM</sup> will be added to the columns as a 5% (v/v) solution.

## **Chapter 5**

### **Column Study Experimental Design**

#### **5.1 Column Set-up and Construction**

Three soil columns were constructed to evaluate three conditions. One column was unamended and acted as the control column, the second column was amended with nBA, nutrients, and was bioaugmented with KB-1<sup>®</sup>, and the third was amended with SRS<sup>™</sup>, nutrients, and was bioaugmented with KB-1<sup>®</sup>. The following sections describe the set-up and construction of the columns.

##### **5.1.1 Soil Collection and Artificial Groundwater**

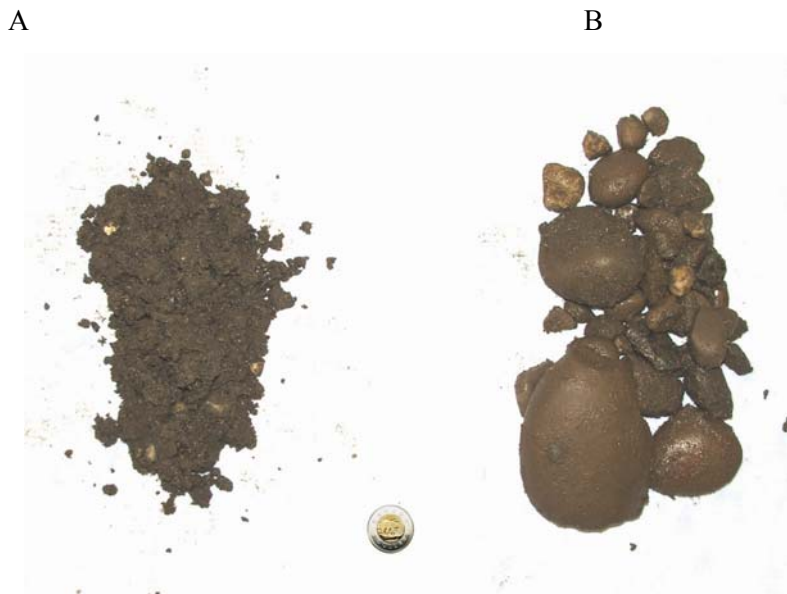
The soil used in this study was obtained from the field site in the UK. Soil samples were collected on 21 October 2005 several meters below ground surface inside the DNAPL source area using a direct push rig and 4-inch diameter, acetate sleeves. The acetate sleeves were capped and sealed with wax immediately upon retrieval. The soil was packed in iced coolers and sent by courier to SiREM and was received on 31 October 2005. The soil was stored in a cold room at 4°C until processing. The material was dominantly sandy gravel, with silty sandy gravel, gravel and minor clay and silt horizons.

Due to the expense of collecting and shipping the amount of ground water needed for the column studies an artificial ground water (AGW) was created in the lab that was designed to match the geochemistry of the site ground water. This allowed all three laboratories to make AGW as needed throughout the study. The procedure for making the AGW is described in Appendix A.

##### **5.1.2 Column Packing**

On 4 January 2006 the soil material was removed from the cold room and placed in a temporary anaerobic glove bag. The cores were opened and the material was sieved through a 12 millimeter (mm) screen to remove larger rocks and debris and then mixed by hand inside a large plastic bag to homogenize the contents. Figure 5.1 shows the material that was used to pack the columns and the larger material that was removed during the sieving process.





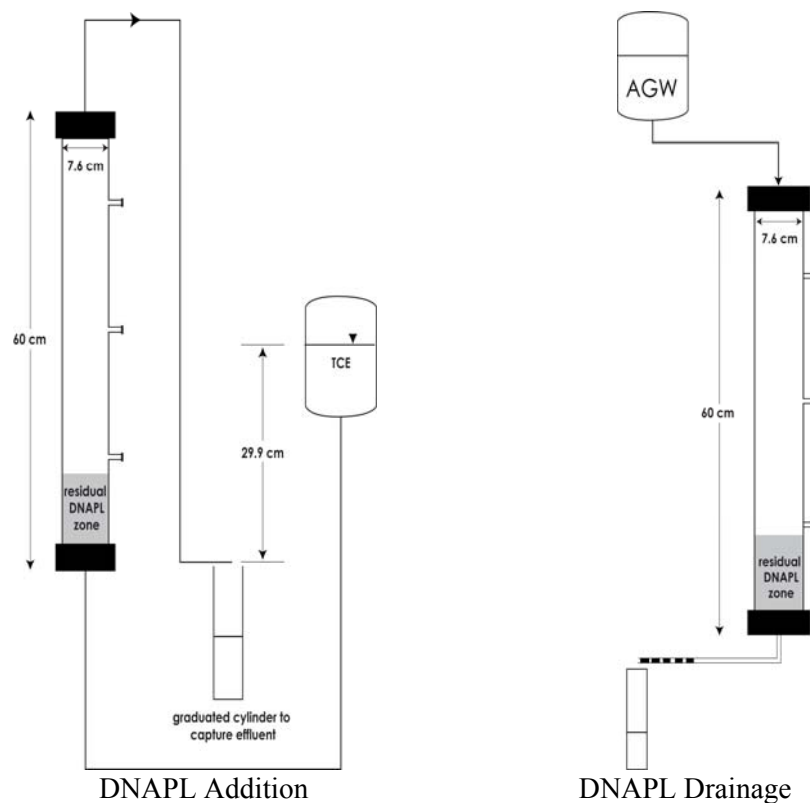
**Figure 5.1** Sabre test cell aquifer solids. (A) Material that passed through a 12 mm screen used to pack the columns, and (B) material that was removed with the 12 mm screen and was not used in the columns.

The SRS<sup>TM</sup> amended and nBA amended columns were packed on 11 January 2006, while the control column was packed on 15 June 2006 because the original control column broke and was not used in the study. A 2 cm layer of clean, non-sterilized, Federal Fine Ottawa sand (30-140 mesh) was placed in the column immediately above the end cap to prevent the study soil from entering the disk and plugging the glass frit. The site soil was added to the column in 2 cm lifts. After each lift was added the soil was uniformly compacted with a compacting tool. During column packing a constant stream of CO<sub>2</sub> gas was passed through the column to remove oxygen and minimize residual air pockets, since CO<sub>2</sub> more readily dissolves into water than oxygen. The top portion of the column (approximately 2 cm) was packed tightly with inert wet glass wool. The end cap was installed and tightened. The weight of soil added to each column was recorded. Columns 1, 2, and 3 received 4.9, 5.1, and 5.3 kilograms (kg) of site material respectively.

### **5.1.3 DNAPL Emplacement**

TCE was dyed with oil red-O at a concentration of 0.1 grams per liter (g/L) prior to adding TCE to the columns. The dye allowed visual observations of the TCE in the columns. A constant head tank containing dyed TCE connected to the bottom of the columns was used to add the TCE to the columns. The height of TCE in the constant head tank provided enough pressure at the bottom of the

column to force the TCE to displace water in the column and move upwards. The amount of TCE in the tank was recorded. The target was to create a TCE DNAPL zone in the bottom 20 cm of the column. TCE was visually observed to reach 30 cm in the nBA and SRS columns and 33 cm in the control column. After the TCE was added to a column the constant head tank was disconnected and the volume of TCE remaining in the tank was recorded. A tank of AGW was connected to the top of the column for approximately two hours to drain the free DNAPL phase and leave a residual DNAPL zone in the column. The DNAPL leaving the bottom of the column was collected in a graduated cylinder and the volume collected was recorded. The amount of TCE remaining in the column was calculated as the difference between the amount added from the constant head tank and the amount collected in the graduated cylinder. The procedure is graphically presented in Figure 5.2. The volume of TCE added to each column was different, which resulted in different saturations of the source zones of the three columns (Table 5.1).



**Figure 5.2** Schematic of DNAPL Addition and Drainage

**Table 5.1** Summary of DNAPL Additions

Column ID	Volume TCE added (mL)	Volume TCE Retained (mL)	Volume of Source Zone (mL)	DNAPL Saturation (%)
Control	73	62	467	13.3
SRS <sup>TM</sup>	35	21.5	397	5.4
nBA	70	40	451	8.9

Notes: DNAPL saturation (%) is equal to the volume of TCE retained / (porosity X volume of source zone (mL)) X 100

Volume of source zone is based on the observed height of DNAPL in the columns (33 cm, 30 cm, and 30 cm for control, SRS, and nBA columns respectively), a cross sectional area of 45.6 cm<sup>2</sup>, and the porosity of the individual columns as calculated by a tracer test described in Section 6.1 below. (Volume of source = cross sectional area X length of source zone X porosity)

## 5.2 Electron Donor and Nutrient Addition and Bioaugmentation

### 5.2.1 nBA Column Amendments

nBA was amended to the nBA column four different times throughout the column study. Table 5.2 summarizes the nBA amendments. The first two additions were as saturated stocks added for approximately one pore volume. This allowed the greatest amount of mass possible to be loaded into the column in a given volume of water. The third addition was at a concentration of 2,000 mg/L based on the results of the toxicity test described in section 4.1.3. The fourth and final addition was as neat nBA in an attempt to directly load the nBA into the DNAPL phase of the column. The total amount of nBA added throughout the column study period was equal to 1.3 times the electron donor demand based on the amount of TCE in the column and the amount of sulfate that would be pumped through the column for one year of column operation. Refer to Appendix B for details of the nBA electron donor demand calculations and for the procedures for making the nBA solutions.

**Table 5.2** Summary of nBA Amendments

Duration (Days)	Amendment	Amount nBA added (g)
47-53	6,000 mg/L	5.4
118-125	6,000 mg/L	4.9
209-246	2,000 mg/L	10.9
280	neat	5.3

A hole was drilled in the cap of a 1L Pyrex bottle containing the saturated nBA solution allowing a piece of Teflon tubing to fit through it tightly. The Teflon tubing was lowered to the bottom of the bottle. The Teflon tubing was connected to the pump tubing and the nBA solution was pumped through the column at the experimental flow rate. Due to the volatility of nBA there was the potential to have losses of the free phase nBA on top of the AGW. The saturated solution was visibly observed every day during the amendment phases to verify that the free phase of nBA was present. The solutions were also analyzed for nBA at the start and end of the amendment phases that confirmed the concentration of the solution. At the end of the saturated stock amendments the flow was stopped for 72 hours followed by the continuation of the AGW solution at the experimental flow rate.

The third nBA amendment phase used a concentrated stock solution of nBA at a concentration of approximately 2,000 mg/L. This solution was pumped through the column from days 209 to 215. On day 215 the concentration of nBA in the influent bag was measured and found to be at a concentration of 500 mg/L indicating that losses were occurring in the influent bag. On 6 November 2006 (Day 222) the bag was removed, cleaned, and a new 2,000 mg/L stock solution was made following the same procedure. On 7 November 2007 (Day 223) the new solution was pumped through the column. On 30 November 2006 (Day 246) the nBA solution was replaced with the regular AGW solution. Table 5.3 shows the concentration of nBA, acetate, and butyrate in the influent bags over time. Although only one data point was collected for acetate and butyrate the data indicate that degradation of nBA in the influent bag was occurring. The AGW contained a redox indicator, resazurin, which is blue in the presence of oxygen and turns to pink and then clear as oxygen is consumed. The AGW in the influent bag turned pink over the addition phase indicating that oxygen was being consumed by microorganisms and that the influent bag was not sterile. This microbial activity may have increased

the rate of nBA hydrolysis. Hydrolysis of nBA did not occur over the duration of the TCE partitioning batch test (Section 4.1.2) when sterile dionized water was used. This suggests that the rate of hydrolysis was faster in the influent bag, either due a component of the AGW or microbial activity. Butanol measurements were not taken, but since there was very little butyrate found, butanol would be expected to be present at an equimolar amount as acetate (assuming acetate degradation was not occurring). This is because hydrolysis of nBA forms one mole of acetate and one mole of butanol (equation 2.1). If butanol degradation were occurring butyrate would be expected as a degradation product along with more acetate. Since, butyrate concentrations were very low, butanol degradation is minimal and the amount of butanol present will be equal to the moles of acetate.

**Table 5.3** Concentrations of nBA, acetate, and butyrate in the influent bag during the 2,000 mg/L nBA amendment phase

Sample ID	Date	Day	nBA (mg/L)	Acetate (mg/L)	Butyrate (mg/L)
influent bag	20-Oct-06	205	2122	--	--
influent bag	30-Oct-06	215	500	--	--
Comment	6-Nov-06	222	bag removed, cleaned and re-amended with 2,000 mg/L of nBA		
influent bag	7-Nov-06	223	2038	--	--
influent bag	14-Nov-06	230	770	--	--
influent bag	17-Nov-06	233	267	--	--
influent bag	21-Nov-06	237	114	1691	4.0
influent bag	24-Nov-06	240	79	--	--
influent bag	27-Nov-06	243	88	--	--

The fourth and final nBA amendment was added as neat nBA directly to the column inlet. A volume of 6 mL of neat nBA was added slowly to the column inlet on 3 January 2007 (Day 280). This volume was chosen because it was similar to the amount added during the first two nBA additions. Immediately after the neat addition regular AGW was pumped through the column at the experimental flow rate. On 2 February 2007 (Day 310) a crack developed in the bottom of the nBA column that could not be fixed and the nBA column was discontinued.

### **5.2.2 SRS™ electron donor**

SRS™ was amended to the column only once during the column operational period. The total amount of SRS™ added was equal to approximately 2.2 times the electron donor demand based on the amount of TCE in the column and the amount of sulfate that would be pumped through the column for one year of column operation. Refer to Appendix B for details on the SRS electron donor demand calculations and for the procedure for making the SRS™ solution.

A 5% solution of SRS™ was created in AGW to amend the column. A hole was drilled in the cap of the Pyrex bottle and a piece of Teflon tubing was lowered to the bottom of the bottle similar to the nBA set-up. The Teflon tubing was connected to the pump tubing and the SRS™ solution was pumped through the column at the experimental flow rate. The SRS™ solution was amended to the column from 12 May 2006 to 19 May 2006 (Days 47 to 54). The SRS™ solution was placed on a stir plate and continuously stirred for the amendment period to keep the emulsion together. If the emulsion were to break separate oil and water phases would be created leading to improper loading of the electron donor into the column. The SRS™ solution was observed daily during the amendment phase. A small amount of oil was observed at the top of the SRS™ solution, but the emulsion remained intact throughout the addition phase. At the end of the SRS™ amendment phase the flow was stopped for 72 hours followed by the continuation of the AGW solution at the experimental flow rate.

### **5.2.3 Nutrient Addition, Bioaugmentation, and Buffering**

The microcosm results indicated that both nutrient addition and bioaugmentation were beneficial to the dechlorination process and therefore were used in the column study.

Bioaugmentation of both electron donor amended columns took place on 15 June 2006 (Day 78). The bioaugmentation culture used was the KB-1® culture grown and maintained at the SiREM lab. A sample was removed from the KB-1® growth vessel using a one liter transfer bottle. The culture was purged for 20 minutes with a nitrogen/carbon dioxide gas mixture to remove VOCs and DHGs from the culture. The culture was transferred to an anaerobic chamber (Coy Laboratory Products, Grass Lake, MI) filled with approximately 80 % nitrogen, 10 % carbon dioxide, and 10 % hydrogen (BOC Gases, Guelph, Ontario). For each column bioaugmentation, a 10 mL sample was removed from the transfer bottle using a 10 mL syringe that was stored in the anaerobic chamber for several days prior to use. The syringe was also rinsed with the culture three times to remove any trace amounts of

oxygen present in the syringe. The syringe filled with KB-1<sup>®</sup> was removed from the anaerobic chamber and 10 mL of KB-1<sup>®</sup> was injected directly into the column inlets over a time span of 10 minutes (flow rate of 1 mL/min). Immediately after bioaugmentation AGW was pumped through the column at the experimental flow rate. QPCR analysis performed directly on the KB-1<sup>®</sup> used to bioaugment the columns indicated that there were  $1.8 \times 10^{11}$  gene copies/L of Dhc in the anaerobic culture. This corresponds to a concentration of  $2.1 \times 10^9$  gene copies/L in the columns assuming the culture was distributed throughout the entire volume of the columns.

Nutrients were added to the treatment columns starting on 1 August 2006 (Day 125) and continued for the remainder of the column operation period. Reagent grade diammonium phosphate (DAP) (Fisher Scientific) was added to the influent AGW at a concentration of 25 mg/L. This concentration was chosen because concentrations higher than this caused precipitation to occur in the AGW solution. For comparison, the DAP used in the microcosm study was a one time addition at a concentration of 500 mg/L.

On 30 November 2006 (Day 246) additional sodium bicarbonate was added to the influent AGW of the nBA and SRS columns because the pH in the effluent was continually below a pH of 6.5. The optimum pH for dehalorespiring bacteria is 6.8-7.8 (Middledorp et al., 1999 and SiREM unpublished data). To address this it was decided to increase the pH of the influent AGW to a target pH of 8.0 to increase the pH throughout the columns. A volume of 3 mL of a saturated bicarbonate solution was needed to reach this target pH in 2L of AGW. This increased the concentration of bicarbonate in the influent AGW from 363 mg/L to 455 mg/L. The pH was not adjusted in AGW used in the control column.

## Chapter 6

### Column Study Results

#### 6.1 Column Study Results

Prior to the addition of DNAPL to the columns bromide tracer tests were performed for each column. Breakthrough curves were obtained by continuous injection of a sodium bromide solution at a concentration of 91 mg/L at the experimental flow rate of 0.07 mL/min. In order to determine the soil properties of the columns, namely the porosity and dispersivity, modeling of the bromide tracer tests was performed by Kokkinaki (2007) as part of the University of Toronto's contribution to the Sabre project. The model used was the COMPSIM model developed by Sleep and Sykes (1993). The model is a finite difference groundwater fate and transport model. The porosity, dispersivity, velocity and residence time values calculated are summarized in Table 6.1 and the breakthrough curves and raw data from the tracer test is presented in Appendix C.

**Table 6.1** Summary of Column Properties

Column	Porosity	Dispersivity (m)	Velocity (cm/day)	Residence Time (Days)
Control Column	0.31	0.08	7.6	7.97
SRS Amended Column	0.31	0.06	8.2	7.45
nBA Amended Column	0.32	0.099	7.4	8.29

The results indicate that the columns were all packed similarly, had similar properties, and that there were likely no preferential flow paths.



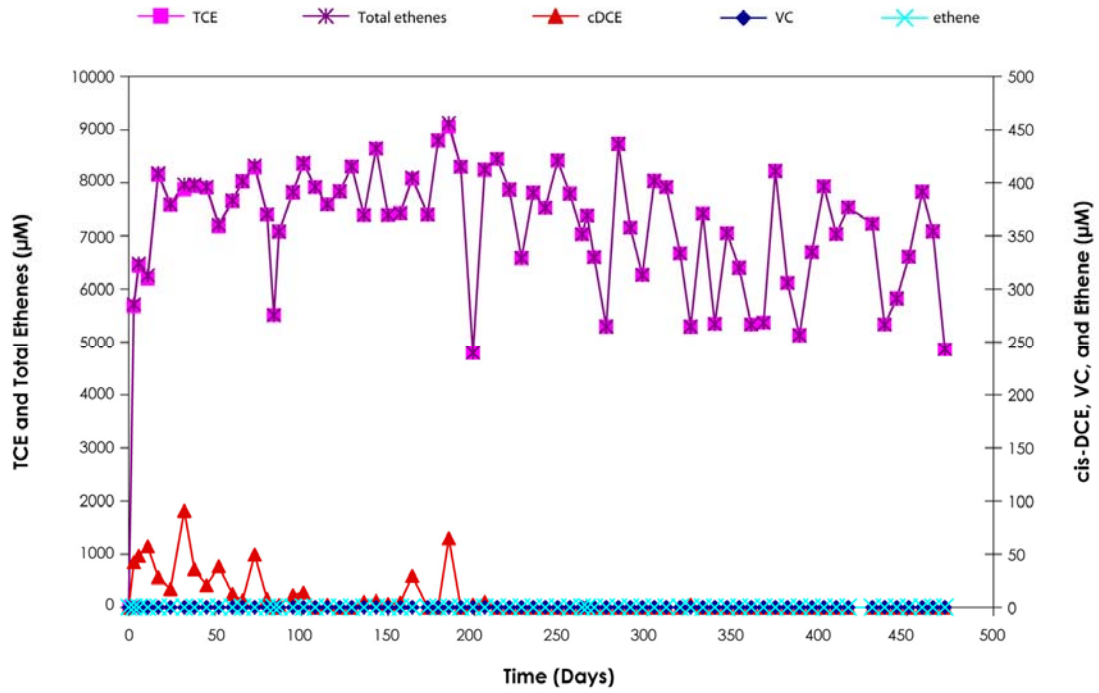
### 6.1.1 Control Column Results

The control column operational period discussed in this report was from 7 July 2006 (Day 0) to 22 October 2007 (Day 472). The VOC and ethene results of the effluent samples collected over the duration of the control column operational period are presented graphically in Figure 6.1. The first effluent samples were collected on day 3. The results showed TCE in the effluent at a concentration of 746 mg/L (5.7 mM) and *cis*-DCE at a concentration of 4.1 mg/L (0.043 mM), VC and ethene were not detected. TCE concentrations increased to day 17 reaching the near saturation concentration of 1,070 mg/L (8.1 mM). TCE concentrations remained relatively stable through day 255 with concentrations in the effluent ranging between 900 mg/L (6.85 mM) and 1,200 mg/L (9.1 mM). *cis*-DCE concentrations remained low up to day 52, after which concentrations decreased to non-detect values. There were low level detections of *cis*-DCE periodically throughout the incubation period, while VC and ethene remained non-detect. After day 255 more variability in the TCE effluent values was observed. Periodic values as low as 610 mg/L (4.6 mM) were measured, but values as high as in the first 255 days of column operation were still consistently observed. The reason for the increased variability is likely due to heterogeneity of the source as the DNAPL phase in the column shrinks due to dissolution. There were no detections of VC or ethene, and methane concentrations remained at or below the detection limit throughout the operation period.

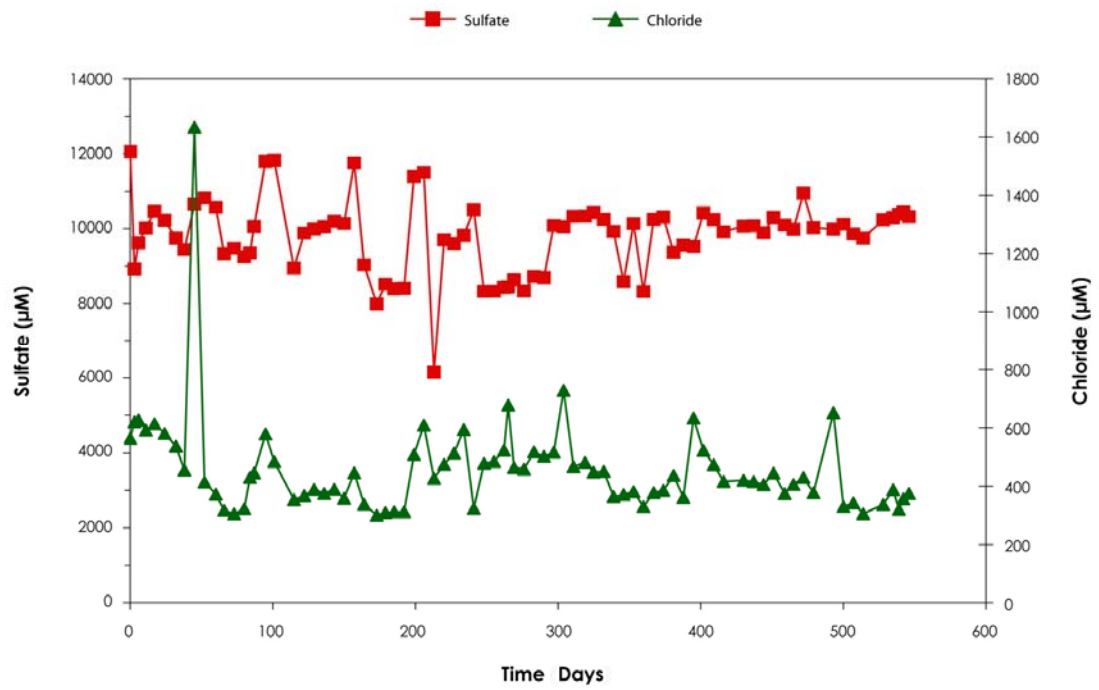
The anion data is presented in Figure 6.2. Sulfate and chloride concentrations remained relatively stable throughout the operation period. Concentrations of VFAs were below the detection limit in the control column throughout the operation period (data not shown).

The pH and ORP data is presented in Figure 6.3. The pH in the control column initially increased to a maximum value of 7.4 on day 115, followed by a slow decrease throughout the remainder of the operational period, but the pH remained above 6.5 over the course of the experiment. The ORP initially slowly decreased to a value of -29 millivolts (mV) on day 213 followed by a rapid increase to 145 mV on day 227. The ORP remained mainly positive throughout the rest of the operational period with periodic decreases below zero observed.

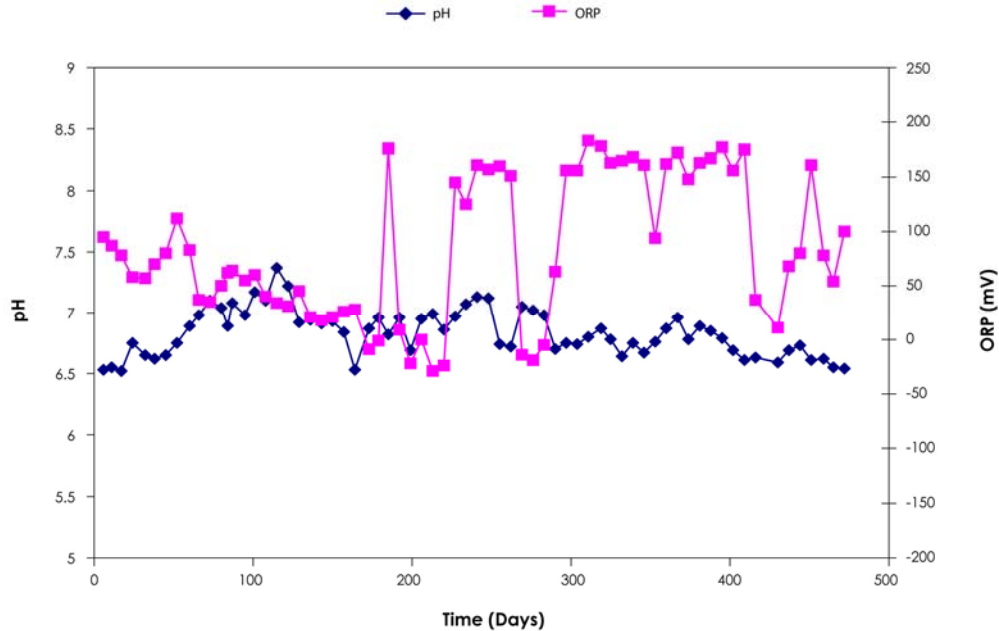
The *Dhc* and *vcrA* concentration in the control column remained at or below the detection limit throughout the column operational period (data not shown).



**Figure 6.1** VOC and ethene concentration trends in the control column



**Figure 6.2** Anion concentration trends in the control column



**Figure 6.3** pH and ORP trends in the control column

### 6.1.2 nBA column results

The nBA column operational period was from 21 March 2006 (Day 0) to 29 January 2007 (Day 306). The VOC and ethene results of the effluent samples collected over the duration of the nBA column operation are graphically presented in Figure 6.4. Prior to the first nBA amendment TCE and *cis*-DCE concentrations reached 148 mg/L (1.1 mM) and 92 mg/L (0.95 mM) respectively. *cis*-DCE concentrations leveled off until day 32 and then decreased to 16 mg/L (0.17 mM) on day 47. TCE concentrations continued to increase to 601 mg/L (4.6 mM) on day 45. After the first nBA amendment on day 47 TCE concentrations continued to increase reaching 842 mg/L (6.4 mM) on day 75. *cis*-DCE concentrations initially decreased slightly followed by an increase to 35 mg/L (0.34 mM) on day 65 and then a decrease to 11 mg/L (0.12 mM) on day 75. After bioaugmentation on day 78 TCE concentrations continued to increase reaching 924 mg/L (7.4 mM) on day 92. TCE concentrations remained relatively stable through the second nBA addition, increasing only slightly to 1,060 mg/L (8.1 mM) prior to the third nBA addition on day 209. *cis*-DCE concentrations remained relatively stable after bioaugmentation until the second nBA addition, after which *cis*-DCE increased to 27 mg/L (0.28 mM) on day 145. *cis*-DCE decreased to 6.0 mg/L (0.062 mM) by day 173 and remained stable prior to the third nBA addition. TCE concentrations remained relatively stable during the 2,000 mg/L nBA addition from day 209 to day 246. However, *cis*-DCE began to increase

slowly to day 222 and then more rapidly to day 252 reaching a concentration of 430 mg/L (4.5 mM). After the 2,000 mg/L nBA addition phase TCE and *cis*-DCE concentrations decreased to 772 mg/L (5.9 mM) and 84 mg/L (0.87 mM) respectively by day 279 prior to the final nBA addition on day 280. Following the neat nBA addition on day 280, TCE continued to decrease slightly and reached 677 mg/L (5.1 mM) by the final sampling event on day 306. *cis*-DCE increased to 105 mg/L (1.1 mM) on day 290 followed by a decrease to 41 mg/L (0.42 mM) on day 306. There were no detections of VC or ethene throughout the column operation period and methane concentrations remained consistently at or below the detection limit.

The anion data for the nBA column is presented in Figure 6.5. Sulfate concentrations increased from day 0 to the first nBA addition on day 46 reaching a concentration of 1,041 mg/L (11 mM). Sulfate concentrations remained relatively stable after bioaugmentation and the second nBA addition prior to the third nBA addition on day 209. Chloride concentrations increased initially to day 13 reaching a concentration of 68 mg/L (1.9 mM), remained stable to day 29 and decreased to 25 mg/L (0.71 mM) on day 47. Chloride concentrations remained relatively stable after bioaugmentation and the second nBA addition prior to the third nBA addition on day 209. Sulfate concentrations remained stable to day 216, following the 2,000 mg/L nBA addition that began on day 209, and then decreased rapidly to 28 mg/L (0.29 mM) on day 249. The sulfate level remained at this low level to day 251. Following the end of the 2,000 mg/L nBA addition phase on day 246 sulfate rapidly increased to influent levels by day 263. Sulfate concentrations decreased slightly to 720 mg/L (7.6 mM) on day 279. Chloride concentrations increased following the 2,000 mg/L nBA addition reaching a maximum concentration of 162 mg/L (4.6 mM) on day 249 followed by a decrease to influent levels by day 280. Following the neat nBA addition on day 280 sulfate concentrations remained relatively stable to day 288 and then increased to influent levels by day 306. Chloride initially increased to 61 mg/L (1.7 mM) on day 289 followed by decrease to 14 mg/L (0.39 mM) on day 292. The chloride concentration was showing an increasing trend on the last sampling event on day 306 reaching 68 mg/L (1.9 mM).

The VFA data for the nBA column is presented in Figure 6.6. The acetate concentrations are displayed on a secondary axis because acetate was the dominant VFA and was present at concentrations much higher than the other VFAs. The time scale for this figure starts at the beginning of the first nBA amendment period (Day 47). All VFAs were below detection limits prior to electron

donor addition. After the nBA addition all VFAs remained below the detection limit to day 56 when a spike of butyrate was observed to reach 9.1 mg/L (0.10 mM), followed by a spike of acetate at 65 mg/L (1.1 mM) on day 58. Butyrate concentrations decreased and remained below the detection limit until the second nBA addition, while acetate concentrations decreased slowly from day 58 to day 120. The second nBA addition occurred from day 118 to 125. Acetate increased rapidly to 432 mg/L (7.3 mM) on day 127, followed by a decrease to non-detect on day 145. Butyrate concentrations also increased after the second nBA addition reaching 14 mg/L (0.16 mM) on day 130, followed by a decrease to below the detection limit on day 145. All VFAs remained near or below the detection limit prior to the 2,000 mg/L nBA addition starting at day 209. Acetate increased rapidly to 1,430 mg/L (24 mM) on day 240 followed by a rapid decrease after the 2,000 mg/L addition phase to below the detection limit on day 279. Butyrate concentrations also increased reaching 484 mg/L (5.5 mM) on day 245, followed by a decrease to below detection on day 264. All VFAs remained near or below detection prior to the neat nBA addition on day 280. Acetate concentrations increased slightly at day 286 and then increased rapidly to 879 mg/L (15 mM) on day 294, followed by a decrease to 193 mg/L (3.3 mM) at the end of the operational period on day 306. Acetate concentrations were showing a declining trend at the end of the incubation period. Butyrate increased only slightly after the neat nBA addition with concentrations of approximately 2.0 mg/L (0.022 mM) persisting in the effluent until the end of the operational period. The other VFAs (lactate, formate, and pyruvate) were below detection limits for the entire column operational period.

The nBA and butanol data is presented in Figure 6.7. There were no detections of nBA in the column effluent during the first three nBA addition phases. Following the neat nBA addition on day 280 nBA increased to 752 mg/L (6.5 mM) on day 294 followed by a decrease to 8.6 mg/L (0.073 mM) at the end of the operational period on day 306. Butanol measurements were not taken until the 2,000 mg/L nBA addition phase began on day 209. Butanol increased to 731 mg/L (9.9 mM) on day 222 followed by a rapid decrease to 97.4 mg/L (1.3 mM) on day 232. Butanol concentrations increased to 1,575 mg/L (21 mM) on day 294 followed by a decrease to 275 mg/L (3.7 mM) at the end of the operational period on day 306.

The pH and ORP data for the nBA column is presented in Figure 6.8. The pH in the nBA column initially decreased to 6.32 on day 78, followed by a slow increase to 7.42 on day 215. The pH then decreased to 6.30 on day 264 despite the increased bicarbonate addition to the AGW on day 236. The

pH increased slightly to 6.60 at the end of the operational period on day 306. ORP measurements were not taken in the nBA column until day 99. The ORP initially increased to a value of 112 on day 152 followed by a decrease throughout the operational period to -27mV on day 306.

The Dhc and vcrA concentrations in the nBA column were at concentrations of  $1 \times 10^6$  gene copies/L and  $4 \times 10^5$  gene copies/L respectively prior to electron donor addition. Both Dhc and vcrA decreased to near non-detect levels prior to bioaugmentation and then remained at this low level throughout the operational period (data not shown).

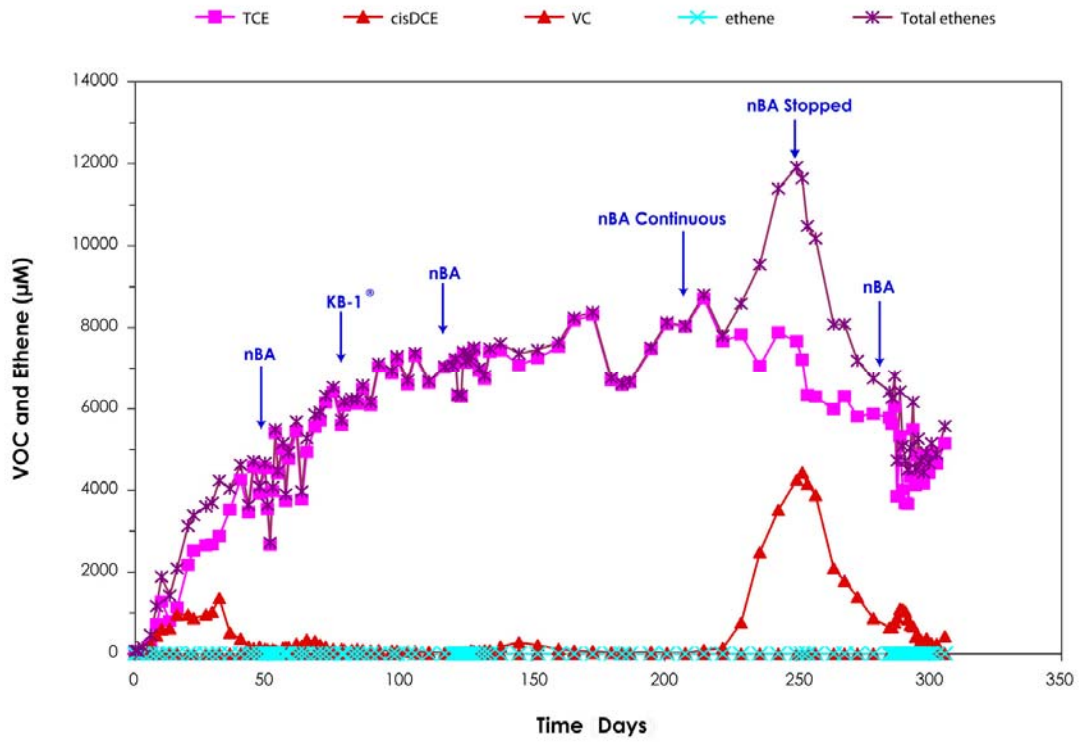


Figure 6.4 VOC and ethene concentration trends in the nBA Amended column

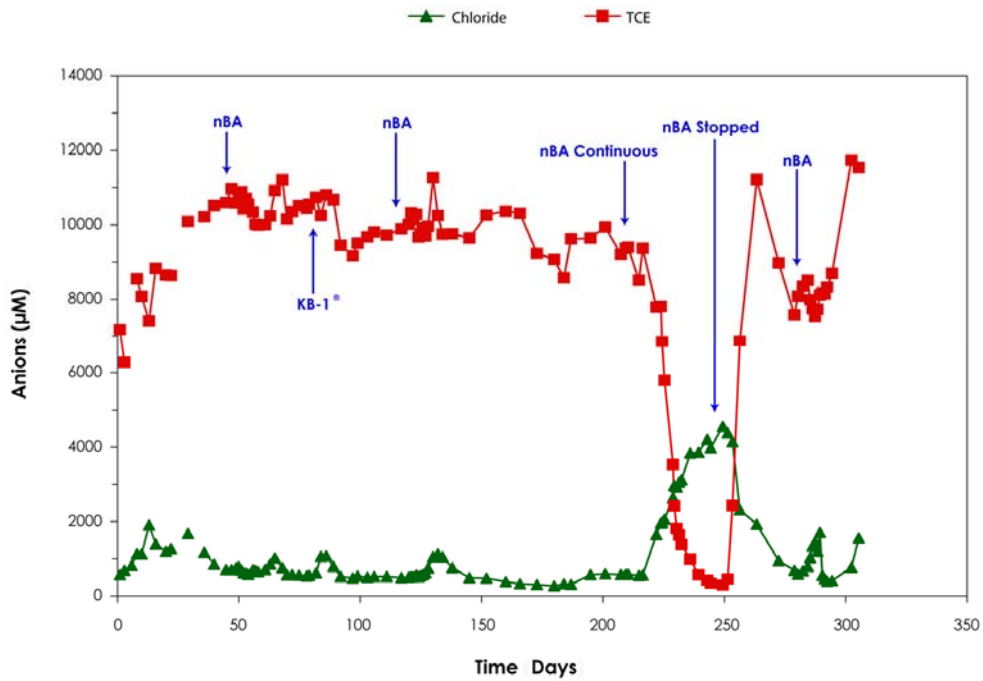
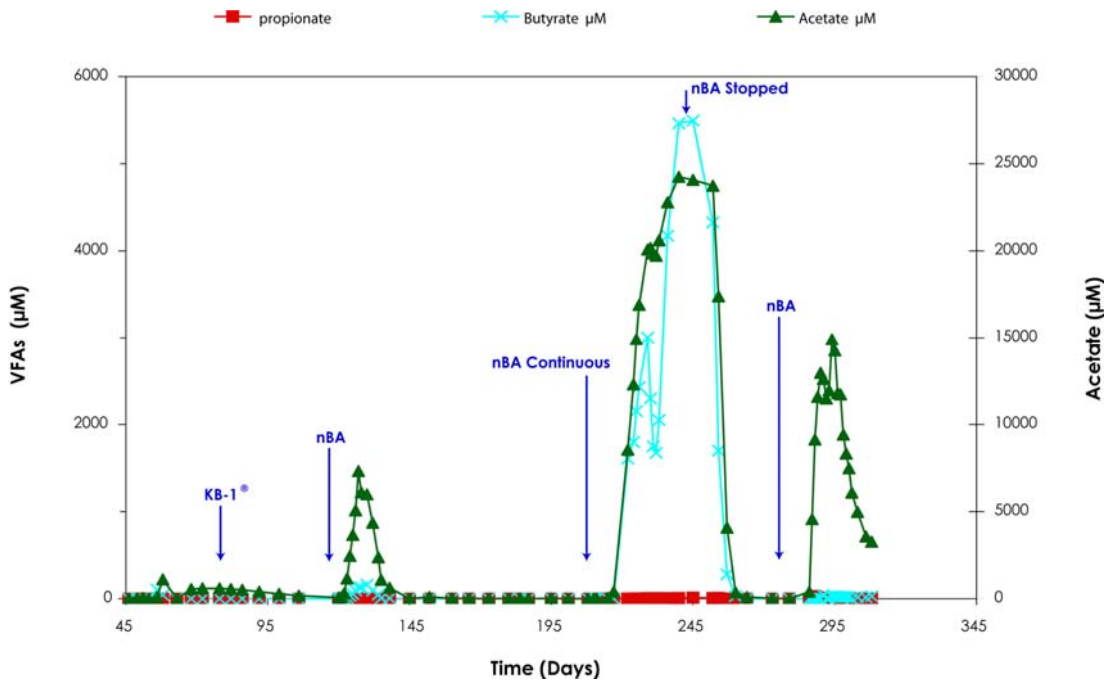
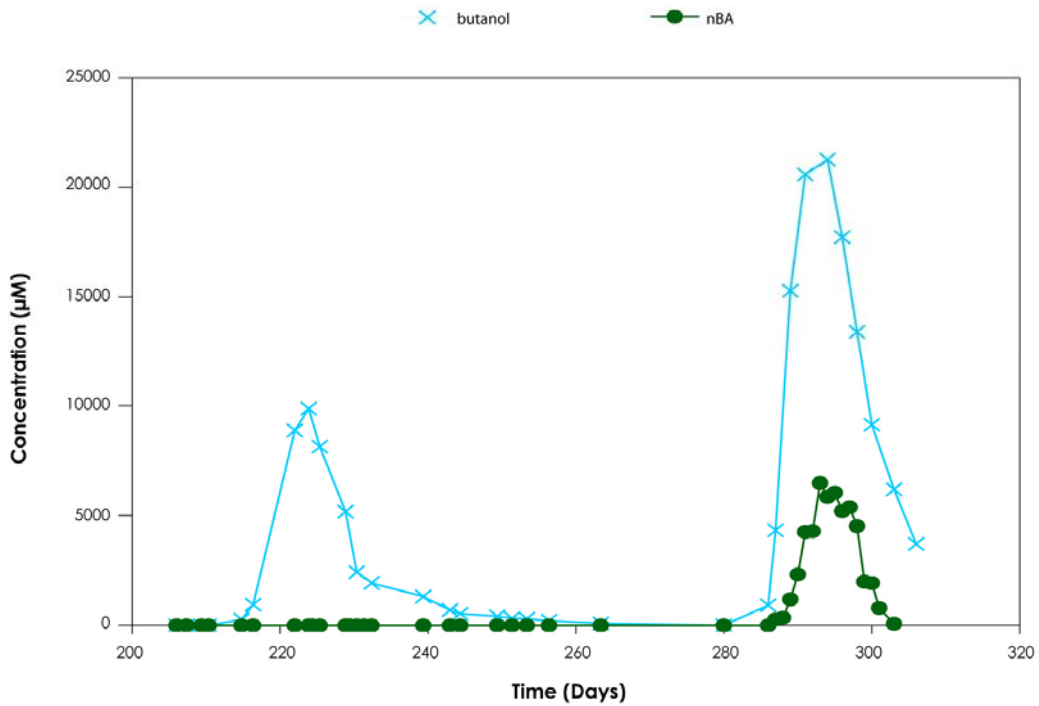


Figure 6.5 Anion concentration trends in the nBA Amended column

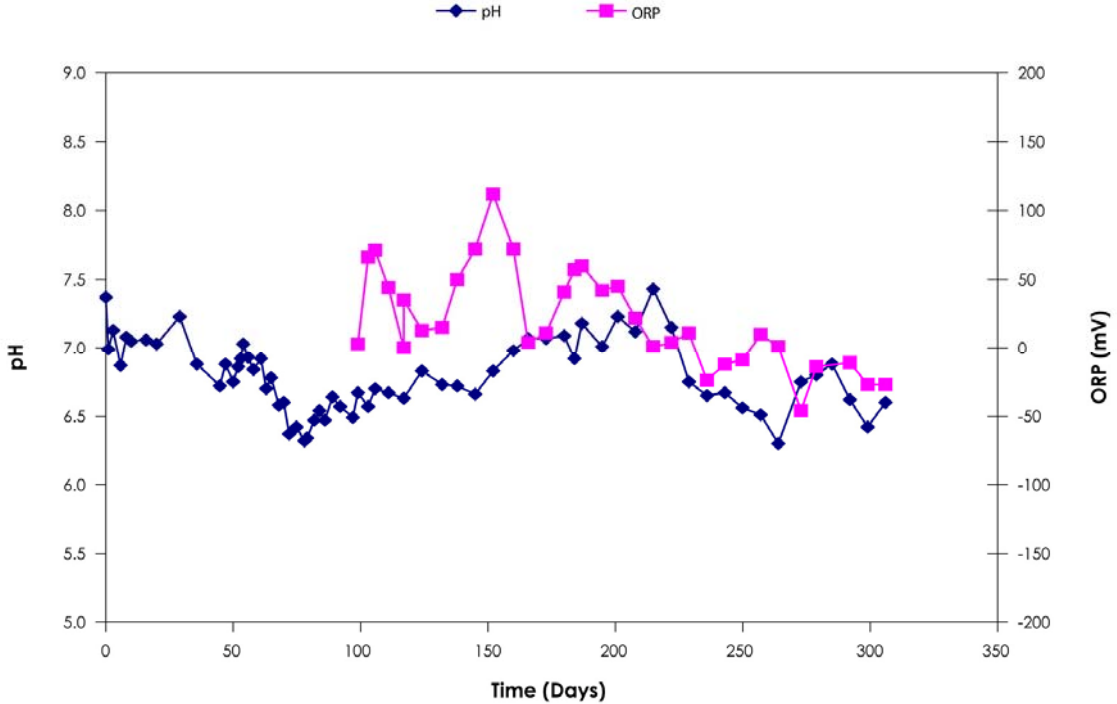


**Figure 6.6** VFA concentration trends in the nBA Amended column



**Figure 6.7** nBA and butanol concentration trends in the nBA Amended column





**Figure 6.8** pH and ORP trends in the nBA amended column

### 6.1.3 SRS™ Column Results

The SRS™ column operational period discussed here was from 21 March 2006 (Day 0) to 22 October 2007 (Day 572). The operational period was divided into two phases; the DNAPL phase from day 0 to day 329, and the plume phase from day 329 to 572. The VOC and DHG results of the effluent samples collected over the duration of the SRS™ column operational period are presented in Figure 6.9. Prior to SRS™ amendment TCE and *cis*-DCE concentrations increased similarly to the trend observed in the nBA column. On day 16 TCE reached a concentration of 268 mg/L (2.0 mM) and *cis*-DCE reached 92 mg/L (0.95 mM). *cis*-DCE concentrations decreased after day 16 to 3.4 mg/L (0.035 mM) on day 47. TCE concentrations continued to increase to 785 mg/L (6.0 mM) on day 45. After the SRS™ amendment on day 47 TCE concentrations continued to increase reaching 913 mg/L (6.9 mM) on day 78. *cis*-DCE concentrations increased slightly to 14 mg/L (0.14 mM) on day 72 followed by a decrease to 4.2 mg/L (0.043 mM) on day 78. After bioaugmentation on day 78 TCE concentrations continued to increase reaching 1,068 mg/L (8.1 mM) on day 89 followed by a slow decrease to a concentration of 549 mg/L (4.2 mM) on day 229 and a rapid decrease to 4.8 mg/L (0.036 mM) on day 257. TCE concentrations remained at or near the detection limit for the

remainder of the DNAPL phase. Following bioaugmentation on day 78 *cis*-DCE concentrations remained relatively stable to day 89 followed by an increase to 131 mg/L (1.3 mM) on day 145. *cis*-DCE concentrations increased slowly to a concentration of 154 mg/L (1.6 mM) on day 195 followed by a rapid increase to a concentration of 498 mg/L (5.1 mM) on day 243. The *cis*-DCE concentration decreased rapidly to near the detection limit on day 299 and remained near or below the detection limit for the remainder of the DNAPL phase. VC and ethene concentrations remained below detection limits for the first 138 days of the operational period. On day 145 low levels of VC and ethene were detected and slowly increased to day 243 to concentrations of 0.81 mg/L (0.013 mM) and 0.32 mg/L (0.011 mM) respectively. After day 243 VC concentrations increased corresponding to the decreasing concentration of *cis*-DCE reaching a maximum concentration of 24 mg/L (0.039 mM) on day 268, followed by a rapid decrease to below the detection limit on day 299. VC remained near or below the detection limit for the remainder of the DNAPL phase. Ethene concentrations remained relatively stable from day 243 to day 264 followed by an increase to 1.0 mg/L (0.036 mM) on day 306 corresponding to the decreasing concentration of VC. Ethene decreased slightly to 0.65 mg/L (0.023 mM) on day 327. Methane concentrations remained stable up to day 152, after which methane began to increase reaching concentration of 8.6 mg/L (0.54 mM) on day 320.

From day 306 to day 327 the main component measured in the column effluent was ethene. Since the DNAPL was depleted it was decided to amend the influent groundwater with high concentrations of TCE to simulate a TCE plume area down gradient of a TCE source zone. On day 329 AGW spiked with approximately 400 mg/L of TCE was pumped through the SRS<sup>TM</sup> column for the remainder of the operational period. Ethene began to increase almost immediately followed by detections of VC and *cis*-DCE on day 362. Ethene continued to increase to a maximum concentration of 35 mg/L (1.2mM) on day 509 followed by a decrease to 16 mg/L (0.56 mM) on day 572. VC increased rapidly from day 362 to day 390 reaching a concentration of 62 mg/L (0.99 mM). VC remained relatively stable to day 495 followed by a decrease to a concentration of 32 mg/L (0.52 mM) at the end of the operational period on day 572. *cis*-DCE increased from day 362 to day 404 reaching a concentration of 139 mg/L (1.4 mM). *cis*-DCE decreased to a concentration of 22 mg/L (0.23 mM) on day 460 and remained relatively stable to day 572. TCE concentrations remained below the detection limit over the plume phase of the study. Methane concentrations increased following the end of the DNAPL phase to a concentration of 12 mg/L (0.74 mM) on day 365 followed by a rapid

decrease to 1.8 mg/L (0.12 mM) on day 404 and remained relatively stable for the remainder of the plume phase.

The target TCE concentration in the influent AGW was 400 mg/L. Losses of TCE were observed in the influent AGW bag so the concentration of TCE added to the column varied with time. Figure 6.10 shows the fluctuations of TCE in the influent bag over time during the plume phase.

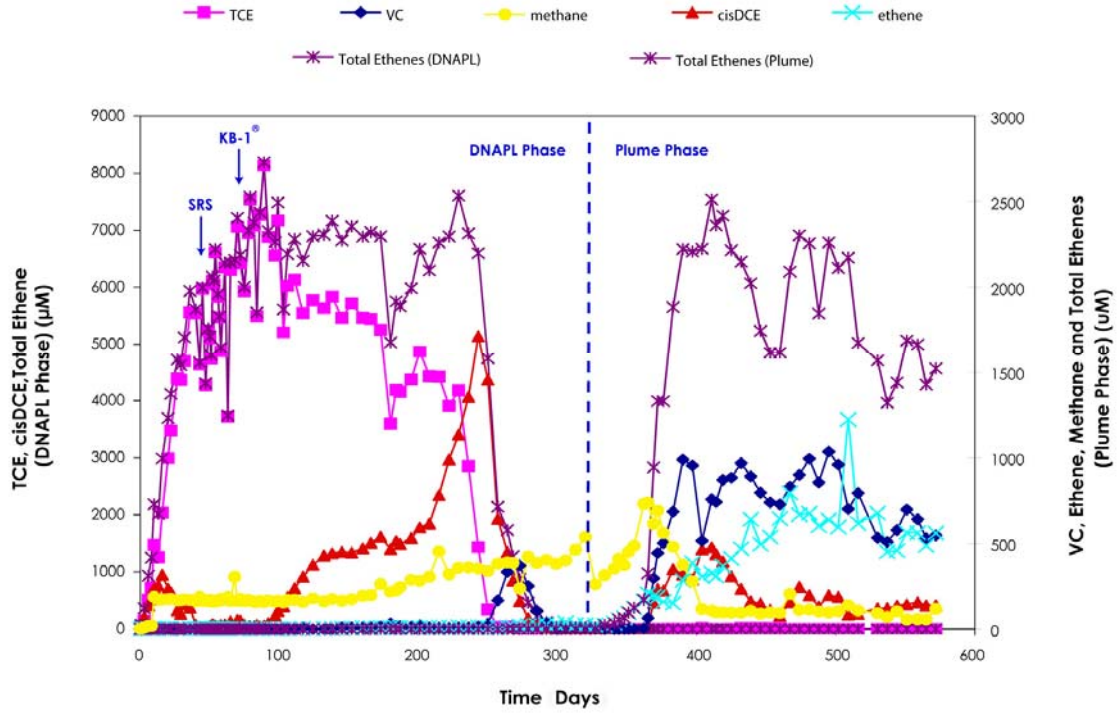
The anion data for the SRS<sup>TM</sup> column is presented in Figure 6.11. Sulfate concentrations increased from day 0 to day 47 reaching a concentration of 1,015 mg/L (11 mM). Sulfate concentrations decreased slightly after SRS addition to a concentration of 992 mg/L (10 mM) on day 79. Chloride concentrations increased initially to day 16 reaching a concentration of 48 mg/L (1.4 mM) followed by a slow decrease to 20 mg/L (0.56 mM) on day 79. After bioaugmentation on day 79 sulfate concentrations remained stable to day 86 followed by a rapid decrease to 31 mg/L (0.33 mM) on day 173. A rebound of sulfate to 242 mg/L (2.5 mM) on day 295 was observed followed by a decrease to 1.8 mg/L (0.019 mM) on day 264. The concentration of sulfate remained low at the start of the plume phase until day 348, after which the sulfate concentration began to increase reaching a maximum concentration of 562 mg/L (5.9 mM) on day 404. After day 404 sulfate decreased to 244 mg/L (2.6 mM) on day 460 and then slowly increased to the end of the operational period reaching a concentration of 514 mg/L (5.4 mM) on day 572. The concentration of chloride began to increase after bioaugmentation reaching 72 mg/L (2.0 mM) on day 103 and remained relatively stable to day 195. Chloride increased to a maximum concentration of 193 mg/L (5.4 mM) on day 243 followed by a decrease to 16 mg/L (0.46 mM) by day 285. The concentration of chloride remained stable to day 327 and began to increase at the start of the plume phase reaching a maximum concentration of 214 mg/L (6.0 mM) on day 488. Chloride showed a decreasing trend to the end of the operational period reaching a concentration of 112 mg/L (3.2 mM) on day 572.

The VFA data for the SRS<sup>TM</sup> column is presented in Figure 6.12. The acetate concentrations are displayed on a secondary axis because acetate was the dominant VFA measured and was present at concentrations much higher than the other VFAs. The time scale for this figure starts at the beginning of the SRS<sup>TM</sup> amendment period (Day 47). All VFAs were below detection limits prior to electron donor addition. After the SRS<sup>TM</sup> addition increases in lactate, butyrate, propionate, and acetate were observed. Lactate increased first reaching a maximum concentration of 363 mg/L (4.1 mM) by day

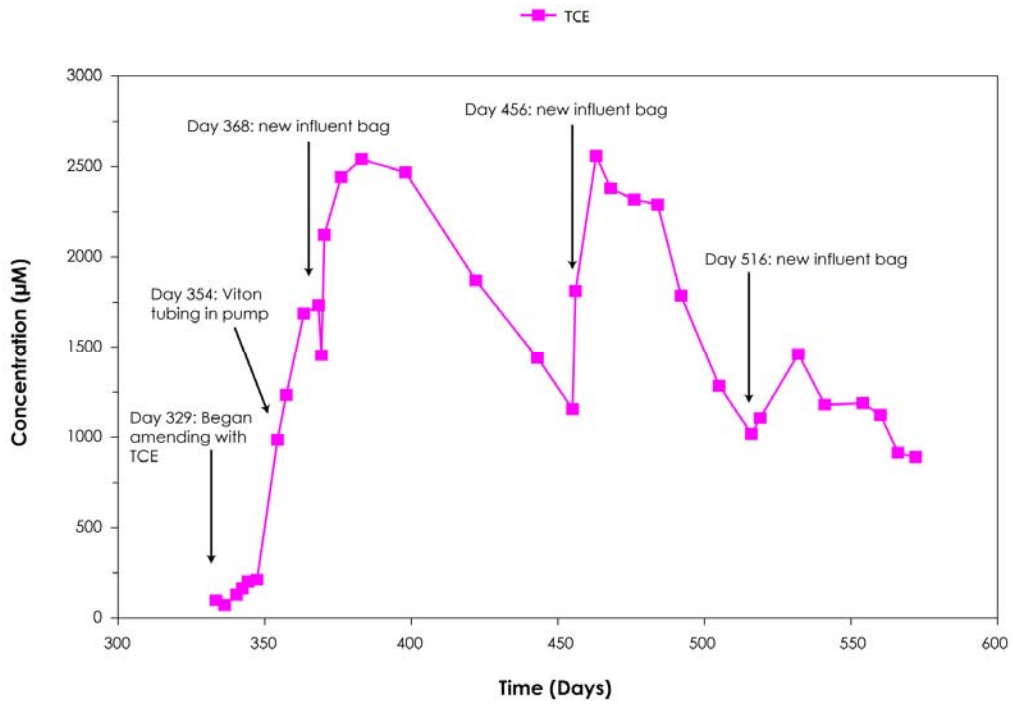
56 followed by a decrease to below the detection limit at day 78. Butyrate and propionate increases followed reaching maximum concentrations on day 58 of 421 mg/L (4.8 mM) and 153 mg/L (2.1 mM) respectively. Butyrate decreased to below detection limit on day 78, while propionate decreased to 41.5 mg/L (0.57 mM) on day 63 and remained relatively stable through to day 78. Acetate increased reaching 133 mg/L (2.3 mM) on day 58 followed by a decrease to near the detection limit on day 63. Low levels of acetate persisted through to day 78. After bioaugmentation on day 78 the concentration of propionate decreased to 12 mg/L (0.16 mM) on day 124 and remained relatively stable through day 292 followed by a decrease to below the detection limit on day 327. Propionate remained near or below the detection limit for the remainder of the operational period. Butyrate was detected sporadically between days 78 and 327 and remained below the detection limit to the end of the operational period. Acetate concentrations increased rapidly after bioaugmentation reaching a maximum concentration of 1,623 mg/L (28 mM) on day 184, followed by a decrease to non-detect on day 404. Acetate remained below detection limit to the end of the operational period on day 572.

The pH and ORP data is presented in Figure 6.13. The pH in the SRS<sup>TM</sup> column remained stable up to day 58 after which it began to decrease reaching a value of 6.30 on day 236. Following the bicarbonate addition, which began on day 236, the pH increased rapidly to 6.77 on day 273 followed by a slow increase to 7.28 on day 467. The pH decreased slightly to 6.81 on day 572. ORP measurements were not taken in the SRS column until day 99. The initial ORP was slightly less than zero and initially increased to +50 mV on day 111, followed by a decrease reaching a value of -281 mV on day 432. The ORP remained stable to day 502, after which it increased to a value of 161 mV on day 551, followed by a decrease to -9 mV at the end of the operational period.

The molecular data is presented in Figure 6.14. The Dhc and vcrA concentrations in the SRS column were at concentrations of  $1 \times 10^5$  gene copies/L and  $2 \times 10^5$  gene copies/L respectively prior to electron donor addition. These concentrations were associated with the detection limit of the method. Both Dhc and vcrA remained at the detection limit up to day 230, after which a rapid increase in both values was observed reaching  $1 \times 10^8$  and  $8 \times 10^7$  gene copies/L respectively by day 324. The concentration of Dhc and vcrA remained at these high levels for the remainder of the column operational period.



**Figure 6.9** VOC, Ethene and Methane concentration trends in the SRS amended column



**Figure 6.10** TCE concentration trends in the influent AGW bag

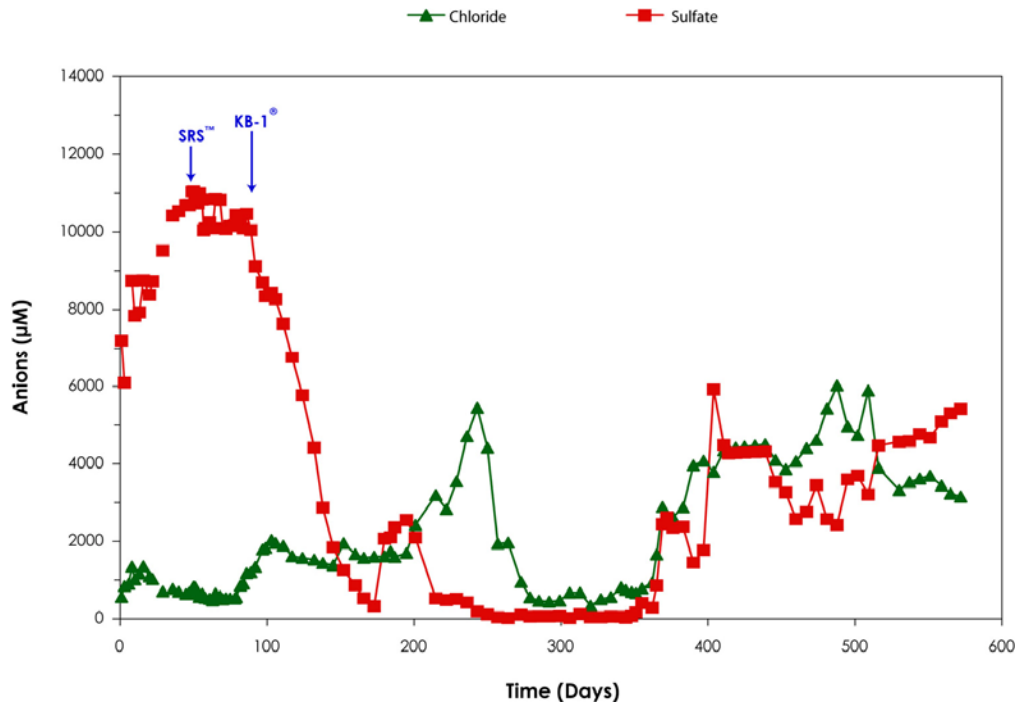


Figure 6.11 Anion concentration trends in the SRS amended column

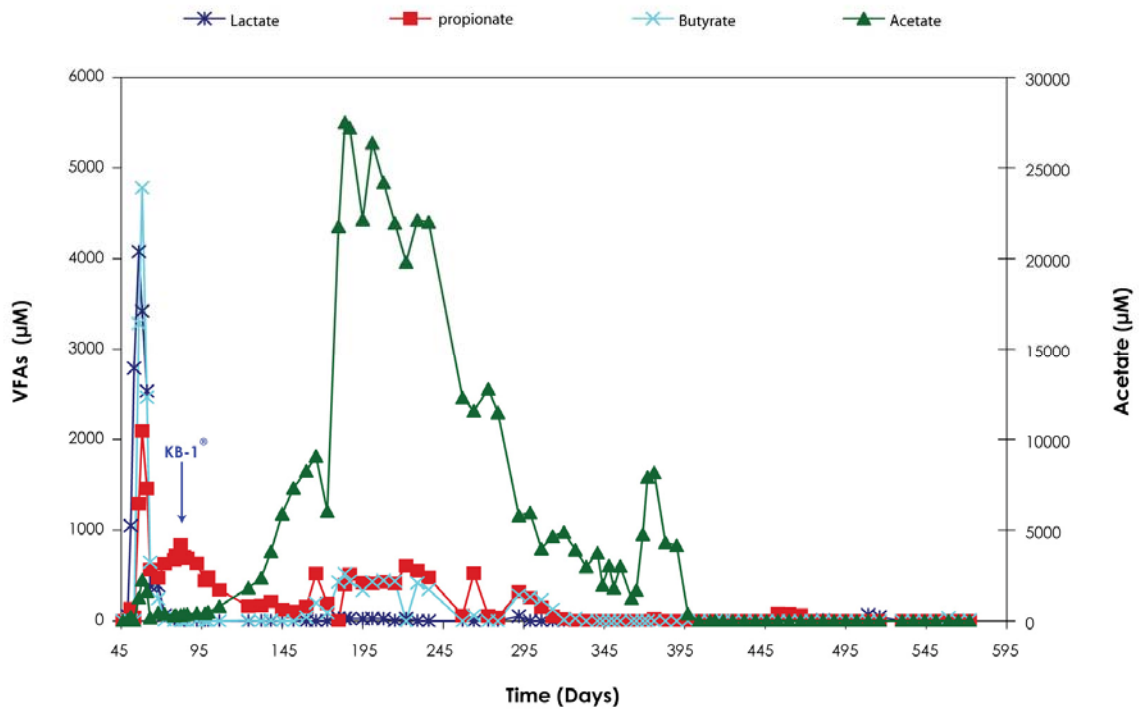
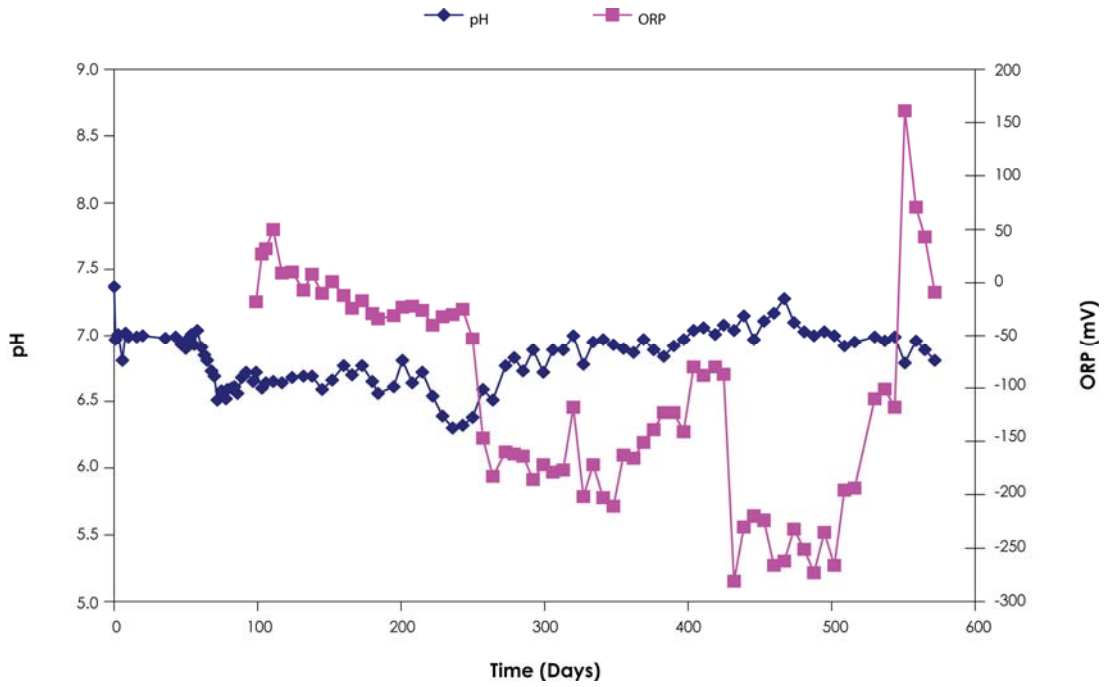
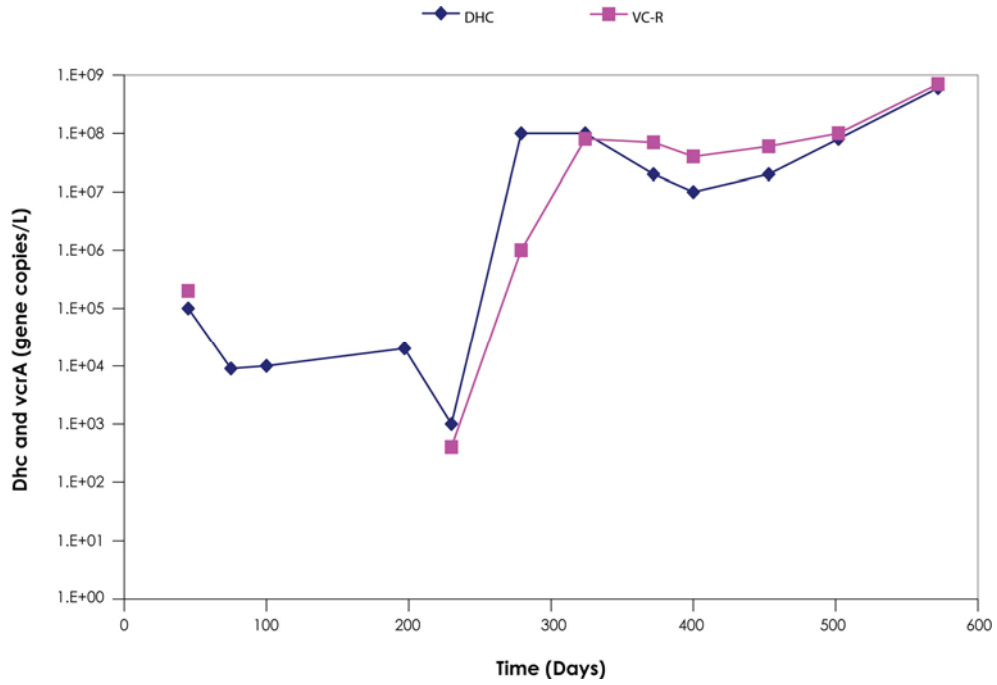


Figure 6.12 VFA concentration trends in the SRS amended column



**Figure 6.13** pH and ORP trends in the SRS amended column



**Figure 6.14** Dhc and vcrA concentration trends in the SRS<sup>TM</sup> amended column

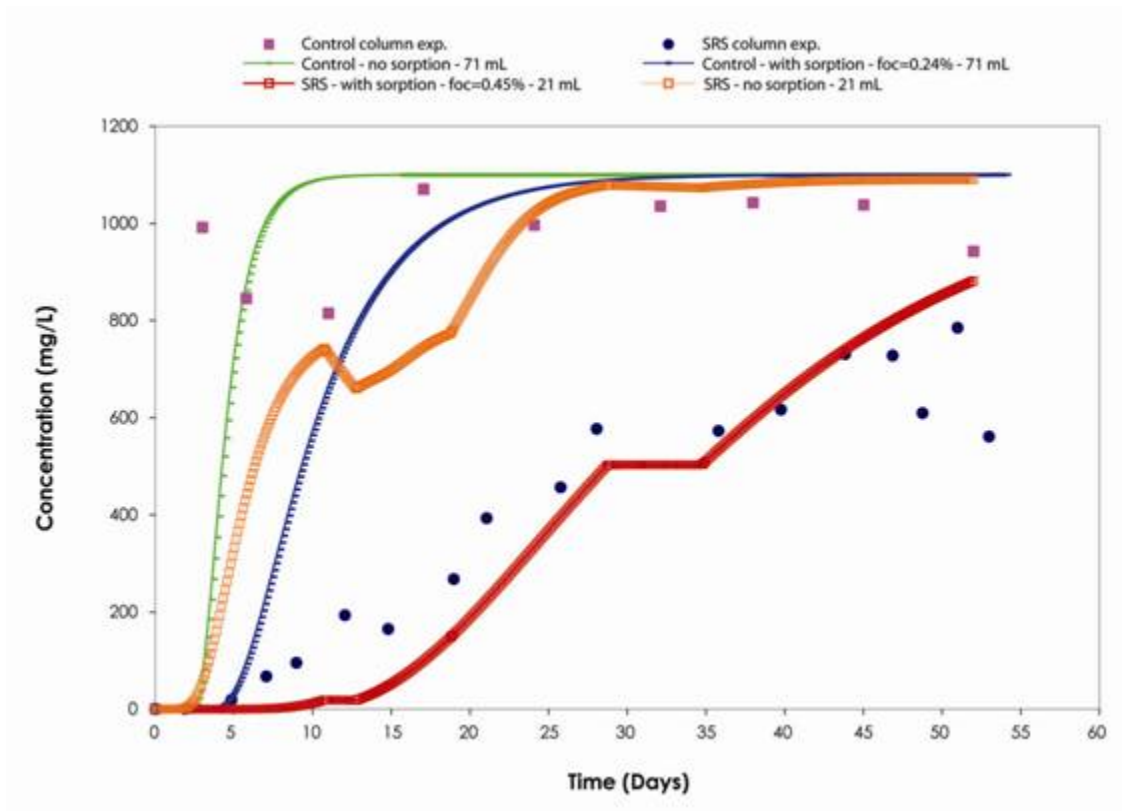
## Chapter 7

### Column Study Discussion

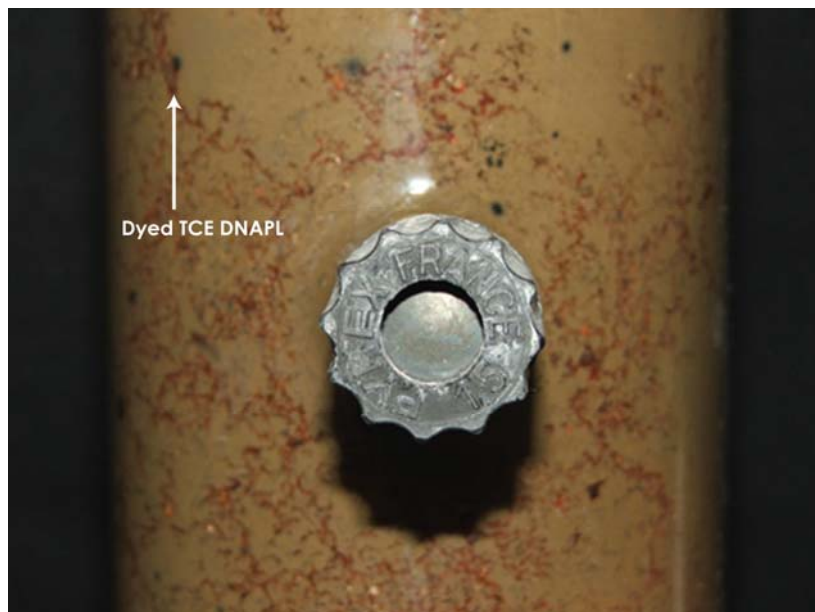
#### 7.1 TCE DNAPL Distributions

The effluent TCE concentration profile for the control column over the first 47 days of column operation was different than the nBA and SRS<sup>TM</sup> columns. TCE concentrations increased very rapidly in the control column to near saturation levels, whereas the increase in TCE was much more gradual in the nBA and SRS<sup>TM</sup> amended columns. The COMPSIM fate and transport model was used to simulate the TCE breakthrough curves for the SRS and control columns. The results are presented in Figure 7.1. The results indicate that the reason for the different breakthrough profiles was primarily due to the greater amount of TCE retained in the control column compared to the treatment columns. This is evident by comparing the curves for each column that do not account for sorption effects because the amount of TCE retained is the only variable for those simulations. The simulated concentration of TCE increased much faster in the control column than in the SRS<sup>TM</sup> column as was observed from the experimental data. Another factor that appears to add to the differences in TCE dissolution profiles is the sorption properties of the columns. The best fit for the experimental data from the control column was when sorption was not considered, while the best fit for the SRS<sup>TM</sup> column was when a fraction of organic carbon ( $F_{oc}$ ) of 0.45% was used. The aquifer material was homogenized prior to packing the columns, but it is difficult to completely homogenize soil and it is possible that the SRS<sup>TM</sup> and nBA columns received a higher percentage of organic matter than the control column. Also, the control column was constructed several months after the treatment columns and it is possible that the organic matter in the material was degraded or changed during storage. A picture of the TCE distribution in the control column is shown in Figure 7.1. The TCE distributions in the nBA and SRS<sup>TN</sup> columns looked similar to the control column.





**Figure 7.1** Experimental and Simulation Results for the TCE Breakthrough curves in the Control and SRS™ Columns



**Figure 7.2** Dyed TCE distribution in the Control Column

## 7.2 Effect of TCE Concentrations on Indigenous Microorganisms

The *cis*-DCE concentration profiles increased similarly in the nBA and SRS<sup>TM</sup> amended columns, while *cis*-DCE was not observed to increase in the control column. The initial increase in *cis*-DCE in the nBA and SRS<sup>TM</sup> amended columns is likely from dechlorination of TCE to *cis*-DCE by indigenous microorganisms present in the aquifer material. Partial dechlorination of TCE to *cis*-DCE was observed in intrinsic control microcosms in the Sabre microcosm study indicating that the aquifer material did indeed contain dechlorinating organisms. The  $F_{oc}$  of the aquifer material was measured and found to be 0.6%, which may have served as electron donor for the dechlorination of TCE to *cis*-DCE. Once the TCE concentrations reached approximately 400 mg/L (3.0 mM) in the effluent of the nBA and SRS<sup>TM</sup> columns the concentration of *cis*-DCE began to decrease. This decrease in the dechlorination of TCE to *cis*-DCE could be due to: (1) the  $F_{oc}$  in the aquifer material that was providing electron donor was depleted, or (2) the concentration of TCE in the areas of the column where dechlorination was occurring became toxic to the indigenous microorganisms, and 3) nutrient levels had declined, slowing the biotransformation. The lack of *cis*-DCE production in the control column is likely due to inhibition of TCE to *cis*-DCE dechlorination by the initial high TCE concentrations; therefore the second explanation above is most likely the reason the *cis*-DCE concentrations declined in the treatment columns. This is also supported by previous reports that concentrations of TCE in this range are toxic to dechlorinators (Yang and McCarty, 2000).

## 7.3 Electron donor retention in columns

### 7.3.1 SRS<sup>TM</sup> retention in column

The amount of SRS<sup>TM</sup> added was based on the amount required to support sulfate reduction and complete dechlorination of TCE to ethene for one year of column operation. In order for the one time addition of SRS<sup>TM</sup> to support dechlorination over that period the SRS<sup>TM</sup> would need to be retained in the column. Visual observations were made during the SRS<sup>TM</sup> addition phase. The SRS<sup>TM</sup> material entered the column and followed similar pathways on the outside of the column as the TCE. The SRS<sup>TM</sup> appeared to surround the TCE DNAPL and the two phases appeared to mix as evident by the SRS<sup>TM</sup> turning a pinkish color. The SRS<sup>TM</sup> material was observed to move through the entire column by the end of the addition phase with a slight amount of visible SRS<sup>TM</sup> in the effluent collection vial on 19 May 2006 (Day 54). There was no flow through the SRS<sup>TM</sup> column for three days following the SRS<sup>TM</sup> addition and once the AGW flow continued visible SRS<sup>TM</sup> was not detected in the column

effluent for the remainder of the operational period. SRS<sup>TM</sup> breakdown products (VFAs) were detected in the effluent for the first 400 days of column operation indicating that SRS<sup>TM</sup> was retained in the column for at least that time frame.

GE was able to analyze for TOC and used TOC measurements as a surrogate for SRS<sup>TM</sup> measurements. The GE experiment consisted of two columns, one that contained TCE DNAPL and one that did not, and SRS<sup>TM</sup> was added in equal amounts to each column. They calculated the amount of TOC added to each column and compared it to the amount of TOC measured in the effluent of each column (Table 7.1). Almost all of the TOC added was retained in TCE DNAPL column versus 71% in the control column indicating that SRS partitioned into the TCE DNAPL resulting in enhanced retention of electron donor in the system.

**Table 7.1** Summary of TOC Retention in the GE Control and DNAPL Columns

Column	TOC In (g)	TOC Out (g)	Amount Retained (%)
Control	6.34	1.85	70.82
DNAPL	6.18	0.09	98.54

### 7.3.1.1 SRS<sup>TM</sup> mass balance

A mass balance was performed in an attempt to account for the fate of the SRS<sup>TM</sup> that was added to the SRS<sup>TM</sup> column. Because there was no direct measure of SRS<sup>TM</sup> in the effluent, the mass balance performed used effluent parameters that would consume electron donor and effluent acetate concentrations. For this mass balance the useable electron donor in SRS<sup>TM</sup> was assumed to be linoleic acid, which was 60% of the initial SRS<sup>TM</sup> mass (because SRS<sup>TM</sup> is 60 % soybean oil). First, the amount of H<sub>2</sub> consumed during dechlorination, sulfate reduction, and methanogenesis was calculated (H<sub>2</sub> consumed per electron acceptor). Then the amount of SRS<sup>TM</sup> consumed was estimated from acetate production because linoleic acid is fermented to acetate and hydrogen (equation 2.5) and acetate was easily measured in the column effluent. The amount of H<sub>2</sub> consumed during dechlorination, sulfate reduction and methanogenesis was greater than the amount of H<sub>2</sub> produced from linoleic acid fermentation to acetate (equation 2.5). Therefore the missing H<sub>2</sub> (H<sub>2</sub> required) was assumed to come from the complete oxidation of a portion of the linoleic acid (equation 2.7). The results are summarized in Table 7.2 and the calculations are provided in Appendix D.

The calculations were performed for two time periods. The first was from the day of SRS<sup>TM</sup> addition (Day 47) to day 404. This time period was chosen because after day 404 measurable VFA's (mainly acetate) were no longer detected in the effluent and it is likely that the SRS<sup>TM</sup> was completely consumed at this time. Sulfate reduction and dechlorination were observed to occur after day 404 to the end of the operational period on day 572, but it is possible that the electron donor used during this phase was not from the SRS<sup>TM</sup>, but was from decaying biomass. However, the calculations were also performed for the entire operational phase because it is possible that there were pockets of SRS<sup>TM</sup> remaining in the column that were slowly fermenting providing the needed electrons for sulfate reduction, dechlorination, and methanogenesis. The results indicate that 58 % of the SRS<sup>TM</sup> was consumed by day 404 and 70 % by day 572. In either case the amount of SRS<sup>TM</sup> added to the column was not accounted for in the mass balance. There are several factors that may contribute to the missing mass in the mass balance calculations: (1) there was some visible oil observed in the effluent vial on the last day of SRS<sup>TM</sup> amendment that was not quantified, (2) acetate can be used directly by sulfate reducers (Widdel, 1988) biasing the estimate of SRS consumed from acetate production to be lower than actual (3) other possible H<sub>2</sub> consuming reactions, such as iron and manganese reduction, that were not accounted for, (4) H<sub>2</sub> produced from SRS<sup>TM</sup> fermentation may have not been used and left the column directly and therefore would not have been used in the mass balance calculation, and (5) possible incorporation of carbon into biomass as suggested by Cope and Hughes, (2001). Long and Borden (2006) performed carbon mass balance equations to account for EVO that was amended to columns and obtained errors ranging from 2% to 67% indicating that these measurements are difficult and variability is high.

Even though the total SRS<sup>TM</sup> added to the column was not accounted for, the addition of SRS<sup>TM</sup> was able to promote dechlorination and sulfate reduction over the operational period (572 days) and possibly longer. Long and Borden (2006) reported good performance with no evidence of substrate depletion or reduced degradation rates after 14 months of column operation and calculated that all of the injected carbon would be released from the column in 6 to 7 years. In this case the SRS<sup>TM</sup> was likely depleted from the column after approximately 404 days (last measurable acetate in effluent), but dechlorinating activity and sulfate reduction continued indicating that SRS<sup>TM</sup> amendment was able to promote the appropriate conditions for these processes longer than anticipated.

**Table 7.2** Summary of SRS™ Mass Balance Results

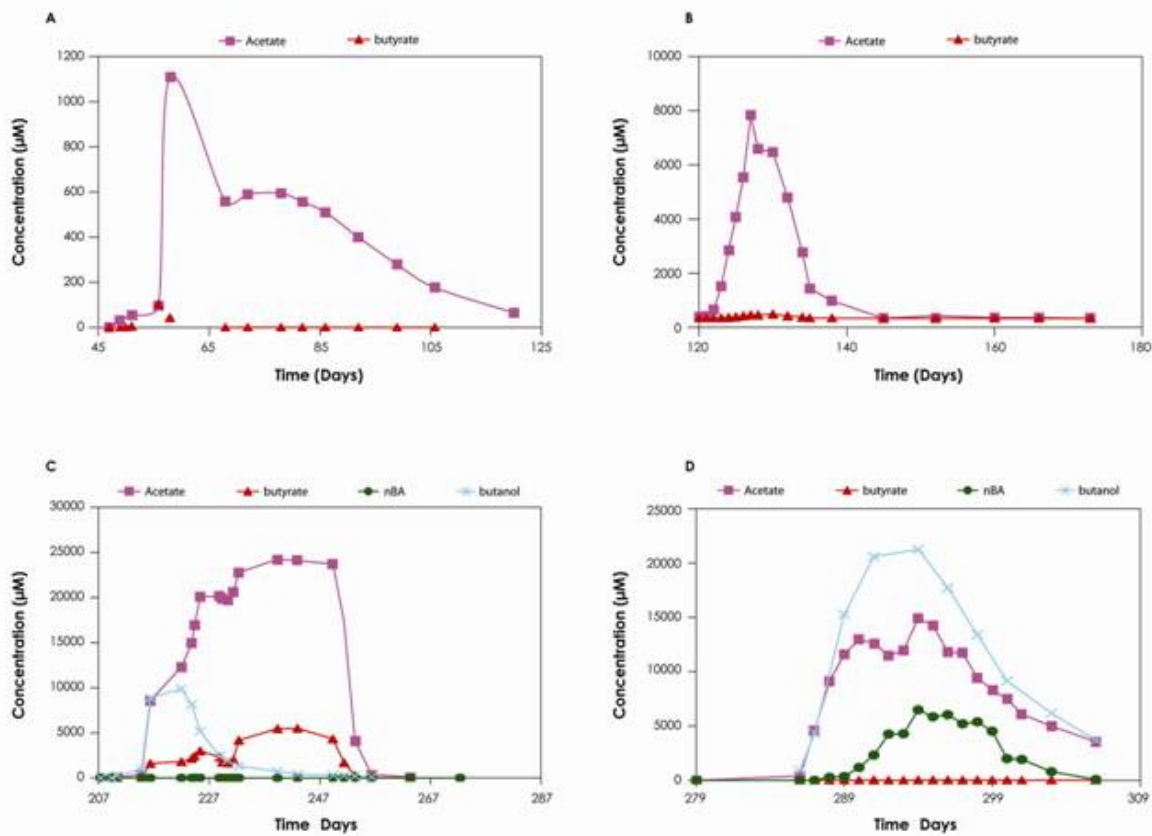
Phase	H <sub>2</sub> consumed per Electron Acceptor (mmoles)						Acetate Produced (mmoles)
	TCE	<i>cis</i> -DCE	VC	methane	sulfate	Sum	
Day 47 - 404	33	2.3	0.59	49	1215	1300	330
Day 47 - 572	53	18	7.5	59	1687	1824	330

linoleic acid consumed to acetate and H <sub>2</sub> (mmoles)	H <sub>2</sub> produced (mmoles)	H <sub>2</sub> required (mmoles)	linoleic acid consumed to CO <sub>2</sub> and H <sub>2</sub> (mmoles)	Total linoleic consumed (mmoles)	Total linoleic consumed (g)	Mass SRS added (g)*	% SRS™ Consumed
37	513	787	16	52	15	25	58
37	513	1311	26	63	18	25	70

Notes: \* corrected for mass of soybean oil in SRS™. SRS™ is 60 % soybean oil

### 7.3.2 nBA Retention in Column and Mass Balance

nBA was added to the nBA column four times throughout the operational period. nBA was not detected in the effluent until the final nBA addition, when it was added as a neat compound. It should be noted that butanol measurements were not taken during the saturated nBA additions, but were taken during the 2,000 mg/L and neat nBA addition phases. nBA is a colorless liquid and visual observations were not possible as with the SRS™ amendment. Mass balance calculations were performed by comparing the amount of nBA added to the column with the amount of nBA and breakdown products measured in the effluent for each amendment phase. The results are summarized in Table 7.3 and the data is presented graphically in Figure 7.3.



**Figure 7.3** Effluent concentration trends of nBA, butanol, butyrate, and acetate. (A) first saturated stock addition. (B) second saturated stock addition. (C) 2,000 mg/L addition phase. (D) neat nBA addition.

**Table 7.3** Summary of nBA Amendments and Mass Recoveries

Days added	Concentration (mg/L)	nBA Mass Added (g)	nBA and Breakdown Products Mass Recovered (g)	Amount Recovered (%)
45-53	6,000	5.4	0.39	7.2*
118-125	6,000	4.9	0.89	18.1*
207-250	2,000	10.9	11.4	104
280	neat	5.3	5.7	108

Notes: \* does not include n-butanol

The recoveries from the first two nBA additions were very poor. Butanol data was not available, but since one mole of butanol is produced for every one mole of acetate during nBA hydrolysis (equation 2.1), and there was very little butyrate measured, the butanol concentrations can be estimated from

the acetate data (assuming that acetate is not further transformed). Using estimated butanol concentrations the nBA recovery increases to 14.2% and 35.9% for the first and second additions respectively. The recoveries are still poor when butanol is included. Initially it was believed that the nBA had partitioned into the TCE DNAPL and the poor recoveries were due to the nBA mass being retained in the DNAPL as was expected to occur with a PED. However, the mass was not recovered over time as nBA or breakdown products, and with the lack of dechlorination and sulfate reduction the fate of nBA is unclear for the first two additions. It is possible that a portion of the nBA was completely oxidized to CO<sub>2</sub>, which was not measured in the effluent, or there were losses of nBA from the influent bottle or tubing during the nBA amendment and the intended mass of nBA was not loaded into the column.

Recoveries for the 2,000 mg/L and neat nBA additions were much better. The recoveries were slightly greater than 100%, which is likely due to errors in the measurements and analytical analysis. The excellent recoveries from these two phases suggests that nBA was likely not completely oxidized during the first two nBA additions and further support the theory that the nBA mass from the first two additions was not sufficiently loaded into the column. The apparent losses of nBA during first two additions make it difficult to determine the partitioning behavior of saturated nBA under these conditions.

#### **7.4 Effects of Biostimulation**

Following electron donor addition on day 47 *cis*-DCE increased in both donor amended columns. The concentration of *cis*-DCE did not increase as high as the original increase at the beginning of the operational period, but do indicate that the indigenous microorganisms were able to dechlorinate at least a portion of the TCE to *cis*-DCE with the addition of nBA or SRS<sup>TM</sup> as electron donor, perhaps in microenvironments not exposed to high concentrations of TCE (Isalou and Sleep, 1998 and Dennis et al., 2003). Although the effluent TCE concentration was near saturation levels, the TCE concentration in portions of the column is likely lower. Sulfate reduction was not observed in the nBA or SRS<sup>TM</sup> columns after electron donor addition possibly due to toxic effects of the high TCE concentrations to sulfate reducers. Sulfate reduction was observed in SRS<sup>TM</sup> and nBA amended microcosms in the Sabre microcosm study indicating that the aquifer material did indeed contain organisms capable of sulfate reduction.

Acetate was present in the effluent after the first nBA addition indicating that nBA hydrolysis and/or fermentation occurred releasing electron donor available for sulfate reduction. There was a small amount of butyrate measured indicating that some H<sub>2</sub> may have been produced from fermentation of butanol (equation 2.2) The H<sub>2</sub> needed for the TCE to *cis*-DCE dechlorination likely came from here. The lack of mass balance from the first nBA addition makes it difficult to speculate on the fate of the electron donor and the amount of electron donor that was actually available for use in sulfate reduction and dechlorination during this initial phase. With only 7.2% of the added nBA recovered it is likely that electron donor was limited during this initial phase.

The initial increase in lactate, butyrate, and propionate in the effluent of the SRS<sup>TM</sup> column was attributed to the lactate component of the SRS<sup>TM</sup>. The SRS<sup>TM</sup> electron donor contains 4% lactate to initially stimulate biological activity when added to the subsurface because lactate is readily available to many microorganisms. The initial propionate and acetate concentrations observed were likely due to fermentation of lactate. Aulenta et al. (2007a) observed that lactate was fermented to acetate and propionate in the absence of sulfate reduction. The lactate component of the SRS<sup>TM</sup> was not completely fermented as evident by the lactate measured in the effluent. This suggests that microbial activity in the column was limited initially. The lactate, propionate, and butyrate in the effluent represented a source of hydrogen that was not utilized and essentially wasted in the column effluent. Acetate only increased slightly and since LCFAs fermented from the oil initially sorb to sediment surfaces and then are slowly fermented (Long and Borden, 2006); it is likely that the acetate at this stage was from lactate fermentation to acetate and propionate rather than from fermentation of the oil component of the SRS. This suggests that H<sub>2</sub> production from the fermentation of LCFAs was not occurring, or occurring slowly. Also, given the likelihood that lactate fermentation followed the pathway to acetate and propionate, instead of acetate to H<sub>2</sub>, (as indicated by the increase in propionate concentrations) the H<sub>2</sub> available for dechlorination may have been limited during this initial phase after donor addition and before bioaugmentation.

Biostimulation did not promote significant dechlorination or sulfate reduction in the time frame observed here. However, this phase is important to condition the columns prior to bioaugmentation with the KB-1<sup>®</sup> culture. The KB-1<sup>®</sup> culture requires reducing conditions, which are typically achieved by first amending the subsurface with electron donor and allowing the indigenous organisms to consume oxygen and created the reducing conditions.



## 7.5 Effects of Bioaugmentation

### 7.5.1 nBA column

Bioaugmentation with KB-1<sup>®</sup> promoted limited dechlorination of TCE to *cis*-DCE initially and after the second nBA amendment phase. The mass balance and availability of nBA was slightly better in the second addition than in the first, but it is still possible that there were limited amounts of electron donor available in this phase. During the 2,000 mg/L nBA amendment phase sulfate concentrations dropped rapidly and *cis*-DCE concentrations increased in the presence of high TCE concentrations indicating that sulfate reduction and dechlorination was possible in this column. Acetate, butanol, and butyrate were all present in the effluent and the nBA mass balance was excellent. Butanol and butyrate detected in the column effluent represented untapped H<sub>2</sub> sources, but there was likely more acetate and H<sub>2</sub> available than after the first two additions. The activity observed during this phase, along with the decrease in activity after the 2,000 mg/L nBA addition phase, indicate that there was sufficient electron donor available to promote dechlorination and sulfate reduction in this phase and that dechlorination of TCE to *cis*-DCE was possible with TCE DNAPL present.

VC and ethene concentrations did not increase throughout the operational period, even during the continuous addition phase. The concentrations of *Dhc* and *vcrA* did not increase in the effluent over the operational period indicating the conditions in the column were not favorable for the growth of the organisms responsible for the final dechlorination steps (i.e. *cis*-DCE to VC and ethene). Several other studies have observed similar results in the presence of DNAPL (Adamson et al., 2003, Carr et al., 2000, and Cope and Hughes, 2001), suggesting that complete dechlorination of TCE to ethene should not be expected in the immediate vicinity of the source zone. Bioaugmentation with KB-1<sup>®</sup> and the use of a PED as electron donor were used to overcome this expectation, but were not able to promote complete dechlorination to ethene in this column scenario. The concentrations of *Dhc* and *vcrA* did not increase during the periods of *cis*-DCE production indicating that these organisms were not likely the organisms responsible for the TCE to *cis*-DCE dechlorination. Many organisms have been shown to be capable of PCE and TCE dechlorination to *cis*-DCE (Da Silva et al., 2006). It is likely that either intrinsic dechlorinators or dechlorinators in the KB-1<sup>®</sup> culture performed the observed dechlorination. The KB-1<sup>®</sup> culture contains *Geobacter* species, which is one possible candidate for the TCE to *cis*-DCE dechlorination. Duhamel et al., 2004 observed a *Geobacter* population increase during TCE to *cis*-DCE dechlorination suggesting that this organism grew from

the dechlorination of TCE to *cis*-DCE. Also, Sung et al. (2006) demonstrated that *Geobacter* strain SZ was capable of using PCE and TCE as metabolic electron acceptors and that 16srRNA gene sequences with high similarity to strain SZ were detected in the KB-1<sup>®</sup> culture.

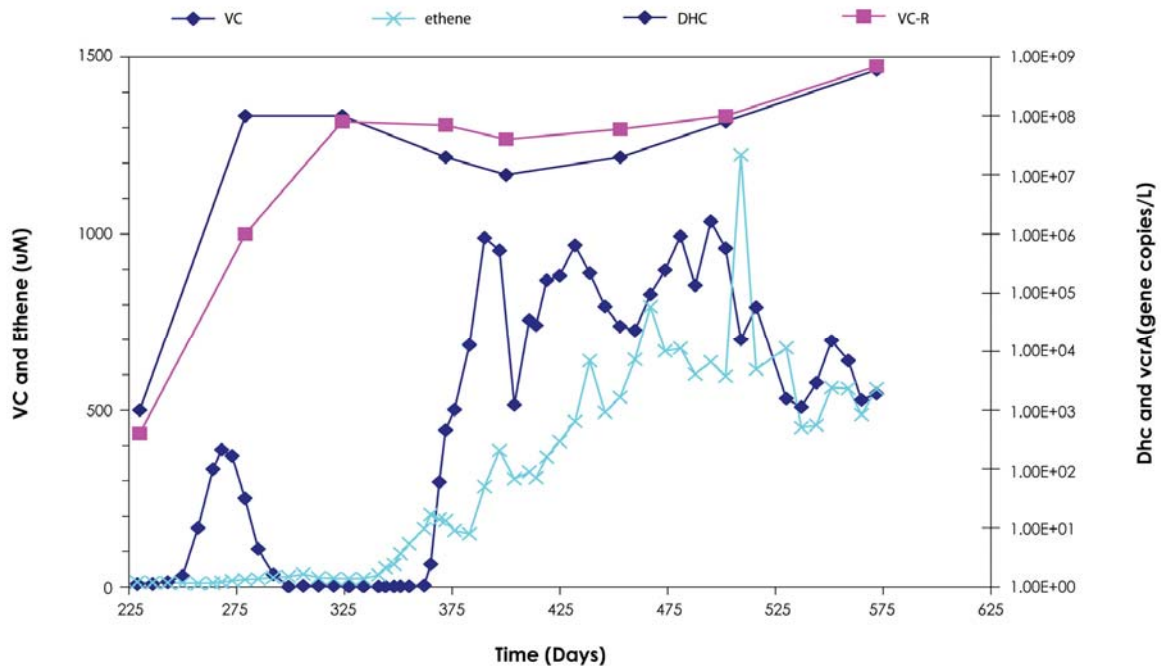
The continuous addition of nBA during the third addition phase when the sulfate reduction and dechlorination were observed was more representative of a soluble electron donor delivery strategy rather than a PED addition. The nBA appeared to break down readily in the column influent bag (Table 5.3), and likely broke down readily once in the column so the nBA may not have made contact with the DNAPL phase and therefore could not partition into the DNAPL. The partitioning properties of the nBA breakdown products are less than nBA. Butanol is the closest with a log  $K_{ow}$  of 0.88 (Montgomery, 2000), butyric acid was estimated to be 0.79 +/-0.12 (Lyman et al, 1990), and acetic acid is negative with a log  $K_{ow}$  of -0.29 (Montgomery, 2000).

There was only limited data obtained after the neat nBA addition. The strategy for the neat nBA addition was to add the nBA, allow the nBA to partition into the DNAPL, and then re-bioaugment the column with KB-1<sup>®</sup>. Re-bioaugmentation would be needed because the neat nBA addition would have inhibited dechlorinating organisms present in the column. The nBA concentrations would be lower when partitioning out of the DNAPL back into the aqueous phase providing the needed electron donor to promote dechlorination. Unfortunately the column broke before it was re-bioaugmented with KB-1<sup>®</sup>. The limited data collected suggests that the neat nBA addition inhibited fermenting organisms in the column. High concentrations of butanol and only very low butyrate concentrations were measured in the effluent indicating that fermentation of butanol to butyrate was not a major process as it was during the 2,000 mg/L nBA addition phase.

### **7.5.2 SRS<sup>™</sup> Column**

During the DNAPL phase bioaugmentation with KB-1<sup>®</sup> promoted dechlorination of TCE to *cis*-DCE with trace amounts of VC and ethene. Sulfate reduction occurred almost immediately following bioaugmentation. TCE concentrations decreased slowly after bioaugmentation, but remained above 500 mg/L (consistent with concentrations representing the presence of DNAPL) with increases in *cis*-DCE production. *Dhc* and *vcrA* concentrations remained low during *cis*-DCE production indicating that other organisms were likely responsible for the TCE to *cis*-DCE dechlorination as discussed above for the nBA column. Following TCE decrease and the production of *cis*-DCE, VC, and ethene

after day 230 concentrations of Dhc and vcrA increased to  $1 \times 10^8$  and  $8 \times 10^7$  gene copies/L respectively. Figure 7.3 shows the increases in Dhc and vcrA with increases in VC and ethene over this time frame and to the end of the operational period. The concentrations of Dhc measured here are consistent with levels where ethene production is expected (Lu et al., 2006). The concentration of Dhc and vcrA remained consistent throughout the plume phase when high levels of VC and ethene continued to be measured in the effluent. These results indicate that the dechlorinating organisms were robust enough to recover from the stresses of high TCE concentrations and dechlorinate through to ethene once the DNAPL was depleted.



**Figure 7.4** Dhc, vcrA, VC, and Ethene Concentration Trends in the SRS Column

Following bioaugmentation the concentration of acetate in the effluent increased rapidly indicating that the LCFA in the SRS<sup>TM</sup> were being fermented and producing the needed H<sub>2</sub> and acetate for dechlorination and sulfate reduction. Dechlorination continued after acetate was no longer detected in the effluent from day 404 through day 572. This was an indication that the useable electron donor was consumed or fermentation of the LCFAs was no longer occurring. There was no noticeable change in dechlorination activity as TCE dechlorination products were continuously measured in the effluent. However, sulfate concentrations began to increase around the same time as the acetate was

no longer detected in the effluent. The sulfate concentrations measured in the effluent never reached as high as influent levels indicating that sulfate reduction was still occurring, which required a source of electron donor. It is possible that there were low flow zones in the column where a portion of the SRS<sup>TM</sup> was trapped and was slowly being fermented supplying the needed electron donor. Sulfate reducers can use acetate directly (Widdel, 1988) so even though acetate wasn't observed in the effluent it is possible that acetate was present in the column and was consumed by sulfate reducers somewhere in the column prior to reaching the effluent. If sulfate reducers were using acetate then it is possible that the dechlorinators benefited from the H<sub>2</sub> produced by the sulfate reducing bacteria as proposed by Aulenta et al. (2007a).

Methane concentrations began to increase in the DNAPL phase following day 150 when high concentrations of TCE and *cis*-DCE were present in the column at concentrations known to inhibit methanogenesis (Yang and McCarty (2000); perhaps in microsites. Methane concentrations continued to increase when chlorinated ethenes were not measured in the effluent. The methanogens benefited during this phase of the column operation, not only from the low chlorinated ethene concentrations, but also from the lack of competition for electron donor from dechlorinators. During this period acetate concentrations were high and it is possible that acetoclastic methanogens as well as hydrogenic methanogens were active. Methane concentrations decreased almost immediately following the start of the plume phase when the column was impacted by high TCE concentrations in the influent AGW. This suggests that methanogens present in the column were inhibited by TCE.

During the plume phase methanogens were at a disadvantage, not only from the high chlorinated ethene concentrations, but also from competition for electron donor by dechlorinators and sulfate reducers. Electron donor was likely limited during much of this phase as evident by the non-detect levels of VFAs in the effluent. Sulfate reducers and dechlorinators have a much lower hydrogen threshold than methanogens so it is possible that the H<sub>2</sub> concentrations were maintained below the threshold for methanogens. Therefore, even if microsites with low TCE concentrations existed, there may not have been enough H<sub>2</sub> available for the methanogens.

The ORP in the column effluent began to increase after day 502 and reached positive values (e.g. +161 mV) on day 551. This corresponds with the decrease in the breakdown products of the SRS<sup>TM</sup> and its apparent depletion indicating that conditions, at least in the column effluent, were not as

reducing as they were earlier in the operational phase. However, dechlorination products were measured in the effluent during this period and so it appears that reductive dechlorination was still active, perhaps in anaerobic microsites.

The pH in the SRS<sup>TM</sup> column decreased to near 6 during the period when *cis*-DCE production was occurring. This is below the optimum pH for reductive dechlorination. This suggests that the acid-buffering capacity of the site materials is not sufficient to maintain a neutral pH during acid production from dechlorination. The pH of the effluent was easily increased to near neutral with the addition of bicarbonate in the influent groundwater. This demonstrates that pH buffering may be required in the field to ensure optimal conditions for dechlorination are maintained.

### **7.5.3 Importance of Sulfate Reduction and Dechlorination in the SRS<sup>TM</sup> Column**

The observed dechlorination of TCE and *cis*-DCE through to VC and ethene in the presence of high sulfate concentrations is a very important finding. There are reports in the literature that suggest dechlorination can take place with sulfate present (Aulenta et al., 2007a) and that sulfate had no inhibition effect on dechlorination (Hoelen and Reinhard, 2004). However, there are other studies that have shown sulfate can negatively affect dechlorination at low H<sub>2</sub> concentrations (Heimann et al., 2005) and that sulfate completely inhibited dechlorination (Nelson et al., 2002). The issue is still controversial and the typical approach used in bioremediation projects is to add enough electron donor to deplete the sulfate first (Aulenta et al., 2007) prior to expecting dechlorination to occur. This is significant because sulfate is often the main electron donor sink and drives the cost of electron donor required for bioremediation projects. It is likely in the plume phase of the SRS<sup>TM</sup> column that H<sub>2</sub> concentrations were low. The results here suggest that dechlorinators are able to compete with the sulfate reducers for H<sub>2</sub> under H<sub>2</sub> limiting conditions and that it is not necessary to deplete sulfate concentrations in order for dechlorination to proceed through to ethene. It is possible that sulfate reducers were inhibited by the high concentrations of TCE allowing the dechlorinating organisms to function in areas where sulfate reducers were not able to compete for electron donor. However, the results from the microcosm study showed that sulfate reducers were able to reduce sulfate in the presence of 800 mg/L of TCE so sulfate reducers and dechlorinators could have co-existed in the column and competed for available electron donor.

## 7.6 Mass balances

### 7.6.1 Chlorinated Ethene and Ethene Mass Balances

Throughout the column study operational period the chlorinated ethene and ethene concentrations in the effluent of the columns were summed to calculate the amount of chlorinated ethene and ethene mass being removed from the columns. The data is summarized in Table 7.4, an example calculation of cumulative mass removal is shown in Appendix D, and the raw data is provided in Appendix E. Over the operational period 57% of the TCE DNAPL in the control column was removed, mainly due to dissolution with 99.87% of the measured chlorinated ethenes removed as TCE. In the nBA column 50 % of the TCE DNAPL was removed in the 306 day operational period before the column broke. The removal was mainly as TCE dissolution (91 %) with 9 % *cis*-DCE, and trace VC measured. The DNAPL was believed to be depleted after about 300 days in the SRS<sup>TM</sup> column as observed by the decrease in TCE and dechlorination products mass measured in the column effluent. The plume phase of the SRS<sup>TM</sup> column did not begin until day 329 and all data up to day 327 was used in the DNAPL phase mass balance calculations. Only 76% of the TCE DNAPL retained in the column was accounted for in the effluent as TCE or dechlorination products. Again, the majority of the mass removed was as TCE (77%); *cis*-DCE accounted for 22% of the total mass removed, with VC and ethene making up the remaining 1%.

The chlorinated ethene and ethene concentrations in the effluent of the SRS<sup>TM</sup> column during the plume phase were summed to compare with the mass of TCE added in the influent AGW. During this phase 94% of the added TCE was accounted for in the effluent with 0% as TCE, 27% as *cis*-DCE, 40% as VC, and 33% as ethene. The mass balance in this case was much better than in the DNAPL phase with only 6% of the added TCE unaccounted for.

The unaccounted for TCE mass in the SRS<sup>TM</sup> column could be due to degradation of the chlorinated ethenes and ethene via a different pathway to end products that were not measured in this study. Anaerobic oxidation of VC to CO<sub>2</sub> by indigenous microorganisms has been demonstrated under redox conditions ranging from methanogenic to nitrate reducing (Bradley and Chapelle, 2000)). Also, Bradley and Chapelle (2002) demonstrated that under sulfate reducing conditions, bed sediment microorganisms completely oxidized ethene to CO<sub>2</sub>. The mass balances obtained in the SRS<sup>TM</sup> column during both the DNAPL and the plume phases did not account for all of the mass added.

Anaerobic oxidation of VC and/or ethene is a possible explanation for the missing mass. Mass loss was also observed in the GE column, which was attributed to anaerobic oxidation processes.

### **7.6.2 Chloride Mass Balances**

Chloride is often used as a metric to verify measurements in dechlorination studies because it is a conservative ion (i.e., it does not degrade) and it is easy to measure. The concentration of chloride is expected to increase during the reductive dechlorination process. As seen in Figure 2.3 one mole of chloride is produced for each dechlorination step and if TCE is completely dechlorinated to ethene three moles of chloride will be produced. Mass balances were performed using the chloride measured in the effluent of the columns with the observed dechlorination products. A chloride mass balance was not performed on the control column because chloride concentrations were not observed to be higher than the background chloride levels in the AGW.

The measured chloride was slightly greater than the chloride expected from dechlorination in the nBA and SRS<sup>TM</sup> columns. Differences were 6.0 %, 2.5 %, and 3.5 % for the nBA column, SRS<sup>TM</sup> column DNAPL phase, and SRS column plume phase respectively (Table 7.5) and are not considered significantly different from 0%. This indicates that chloride is an excellent parameter to measure during the reductive dechlorination process and can be used to verify chlorinated ethene and ethene concentrations in bioremediation studies. Refer to Appendix D for an example calculation of cumulative mass removal and Appendix E for the raw data.

**Table 7.4** Summary of Chlorinated Ethene and Ethene Mass Balance Results

Column	TCE added (mmoles)	TCE mmoles (% of total mmoles)	Cis-DCE mmoles (% of total mmoles)	VC mmoles (% of total mmoles)	Ethene mmoles (% of total mmoles)	Total mmoles	Amount Removed (%)
Control	691	394 (99.87)	0.51 (0.13)	0.00 (0.00)	0.00 (0.00)	394	57
nBA	446	201 (90.66)	21 (9.3)	0.097 (0.044)	0.00 (0.00)	222	50
SRS (DNAPL Phase)	240	140(77.21)	40 (22)	1.0(0.57)	0.31 (0.17)	182	76
SRS (Plume Phase)	52	0.00 (0.00)	13 (27)	19 (39)	16 (33)	49	94

Notes: Bracketed values are the percentage of the total mmoles measured for that compound



**Table 7.5** Summary of Chloride Mass Balance Results

	Control column					nBA Column				
	Measured in effluent (mmoles)	mmoles of Cl released per mole	Cl released during dechlorination (mmoles)	Cl Measured (mmoles)	RPD	Measured in effluent (mmoles)	mmoles of Cl released per mole of VOC	Cl released during dechlorination (mmoles)	Cl Measured (mmoles)	RPD
TCE	394	0	0	--	--	201	0	0	--	--
cis-DCE	0.51	1	0.51	--	--	20.6	1	20.6	--	--
VC	0.00	2	0	--	--	0.097	2	0.19	--	--
Ethene	0.00	3	0	--	--	0	3	0	--	--
Total	395	--	0.51	ND	--	222	--	20.8	21.9	5.58

	SRS Column (DNAPL Phase)					SRS column (Plume Phase)				
	Measured in effluent (mmoles)	mmoles of Cl released per mole	Cl released during dechlorination (mmoles)	Cl Measured (mmoles)	RPD	Measured in effluent (mmoles)	mmoles of Cl released per mole of VOC	Cl released during dechlorination (mmoles)	Cl Measured (mmoles)	RPD
TCE	140	0	0	--	--	0	0	0	--	--
cis-DCE	40.1	1	40.1	--	--	13.3	1	13.1	--	--
VC	1.04	2	2.08	--	--	19.1	2	38.2	--	--
Ethene	0.31	3	0.94	--	--	16.1	3	48.4	--	--
Total	182	--	43.1	44.2	2.46	48.6	--	99.9	104	3.52

## 7.7 Evidence of Enhanced Dissolution

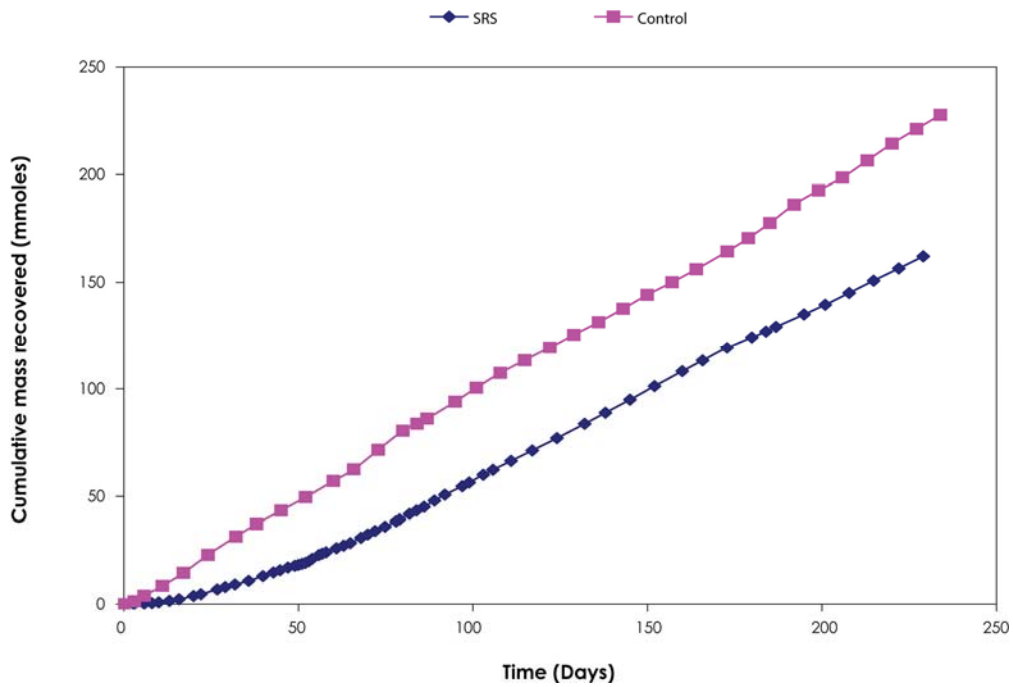
Several studies have demonstrated biologically induced enhanced dissolution of PCE DNAPL in column studies. Enhancement factors of 1.6 (Da Silva et al., 2006), 5.0 (Yang and McCarty, 2000), and 16 (Cope and Hughes, 2001) have been reported. Similar studies using TCE DNAPL could not be found in the literature. The solubility of PCE (140 mg/L) is much less than TCE (1,100 mg/L) and the  $K_{OW}$  of TCE is 9 times less than PCE (Adamson et al., 2003). Therefore, higher enhancement factors can be expected for PCE bioremediation as the difference in the solubility's and partitioning properties of dechlorination products is greater for PCE than for TCE. For example the solubility of cis-DCE (3,500 mg/L) is 23 times greater than PCE, but only 3 times greater than TCE. Nonetheless, the production of less chlorinated compounds from TCE DNAPL will increase the concentration gradient and allow more TCE to dissolve into the aqueous phase and increase the removal rate from a source zone relative to the dissolution of TCE alone.

Enhancement factors are typically calculated as the total moles removed in a treatment column divided by the total moles removed due to TCE dissolution only in a control column (Cope and Hughes, 2001). The cumulative mass recovery in the SRS<sup>TM</sup> column and control column was compared over the first 229 days of column operation (Figure 7.5). This time period was chosen because this was before the total mass in the SRS<sup>TM</sup> column dropped rapidly indicating that DNAPL was no longer present. The cumulative mass recovery in the nBA column and control column was compared over the entire 306 days of column operation (Figure 7.6). In both cases the mass removed due to dissolution alone in the control column was greater than the mass removed in the treatment columns. Enhancement factors of 0.71 and 0.77 were calculated for the SRS<sup>TM</sup> and nBA columns respectively. Enhancement factors less than one indicate the mass removal was actually greater in the control column than in the treatment columns. However, as discussed above, the amount of DNAPL added to each column was different and the dissolution profiles before any amendments were made to the treatment columns were not the same as the control column over the same time period. Therefore the dissolution profile from the control column is likely not a good indication of what the untreated dissolution profile would be in the treatment columns.

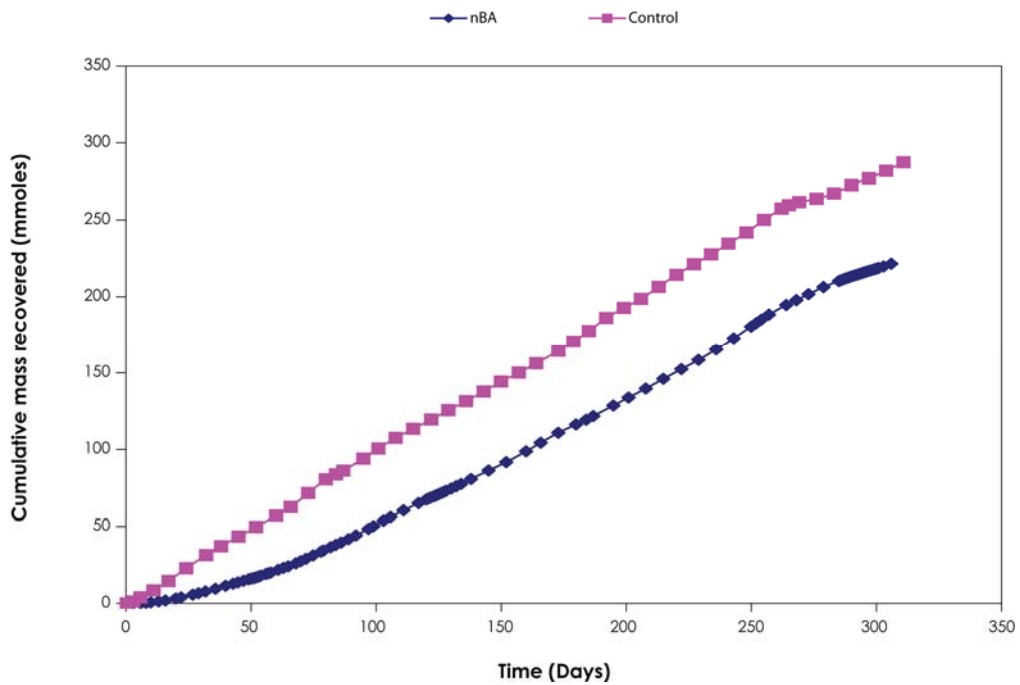
Non treated dissolution profiles representative of the treatment columns were not available so a different method of estimating enhanced dissolution was needed. The treatment columns were normalized to the control column by comparing the cumulative mass removed divided by the initial mass retained in the control column with the cumulative mass removed divided by the initial mass retained in the treatment columns. Refer to Appendix D for an example calculation of cumulative mass removal and Appendix E for the raw data. The normalized enhanced dissolution profiles for the SRS<sup>TM</sup> and nBA columns are shown in Figures 7.7 and 7.8 respectively.

Figure 7.7 shows that before biostimulation in the SRS<sup>TM</sup> column the normalized dissolution in the control column was slightly greater than in the SRS<sup>TM</sup> column. After biostimulation on day 47 the SRS<sup>TM</sup> profile began to increase relative to the control profile and this increased further after bioaugmentation on day 79 indicating that biostimulation and bioaugmentation increased the rate of mass removal relative to the control column. After 229 days of column operation 68% of the mass was removed from the SRS<sup>TM</sup> column versus only 33% of the mass in the control column, or 2.1 times more of the relative mass was removed from the SRS<sup>TM</sup> column compared to the control column.

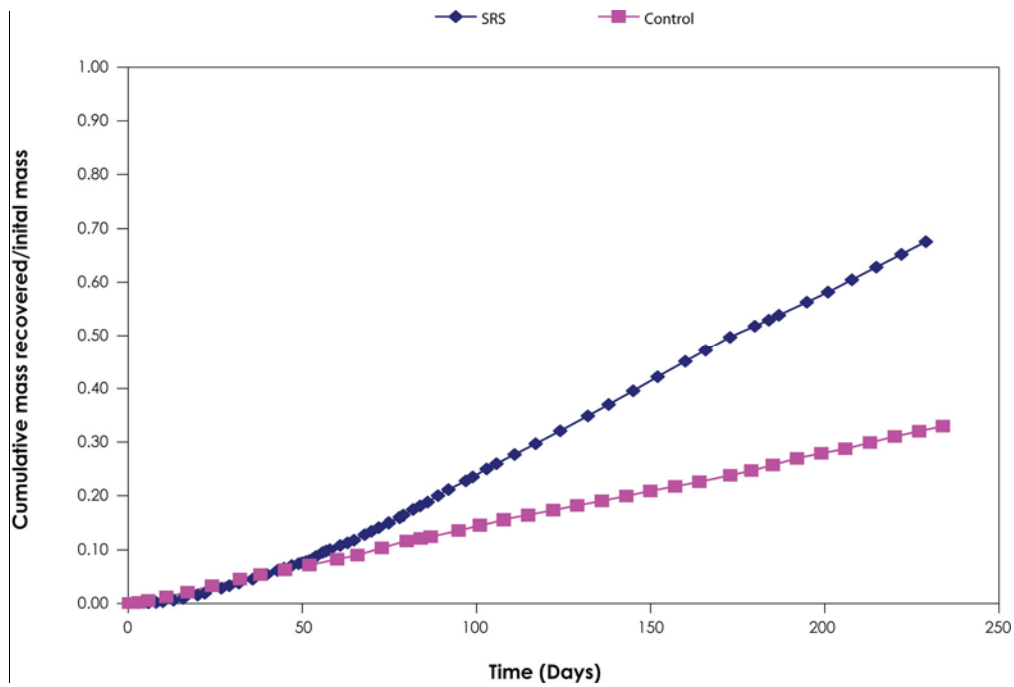
The normalized dissolution in the control column was greater than in the nBA amended column up to the 2,000 mg/L nBA addition phase on day 209 (Figure 7.8). After day 209 the nBA dissolution increased above the control to the end of the operational period on day 306, although the differences weren't as great as in the SRS<sup>TM</sup> column. 50% of the mass was removed from the nBA column versus 42% in the control column, or 1.2 times more of the relative mass was removed from the nBA column compared to the control column.



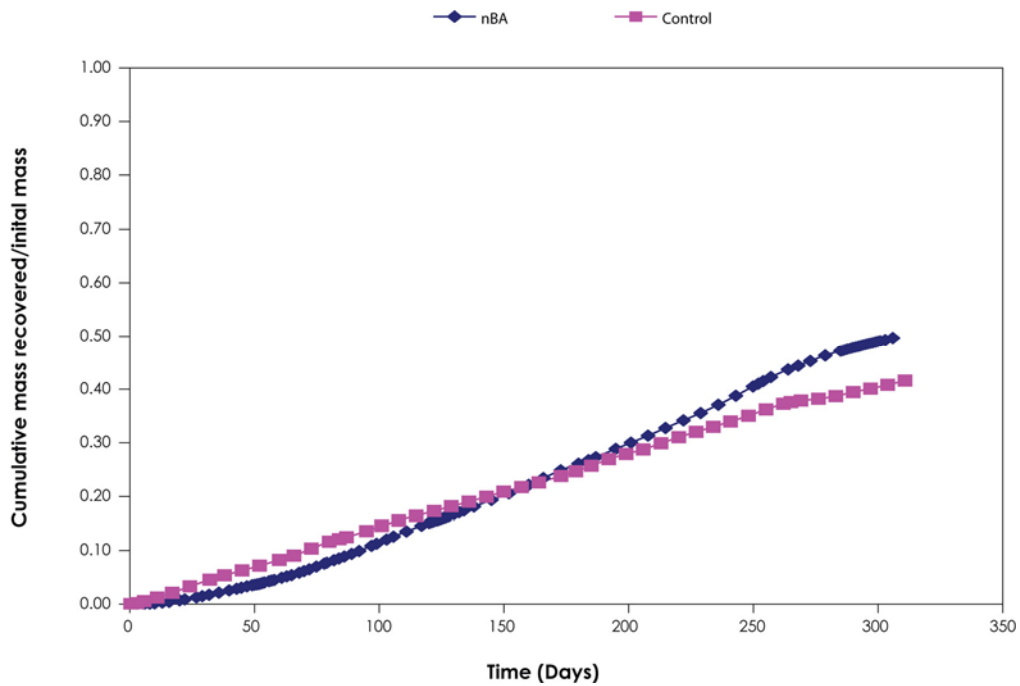
**Figure 7.5** Cumulative Mass Recovered in the SRS<sup>TM</sup> and Control Columns



**Figure 7.6** Cumulative Mass Recovered in the nBA and Control Columns



**Figure 7.7** Normalized Enhanced Dissolution in the SRS<sup>TM</sup> and Control Columns



**Figure 7.8** Normalized Enhanced Dissolution in the nBA and Control Columns

## Chapter 8

### Conclusions and Future Work

#### 8.1 Conclusions

This research focused on enhancing the dissolution from TCE DNAPL source areas using bioremediation with partitioning electron donors and bioaugmentation as the remedial technology. The objectives of this research were to compare two partitioning electron donors in their ability to stimulate reductive dechlorination in the presence of TCE DNAPL and estimate the extent of DNAPL dissolution enhancement created by the bioremediation process. The following conclusions are drawn.

1. Both nBA and SRS<sup>TM</sup> electron donors were capable of supporting dechlorination of TCE to *cis*-DCE when TCE DNAPL was present in the column. The presence of DNAPL inhibited complete dechlorination to ethene.
2. Dhc were not responsible for the TCE to *cis*-DCE dechlorination in the presence of high TCE concentrations, the dechlorination was performed by other dechlorinating organisms. However, Dhc organisms were robust enough to recover from the stresses of the high TCE concentrations and dechlorinated TCE through to ethene following DNAPL depletion in the SRS<sup>TM</sup> column.
3. nBA was unstable in groundwater and flushed through the column rapidly. Activity was not maintained after the 2,000 mg/L nBA addition indicating limited electron donor retention within the column.
4. The addition of SRS<sup>TM</sup> promoted sulfate reduction and dechlorination for 1.5 years, which was longer than anticipated. Anaerobic activity was maintained after SRS<sup>TM</sup> breakdown products (VFAs) were not detected indicating that other metrics (dechlorination and sulfate reduction) are needed before reamending with electron donor.
5. Dechlorination to ethene was achieved in the presence of high sulfate concentrations. It may not be necessary to deplete sulfate prior to achieving complete dechlorination, which will decrease the amount of electron donor required.
6. The buffering capacity of the site material is not sufficient to buffer acid produced during dechlorination. pH needs to be carefully monitored during bioremediation studies and adjusted if required to ensure optimum dechlorinating conditions are maintained.

7. By normalizing the mass in the control column to the treatment columns enhancement factors of 2.1 and 1.2 were calculated for the SRS<sup>TM</sup> and nBA columns respectively.

## 8.2 Future Work

At the time this report was written the Sabre project was still underway and the following future work related to the SiREM column study is planned.

1. Perform a final tracer tests on the SRS<sup>TM</sup> column to compare the porosity and dispersivity values to the initial tracer test. This will be used to determine if the porosity decreased throughout the study due to the bioactivity in the column.
2. Analyze archived DNA samples collected from the nBA and SRS<sup>TM</sup> columns for *Geobacter* to verify the hypothesis that *Geobacter* was responsible for the TCE to *cis*-DCE dechlorination. It is expected that *Geobacter* counts will be high in the SRS<sup>TM</sup> column during the TCE to *cis*-DCE dechlorination and decrease once the Dhc concentrations increased during VC and ethene production.
3. Dismantle the SRS<sup>TM</sup> column and collect soil samples along the profile of the column. Extract DNA from different soil zones and analyze for Dhc, *vcrA*, and *Geobacter* counts to determine how these organisms are distributed in the column. Samples will also be submitted for denatured gradient gel electrophoresis (DGGE) to determine if there are changes in the microbial communities in different zones of the column.
4. Continue to run the control column at the experimental flow rate until the TCE DNAPL is depleted. This will be used to compare the dissolution curve as TCE is depleted from dissolution alone to the bioenhanced dissolution profile in the SRS<sup>TM</sup> column.

## References

- Adamson, D. T.; McDade, J. M.; Hughes, J. B. 2003. Inoculation of a DNAPL source zone to initiate reductive dechlorination of PCE. *Environ. Sci. Technol.*, 37(11), 2525-2533.
- Adamson, D. T., D. Y. Lyon, and J. B. Hughes. 2004. Flux and product distribution during biological treatment of tetrachloroethene dense non-aqueous-phase liquid. *Environ. Sci. Technol.*, 38(7), 2012-2028.
- Adrian, L., Szewzyk, U., Wecke, J., and Gorisch, H. 2000. Bacterial dehalorespiration with chlorinated benzenes. *Nature*. 408:580-583
- Agency for toxic Substance and Disease Registry. 1997. Toxicological Profile for Trichloroethylene (update).
- American Bar Association (ABA). 2005. Site Remediation Committee Newsletter. Vol.1, No. 1.
- Amos, B.K, Christ, J.A, Abriola, L.M, Pennell, D., Loffler, F.E. 2007. Experimental Evaluation and Mathematical Modeling of Microbially enhanced Tetrachloroethene (PCE) Dissolution. *Environ. Sci. Technol.*, 41; 963- 970
- Aulenta, F., Majone, M., Verbo, P., Tandoi, V. 2002. Complete dechlorination of tetrachloroethene to ethene in presence of methanogenesis and acetogenesis by an anaerobic sediment microcosm. *Biodegradation*. 13: 411-424.
- Aulenta, F., Majone, M., Tandoi, V. 2006. Enhanced anaerobic bioremediation of chlorinated solvents: environmental factors influencing microbial activity and their relevance under field conditions. *J Chem technol Biotechnol.*, 81, 1463-1474.
- Aulenta, F., Pera, A., Rossetti, S., Petrangeli, P., Majone, M. 2007. Relevance of side reactions in anaerobic reductive dechlorination microcosms amended with different electron donors. *Water Research*. 41: 27-38.
- Bradley, P.M., and Chapelle, F.H. 2000. Microbial degradation of chloroethenes in groundwater systems. *Hydrogeology J*. 8:104-111
- Bradley, P.M., and Chapelle, F.H. 2002. Mineralization of ethene under sulfate-reducing conditions. *Bioremediation Journal*. 6(1): 1-8.
- Cabirol, N., Jacob, F., Perrier, J., Fouillet, B., and Chambon, P. 1998. Interaction between methanogenic and sulfate-reducing microorganisms during dechlorination of a high concentration of tetrachloroethylene. *J Gen Appl Microbiol*. 44: 297-301
- Carr, C. S.; Garg, S.; Hughes, J. B. 2000. Effect of dechlorinating bacteria on the longevity and composition of PCE-containing nonaqueous phase liquids under equilibrium dissolution conditions. *Environ. Sci. Technol.*, 34(6), 1088-0194.
- Cope, N.; Hughes, J. B. 2001. Biologically-enhanced removal of PCE from NAPL source zones *Environ. Sci. Technol.*, 35(10); 2014-2021.
- Da Silva, M.L.B., Daprot, R.C., Gomez, D.E., Hughes, J.B., Ward, C.H., Alvarez, J.J. 2006. Comparison of Bioaugmentation and Biostimulation for the Enhancement of Dense Nonaqueous Phase Liquid Source Zone Bioremediation. *Water Environment Research*, 78 (13), 2456-2465
- David, R.M., Tyler, T.R, Ouellette, R., Faber, W.D., Banton, M.I. 2001. Evaluation of subchronic toxicity of n-butyl acetate vapor. *Food and chemical Toxicology*. 39: 877-886



- Denis, D.C., Sleep, B.E., Fulthorpe, R.R., Liss, S.N. 2003. Phylogenetic analysis of bacterial populations in an anaerobic microbial consortium capable of degrading saturation concentrations of tetrachloroethylene. *Can. J. Microbiol.* 49: 15-27
- Duhamel, M., Wehr, S.D., Yu, L., Rizvi, H., Seepersad, D., Dworatzek, S., Cox, E.E, and Edwards, E.A. 2002. Comparison of anaerobic dechlorinating enrichment cultures maintained on tetrachloroethene, trichloroethene, cis-dichloroethene, and vinyl chloride. *Water Research.* 36:4193-4202
- Duhamel, M.A., Mo, K. and Edwards, E.A. 2004. Characterization of a highly enriched Dehalococcoides - containing culture that grows on vinyl chloride and trichloroethene. *Applied Environmental Microbiology.* 70:5538-5545
- Duhamel, M.A. 2005. Community Structure and dynamics of Anaerobic Chlorinated Ethene-Degrading Enrichment Cultures. PhD dissertation, University of Toronto, Toronto, Ont.
- EPA (2004a). CERCLIS database. United States Environmental Protection Agency. <http://cfpub.epa.gov/supercpad/cursites/srchsites.cfm>
- Environment Agency. 2003. An illustrated handbook of DNAPL transport and fate in the subsurface. Environment Agency Rio House.
- Fennell, D.E., and Gossett, J.M. 1998. Modeling the production of and competition for hydrogen in a dechlorinating culture. *Environ. Sci. Technol.*, 32(16), 2450
- Fennell, D.E., and Gossett, J.M. 2003. Microcosms for site specific evaluation of enhanced biological reductive dehalogenation, in *Dehalogenation: Microbial Processes and Environmental Applications*, ed. by Haggblom MM and Bossert ID. Kluwer Academic, Boston, MA, pp 385-420.
- Ferguson, J.F. and J.M.H. Pietari. 2000. Anaerobic transformations and bioremediation of chlorinated solvents. *Environmental Pollution.*, 107:209-215.
- Freedman, D.L.; Gossett, J.M. 1989. Biological reductive dechlorination of tetrachloroethylene and trichloroethylene to ethylene under methanogenic conditions. *Appl. Environ. Microbiol.*, 55(9), 2144-2151.
- Johnson, R.L.; Pankow, J.F. 1992. Dissolution of dense chlorinated solvents into groundwater. 2. Source functions for pools of solvent. *Environ Sci. Technol.*, 26, 896-901
- Hanaki, K., Nagase, M., Matsuo, T. 1981. Mechanism of inhibition caused by long-chain fatty acids in anaerobic digester process. *Biotechnol. Bioeng.* 23; 1591-1610
- Harkness, M.R. 2000. Economic considerations in enhanced anaerobic degradation. *Bioremediation and Phytoremediation of Chlorinated and Recalcitrant Compounds. 2<sup>nd</sup> Internat. Conf. Remediation of Chlorinated and Recalcitrant Compounds.* Monterey, CA.
- Harkness, M.R.; Bracco, A.A.; Brennan, M.J. Jr.; DeWeerd, K.A.; Spivack, J.L. 1999. Use of bioaugmentation to stimulate complete reductive dechlorination of trichloroethene in Dover soil columns. *Environ. Sci. Technol.* (7), 33:1100-1109.
- He, J., Ritalahti, M.K, Yang, K.U, Koeningsberg, S.S, Loffler, F.E. 2003. Detoxification of vinyl chloride to ethene coupled to growth of an anaerobic bacterium. *Nature.* 424: 62-65
- He, J., Sung, Y., Krajmalnik-Brown, R., Ritalahti, K.M, Loffler, F.E. 2005. Isolation and characterization of Dehalococcoides sp. Strain FL2, a trichloroethene (TCE)- and cis-1,2-dichloroethene-respiring anaerobe. *Environmental Microbiology.* 7:1442-1450.

- Heimann, A.C., Friis, A.K., Jakobsen, R. 2005. Effects of sulfate on anaerobic chloroethene degradation by an enriched culture under transient and steady-state hydrogen supply. *Water Research*. 39: 3579-3586
- Hendrickson, E.R., J. A. Payne, R. M. Young, M.G. Star, M. P. Perry, S. Fahnestock, D. E. Ellis and R.C. Ebersole. 2002. Molecular analysis of Dehalococcoides 16S ribosomal DNA from chloroethene-contaminated sites throughout North America and Europe. *Appl. Environ. Microbiol.* 68:485-495.
- Hoelen, T., and Reinhard, M. 2004. Complete biological dehalogenation of chlorinated ethylene's in sulfate containing ground-water. *Biodegradation*. 15:395-403.
- Isalou, M., Sleep, B.E. 1998. Biodegradation of High Concentrations of Tetrachloroethene in a Continuous Flow column System. *Environ. Sci. Technol.* (22), 32:3579-3585.
- Johnson, R.L., Pankow, J.F. 1992. Dissolution of dense chlorinated solvents into groundwater. 2. Source functions for pools of solvent. *Environ. Sci. Technol.* 26, 896-901.
- Kueper, B.H., Redman, J.D., Starr, R.C., Reitsma, S. and Mah, M. 1993. A field experiment to study the behavior of tetrachloroethylene below the water table: spatial distribution of residual and pooled DNAPL. *Journal of Groundwater*, 31 (5), 756-766
- Kokkinaki, A. 2007. Modeling of enhanced anaerobic biodegradation of trichloroethene. M.A.Sc. thesis, University of Toronto, Toronto, Ont.
- Lee, M. D.; Biehle, A. A.; Ellis, D. E..1999. Anaerobic bioremediation of DNAPL tetrachloroethene in sludge microcosm studies. In: A. Leeson and B. C. Alleman, eds., *Engineered Approaches for In Situ Bioremediation of Chlorinated Solvent Contamination*. 5(2):55-60.
- Lee, M.D., Lieberman, M.T., Beckwith, W., Borden, R.C, Haas, P., Becvar, E.S., Dobson, K., Sandlin, G.J. 2005. Vegetable oil pilots to enhance DNAPL sequestration and reductive dechlorination. *Proceedings of Remediation of chlorinated and Recalcitrant Compounds – 8<sup>th</sup> Internat. Conf., Baltimore, MD*
- Long, C., Borden, R.C. 2006. Enhanced reductive dechlorination in columns treated with edible oil emulsion. *J. of Contaminant Hydrology*. 87: 54-72.
- Lu, X., Wilson, J.T., Kampbell, D.H. 2006. Relationship between Dehalococcoides DNA in ground water and rates of reductive dechlorination at field scale. *Water Research*. 40: 3131-3140.
- Lyman, W., Reehl, W., and Rosenblatt, D. 1990. *Handbook of Chemical Property Estimation Methods: Environmental Behavior of Organic Compounds*. American Chemical Society, Washington, DC.
- Mackay, D., Shiu, W., and Ma, K. 1993-1996. *Illustrated Handbook of Physical-Chemical Properties and Environmental Fate for Organic Chemicals*, Vol. I-IV. Lewis Publishers, Boca Raton, FL.
- May, P. Linoleic acid. 2008. <http://www.chm.bris.ac.uk/motm/linoleic/linoleic.htm> (accessed January 2008)
- Maymo-Gatell, X.; Tandoi, V.; Gossett, J.M.; Zinder, S.H. 1995. Characterization of an H<sub>2</sub>-utilizing enrichment culture that reductively dechlorinates tetrachloroethene to vinyl chloride and ethene in the absence of methanogenesis and acetogenesis", *Appl. Environ. Microbiol.* 61(11), 3928-3933.

- Maymo-Gatell, X., Chien, Y.T., Gosset, J.M., Zinder, S.H. 1997. Isolation of a bacterium that reductively dechlorinates tetrachloroethene to ethene. *Science*. 276: 1568-1571.
- Maymo-Gatell, X., Anguish, T., and Zinder, S.H. 1999. Reductive dechlorination of chlorinated ethenes and 1,2-dichloroethane by “*Dehalococcoides ethenogenes*” 195. *Appl. Environ. Microbiol.* 65: 3108-3113.
- McCarty, P.L. 1994. An overview of anaerobic transformation of chlorinated solvents. EPA Symposium on Intrinsic Bioremediation of Ground Water. Denver, CO: 135-142.
- Middledorp, P.J.M., Luijten, M.L.G.C., van de Pas, B.A., van Eedert, M.H.A., Kengen, S.W.M., Schraa, G., Stams, A.J.M. *Bioremediation J.* 3:151-169.
- Müller, J.A., B.M. Rosner, G. von Abendroth, G. Meshulam-Simon, P. L. McCarty, and A. M. Spormann. 2004. Molecular Identification of the Catabolic Vinyl Chloride Reductase from *Dehalococcoides* sp. Strain VS and Its Environmental Distribution. *Applied and Environmental Microbiology* 2004 August; 70(8): 4880–4888.
- Montgomery, J.H. 2000. *Ground water chemicals*. 3<sup>rd</sup> ed. CRC press LLC. Boca Raton, FL.
- Nielsen, B.R., Keasling, J.D. 1999. Reductive Dechlorination of Chlorinated Ethene DNAPLs by a Culture Enriched from Contaminated Groundwater. *Biotechnology and Bioengineering.*, 62(2), 160-165.
- Norris R.D., Hinchee, Brown, McCarty, Semprini, Wilson, Kambell, Reinhard, Bauwer, Borden, Vogel, Thomas, and Ward. 1994. *Handbook of Bioremediation*. Boca Raton, FL; Lewis Publishers. ID:380192 (UTL Catalogue ckey).
- Nelson, D.K., Hozalski, R.M., Clapp, L.W., Semmens, M.J., and Novak, P.J., effect of nitrate and sulfate on dechlorination by a mixed hydrogen-fed culture. *Bioremediation*. 6: 225-236
- OECD Guideline for Testing of Chemicals. Adopted by council 27<sup>th</sup> July 1995. Partition coefficient (n-octanol/water): shake flask method. <http://titania.sourceoecd.org> (Accessed September 2006)
- Prager, J.C. 1996. *Environmental Contaminant Reference Databook*. Toronto, Van Nostrand Reinhold.
- Russell, H.H., Mathews, J.G., and Sewell, G.W. 1992. TCE removal from contaminated soil and groundwater. Tech. rept. United States Environmental Protection Agency (EPA), Ada, Oklahoma. EPA/540/S-92/002.
- Sleep, B.E., and Sykes, J.F. 1993. compositional simulation of groundwater contamination by organic compounds. 1. model development and verification. *Water resources research*. 29(6), 1697-1708.
- Simmonds, A.C. 2007. Dechlorination Rates in KB-1, a Commercial Trichloroethylene-Degrading Bacterial Culture. M.A.Sc. thesis, University of Toronto, Toronto, Ont.
- Smatlak, C.R., Gossett, J.M., and Zinder, S.H. 1996. Comparative kinetics of hydrogen utilization for reductive dechlorination of tetrachloroethene and methanogenesis in an anaerobic enrichment culture. *Environ. Sci. Technol.*, 30: 2850-2858
- Smidt, H., and De Vos, W.M. 2004. Anaerobic microbial dehalogenation Annual review of Microbiology., 58: 42-73.
- Steinigeweg, S., Gmehling, J. 2002. n-butyl acetate synthesis via reactive distillation: thermodynamic aspects, reaction kinetics, pilot-plant experiments, and simulation studies. *Ind. Eng. Chem. Res.* 41:5483-5490

- Sung, Y., Ritalahti, K.M., Apkarian, R.P., and Löffler, F.E. 2006. Quantification PCR confirms purity of strain GT, a novel trichloroethene-to-ethene-respiring *Dehalococcoides* isolate. *Applied and Environmental Microbiology*. 72:1980-1987.
- The interstate Technology and Regulatory Council (ITRC) Bioremediation of Dense Nonaqueous Phase Liquids (Bio DNAPL) Team. 2005. Overview of In Situ Bioremediation of Chlorinated Ethene DNAPL Source Zones.
- U.S.EPA, 2000. Engineered approaches to in situ bioremediation of chlorinated solvents: fundamentals and field applications. Tech. rept. United States Environmental Protection Agency.
- Waller, A. S., Rosa Krajmalnik-Brown, F.E. Löffler and E.A. Edwards. 2005. Multiple Reductive-Dehalogenase-Homologous Genes Are Simultaneously Transcribed during Dechlorination by *Dehalococcoides*-Containing Cultures *Appl. Environ. Microbiol.* 71: 8257–8264.
- Widdel, F. 1988. Microbiology and ecology of sulfate and sulfur reducing bacteria. In: Zehnder, A.J.B. (Ed.), *Biology of Anaerobic Microorganisms*. Wiley, New York, pp. 469-586
- Yang, Y.; McCarty, P. L. 1998. Competition for Hydrogen within a Chlorinated Solvent Dehalogenating Anaerobic Mixed culture. *Environ. Sci. Technol.*, 32, 3591-3597.
- Yang, Y.; McCarty, P. L. 2000. Biologically enhanced dissolution of tetrachloroethene DNAPL. *Environ. Sci. Technol.*, 34, 2979-2984.
- Yang, Y.; McCarty, P. L. 2002. Comparison between donor substrates for biologically enhanced tetrachloroethene DNAPL dissolution”, *Env. Sci. Technol.*, 36(15), 3400-3404.
- Zawtocki, C., Lieberman, M.T, Border, R.C., Birk, G.M. 2004. Treatment of perchlorate and 1,1,1-trichloroethane in groundwater using edible oil substrate (EOS<sup>®</sup>). *Proceedings of Remediation of chlorinated and Recalcitrant Compounds – 4<sup>th</sup> Internat. Conf., Monterey, CA*
- Zenker, M.J., Borden, R.C., Barlaz, M.A., Lieberman, M.T., Lee, M.D. 2000. Insoluble substrates for reductive dehalogenation in permeable reactive barriers. In: Wickranayake, G.B., Gavaskar, A.R., Alleman, B.C., Magar, V.S. (Eds.), *Bioremediation of Phytoremediation of Chlorinated and Recalcitrant Compounds*. Battelle Press, Columbus, OH, pp. 7-84.

## Appendix A

### Procedure For Making Artificial Ground Water

The artificial groundwater (AGW) used in this study was made in two liter amounts and stored for up to two months in the dark at room temperature. A bicarbonate solution, trace mineral mix solution, and a resazurin solution were made as concentrated solutions and were diluted to form the final AGW solution. Mineral mix and nutrients were added to the AGW solution immediately prior to use to minimize the chances of contamination due to microbial growth.

#### **Bicarbonate Solution**

The bicarbonate solution was made at 10 times the concentration required for the final AGW solution. Each chemical listed below was weighed out separately into a weigh boat and added to approximately 750 mL of dionized water in a 1 L volumetric flask. Dionized water was added to bring the total volume up to 1 L. The solution was mixed well on a stir plate until all components were dissolved, transferred into a 1 L sterile pyrex bottle, and stored in the dark at room temperature.

<u>Chemical</u>	<u>Target amount</u>
NaHCO <sub>3</sub>	5.0 g
KCl	0.2 g
MgSO <sub>4</sub> ·7H <sub>2</sub> O	2.8 g

#### **Trace Mineral Mix**

The trace mineral mix solution was made at 500 times the concentration required for the final AGW solution. Each chemical listed below was weighed out separately into a weigh boat and added to approximately 750 mL of dionized water in a 1 L volumetric flask. A volume of 1 mL of 36 M H<sub>2</sub>SO<sub>4</sub> was added to help dissolve all components. Once dissolved dionized water was added to bring the total volume up to 1 L. The solution was transferred into a 1 L pyrex bottle, autoclaved, and stored in the dark at room temperature.

<u>Chemical</u>	<u>Target amount</u>
H <sub>3</sub> BO <sub>3</sub>	0.3 g
ZnCl <sub>2</sub>	0.1 g
Na <sub>2</sub> MoO <sub>4</sub> ·2H <sub>2</sub> O	0.1 g
NiCl <sub>2</sub> ·6H <sub>2</sub> O	0.75 g
MnCl <sub>2</sub> ·4H <sub>2</sub> O	1.0 g
CuCl <sub>2</sub> ·2H <sub>2</sub> O	0.1 g
CoCl <sub>2</sub> ·6H <sub>2</sub> O	1.5 g
Na <sub>2</sub> SeO <sub>3</sub>	0.02 g
Al <sub>2</sub> (SO <sub>4</sub> ) <sub>3</sub> ·18H <sub>2</sub> O	0.1 g

### **Resazurin**

The resazurin solution was made at 1000 times the concentration required for the final AGW solution. A target mass of one gram of resazurin was weighed into a weigh boat and added to approximately 750 mL of dionized water in a 1 L volumetric flask. Dionized water was added to bring the total volume up to 1 L. The solution was transferred into a 1 L pyrex bottle and stored in the dark at room temperature.

### **Preparing the Artificial Ground Water Solution**

A target mass of 2.78 grams of calcium sulfate dihydrate was weighed into a weigh boat and added to approximately 1,600 mL of dionized water in a 2 L volumetric flask. The solution was heated and mixed on a stir plate until dissolved. The solution was allowed to cool to room temperature. A volume of 200 mL of the concentrated bicarbonate solution and 2 mL of the concentrated resazurin solution were added to the 2 L volumetric flask and dionized water was used to bring the total volume up to 2 L. The solution was mixed, transferred to a sterile 2 L Pyrex bottle, and stored in the dark at room temperature.

Immediately prior to use 4 mL of the concentrated trace mineral mix was added to the AGW and a target mass of 0.05 grams of diammonium phosphate was weighed into a weigh boat and added to the AGW. The AGW was mixed well and was ready for use.

## Appendix B

### Electron Donor Demand Calculations

The theoretical electron donor demand was calculated for the nBA and SRS™ columns. The SRS™ electron donor demand was used to determine how much SRS™ to add to the SRS™ column. The nBA electron donor demand was calculated to compare the actual nBA added with the theoretical demand. Calculations used are outlined below.

#### nBA

##### nBA Electron Donor Calculations

Calculation to determine the amount of sulfate added to the nBA column over one year of column operation:

Sulfate concentration in AGW = 1,000 mg/L

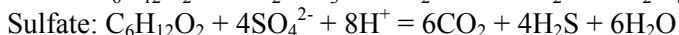
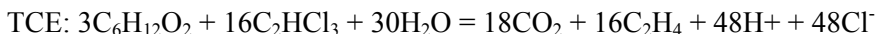
Experimental flow rate = 0.07 ml/min

Time = 365 days

Sulfate added (mg/column) =

$$1,000 \text{ mg/L} \times (0.07 \text{ mL/min} / 1000 \text{ mL/L}) \times 365 \text{ days} \times 24 \text{ hours/day} \times 60 \text{ min/hour} = 36,792$$

Balanced chemical equations for electron acceptor with nBA to determine molar ratio:



Calculation for the nBA demand needed to reduce the amount of sulfate added:

$$\text{nBA demand (mg/column)} = 36,792 \text{ (mg sulfate/column)} / 96 \text{ mg/mmol} \times \frac{1}{4} \times 116.16 \text{ mg/mmol} \\ = 11,130$$

Calculation for the nBA demand needed to reduce the amount of TCE added:

$$\text{nBA demand (mg/column)} = 59,000 \text{ (mg TCE/column)} / 131.4 \text{ mg/mmol} \times \frac{3}{16} \times 116.16 \\ \text{mg/mmol} = 9,779$$

**Table B1:** Summary of nBA demand calculations

	Sulfate	TCE
Initial Concentration (mg/columnn)	36,792 <sup>a</sup>	59,000
Molecular Weight (mg/mmol)	96	131.4
Molar Ratio <sup>b</sup>	1/4	3/16
nBA Demand (mg/columnn)	11,130	9,779
Total nBA Demand (mg/columnn)	20,909	
First nBA addition (mg/columnn)	5,400	
Second nBA addition (mg/columnn)	4,900	
Third nBA addition (mg/columnn)	10,900	
Fourth nBA addition (mg/columnn)	5,300	
Total nBA added (mg/columnn)	26,500 <sup>c</sup>	
Safety Factor	1.3	

Notes: <sup>a</sup>based on the amount of sulfate added to the column for one year of operation at the experimental flow rate

<sup>b</sup>Molar ratio is the number of moles of electron donor consumed per mole of electron acceptor

<sup>c</sup>sum of nBA amended to the column over the operational period

### nBA Amendment Solutions

Procedure used to make nBA saturated stocks for first and second nBA amendments:

nBA solubility = 6,000 mg/L

nBA density = 0.882 mg/ $\mu$ L

The amount of nBA required to make an 800 mL nBA saturated stock in AGW is

Mass nBA needed (mg) = 6,000 mg/L X 0.8L = 4,800 mg

Volume nBA needed ( $\mu$ L) = 4,800 mg / 0.882 mg/ $\mu$ L = 5,442  $\mu$ L

A one liter glass Pyrex bottle was filled with 800 mL of AGW. A volume of 100 mL of nBA was poured into the AGW. This amount of nBA was well above the theoretical amount of nBA needed. This procedure was used so that there was a visible phase of nBA above the AGW to ensure that the saturated stock concentration was maintained throughout the addition phase. The bottle was capped and the solution was shaken well and allowed to equilibrate over night prior to use.

Procedure used to make the 2,000 mg/L nBA solutions for the third nBA amendment phase:

The solution was made in a 10L tedlar bag containing 6L of AGW.

Mass nBA needed (mg) = 2,000 mg/L X 6.0 L = 12,000 mg

Volume nBA needed ( $\mu$ L) = 12,000 mg / 0.882 mg/ $\mu$ L = 14,000  $\mu$ L



A 25 mL syringe was used to inject the 14 mL of nBA through the inlet septa on the tedlar bag. The solution was shaken well and allowed to equilibrate over night prior to use.

Procedure used for the fourth (neat nBA) addition phase:

It was decided to add 6 mL of nBA because this was similar to the mass of nBA added during the first and second nBA additions.

Mass nBA added (mg) = 6 mL X 0.882 mg/ $\mu$ L X 1000 mL/ $\mu$ L = 5,300 mg

A 10 mL syringe was used to inject the nBA directly into the column inlet. The nBA was injected at a rate of 1 mL/min.

### SRS<sup>TM</sup>

#### SRS<sup>TM</sup> Electron Donor Calculations

Calculation to determine the amount of sulfate added to the SRS<sup>TM</sup> column over one year of column operation:

Sulfate concentration in AGW = 1,000 mg/L

Experimental flow rate = 0.07 ml/min

Time = 365 days

Sulfate added (mg/column) =

1,000 mg/L X (0.07mL/min/1000mL/L) X 365days X 24hours/day X 60 min/hour = 36,792

SRS is 60% soybean oil. Soybean oil is composed of:

53.3% Linoleic Acid

23.8% Oleic Acid

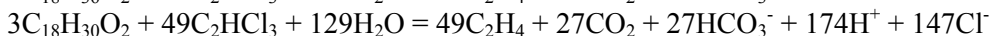
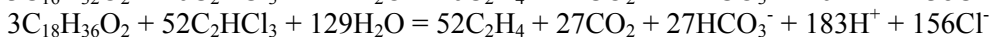
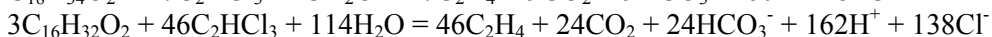
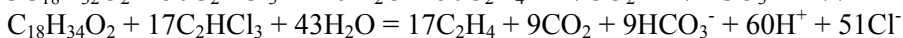
10.8% Palmitic Acid

4.0% Stearic Acid

7.1% Gamma-Linoleic

Balanced chemical equations for electron acceptor with SRS<sup>TM</sup> to determine molar ratio:

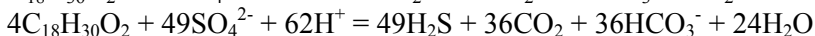
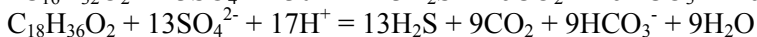
TCE:



Actual molar ratio used is the sum of the molar ratio of each component of soybean oil multiplied by the percentage of that component in soybean oil as follows:

$$(0.533 \times 3/50) + (0.238 \times 1/17) + (0.108 \times 3/46) + (0.04 \times 3/52) + (0.071 \times 3/49) = 4/67$$

Sulfate:



Actual molar ratio used is the sum of the molar ratio of each component of soybean oil multiplied by the percentage of that component in soybean oil as follows:

$$(0.533 \times 2/25) + (0.238 \times 4/51) + (0.108 \times 2/23) + (0.04 \times 1/13) + (0.071 \times 4/49) = 7/88$$

Calculation for the SRS<sup>TM</sup> demand needed to reduce the amount of sulfate added:

Molecular weight used is that of linoleic acid because linoleic acid is the main component in soy bean oil.

$$\text{SRS}^{\text{TM}} \text{ demand (mg/column)} = 36,792 \text{ (mg sulfate/column)} / 96 \text{ mg/mmol} \times 7/88 \times 280 \text{ mg/mmol} = 8,536$$

Calculation for the SRS<sup>TM</sup> demand needed to reduce the amount of TCE added

$$\text{SRS}^{\text{TM}} \text{ demand (mg/column)} = 24,000 \text{ (mg TCE/column)} / 131.4 \text{ mg/mmol} \times 4/67 \times 280 \text{ mg/mmol} = 3,053$$

**Table B2:** Summary of SRS<sup>TM</sup> demand calculations

	Sulfate	TCE
Initial Concentration (mg/column)	36792 <sup>a</sup>	24000
Molecular Weight (mg/mmol)	96	131.4
Molar Ratio <sup>b</sup>	7/88	4/67
Soybean Oil Demand (mg/column)	8536	3053
Total Soybean Oil Demand (mg/column)	11589	
Total SRS Demand (mg/column)	19315 <sup>c</sup>	
Total SRS added (mg/column)	42200 <sup>d</sup>	
Safety Factor	2.2	

Notes: <sup>a</sup>based on the amount of sulfate added to the column for one year of operation at the experimental flow rate

<sup>b</sup>Molar ratio is the number of moles of electron donor consumed per mole of electron acceptor

<sup>c</sup> soybean oil demand multiplied by 60% because SRS<sup>TM</sup> is 60% soybean oil

<sup>d</sup> Amount of SRS<sup>TM</sup> added to the column

### SRS™ Amendment Solution

SRS™ was added to the column as a 5% (w/v) solution based on the results of the batch test described in Section 4.1.4.

5% = 5 parts per 100

Mass SRS (g) = 50g/L X 1L = 50g

A one liter glass Pyrex bottle was partially filled with AGW. 50 g of SRS™ was weighed on a weigh boat and added to the AGW. A syringe with AGW was used to rinse the SRS™ into the bottle. AGW was added to the bottle to bring the total volume of the solution to 1 L. The bottle was capped and the solution was shaken well and allowed to equilibrate over night prior to use. During addition the solution was continually stirred using a magnetic stir bar and stir plate.

## Appendix C

### Bromide Tracer Test

The tracer tests were conducted using AGW amended with sodium bromide to a measured concentration of 91 mg/L. The solution was pumped through all three columns at a target flow rate of 0.07 mL/min. Flow rate measurements were taken throughout the tracer tests and are shown in Table C1. The breakthrough data were fit using the COMPSIM fate and transport model by Kokkinaki (2007) to estimate porosity and dispersivity values. Also, calculations for velocity and residence time were performed for all three columns.

Sample Calculations for velocity and residence time using the SRS<sup>TM</sup> column:

#### Pore water velocity

$$V = Q/(nXA)$$

$$\text{Average } Q = 0.079 \text{ cm}^3/\text{min} / 60 \text{ sec}/\text{min}/1000/1000 = 1.32 \times 10^{-9} \text{ m}^3/\text{sec}$$

$$n = 0.31 \text{ estimated from COMPSIM}$$

$$\text{column radius (r)} = 3.8 \text{ cm} = 0.038 \text{ m}$$

$$A = \pi r^2 = 0.0045 \text{ m}^2$$

$$v = 1.32 \times 10^{-9} \text{ m}^3/\text{sec} / (0.31 \times 0.0045 \text{ m}^2) = 9.43 \times 10^{-7} \text{ m}/\text{sec} \times 60 \text{ sec}/\text{min} \times 60 \text{ min}/\text{hour} \times 24 \text{ hour}/\text{day} \times 100 \text{ cm}/\text{m} = 8.2 \text{ cm}/\text{day}$$

#### Residence Time

$$\tau = n \times (V/Q)$$

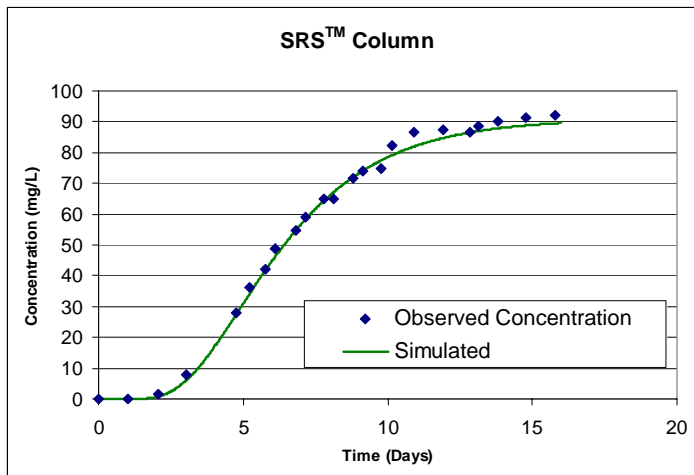
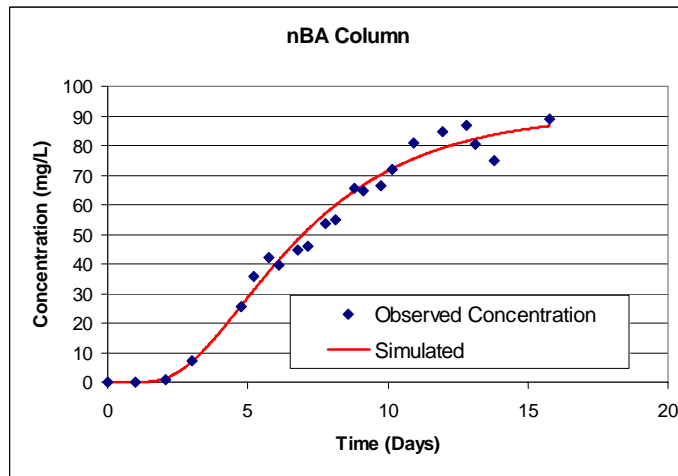
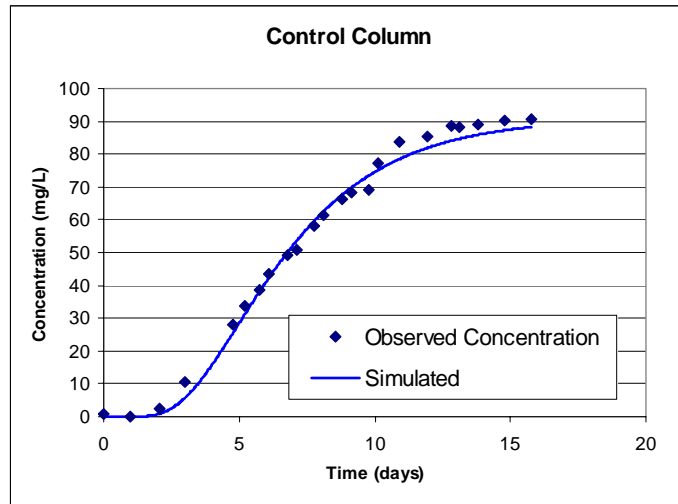
$$V = A \times L$$

$$V = 0.0045 \text{ m}^2 \times 0.6 \text{ m} = 0.0027 \text{ m}^3$$

$$\tau = 0.31 \times 0.0027 \text{ m}^3 / 1.32 \times 10^{-9} \text{ m}^3/\text{sec} / 60 \text{ sec}/\text{min} / 60 \text{ min}/\text{hour} / 24 \text{ hour}/\text{day} = 7.45 \text{ days}$$

**Table C1:** Bromide Tracer Test Data

Control Column				nBA Column				SRS™ Column			
Time (Days)	Br. Conc. (mg/L)	Simulated Br. Conc. (mg/L)	flow rate (cm <sup>3</sup> /min)	Time (Days)	Br. Conc. (mg/L)	Simulated Br. Conc. (mg/L)	flow rate (cm <sup>3</sup> /min)	Time (Days)	Br. Conc. (mg/L)	Simulated Br. Conc. (mg/L)	flow rate (cm <sup>3</sup> /min)
0.00	0.67	0.00	0.09	0.00	<0.39	0.00	0.09	0.00	<0.39	0.00	0.10
1.00	<0.39	0.00	0.07	1.00	<0.39	0.01	0.06	1.00	<0.39	0.00	0.07
2.06	2.5	0.91	0.08	2.06	0.83	1.4	0.09	2.06	1.7	0.80	--
3.02	11	5.9	0.07	3.02	7.4	7.1	0.07	3.02	7.8	6.0	0.08
4.76	28	25	0.07	4.76	26	26	0.07	4.76	28	28	0.09
5.23	34	32	0.07	5.23	36	31	0.07	5.23	36	34	0.08
5.77	38	38	0.08	5.77	42	38	0.07	5.77	42	42	0.07
6.09	44	42	0.07	6.09	40	41	0.07	6.09	49	46	0.08
6.81	49	50	0.07	6.81	45	48	0.07	6.81	55	55	0.07
7.14	51	54	0.08	7.14	46	52	0.08	7.14	59	58	0.08
7.78	58	60	0.07	7.78	54	57	0.07	7.78	65	64	0.08
8.11	61	63	0.08	8.11	55	60	0.08	8.11	65	67	0.08
8.81	66	68	0.07	8.81	65	65	0.07	8.81	72	72	0.08
9.14	68	70	0.08	9.14	65	67	0.08	9.14	74	74	0.09
9.77	69	73	0.07	9.77	66	70	0.07	9.77	75	77	0.08
10.14	77	75	0.07	10.14	72	72	0.07	10.14	82	79	0.08
10.91	84	79	--	10.91	81	75	--	10.91	87	82	--
11.93	85	82	--	11.93	85	79	--	11.93	88	85	--
12.83	88	84	0.07	12.83	87	82	0.07	12.83	87	87	0.08
13.13	88	85	0.09	13.13	80	82	0.08	13.13	89	87	0.08
13.80	89	86	0.07	13.80	75	84	0.07	13.80	90	88	0.08
14.77	90	87	0.07	15.78	89	87	0.07	14.77	91	89	0.08
15.78	90	88	0.07	--	--	--	--	15.78	92	90	0.08



**Figure C1:** Bromide Breakthrough Curves for control, nBA and SRS™ Columns

## Appendix D

### Mass Balance Calculations

In order to perform electron donor, chlorinated ethene and ethene, and chloride mass balance calculations the mass of chlorinated ethenes, ethene, methane, sulfate, chloride, and acetate removed from the column were needed. After each sampling event the cumulative masses of these compounds exiting the column were calculated. The following example shows the method used to determine the cumulative mass of a compound removed from a column after day 50. Day 50 was chosen randomly. The same method was used for all other mass removal calculations. Refer to Appendix E for the raw data sheets.

#### Example Calculation for mass removed from a column:

Mass removed (mg) = (average of current and previous compound concentration [mg/L]) X (current day – previous day [day]) X (average of current and previous flow rate [ml/min]) / (1000 [ml/L] X 60 [min/hour] X 24 [hours/day]) + (previous day cumulative compound mass [mg])

For TCE in the SRS column after 50 days of column operation:

Mass TCE removed (mg) = (average of 670 mg/L and 689 mg/L) X (50 days – 49 days) X (average of 0.062 ml/min and 0.046 ml/min) / 1000 ml/L X 60 min/hour X 24 hour/day + 2,082 mg  
Mass TCE removed (mg) = (679 mg/L) X (1 day) X (0.054 ml/min) / (1000 ml/L X 60 min/hour X 24 hour/day) + 2,082 mg

Mass TCE removed (mg) = 2,134 mg

#### Chlorinated Ethene Consumption:

In order to determine the amount of electron donor used, the mass of chlorinated ethenes consumed was needed. Using the reductive dechlorination sequence the mass of TCE consumed is equal to the mass of *cis*-DCE, VC, and ethene produced. Similarly the mass of *cis*-DCE consumed is equal to the mass of VC and ethene produced and the mass of VC consumed is equal to the mass of ethene produced. Using the masses of chlorinated ethenes and ethenes exiting the column after 572 days (refer to Appendix E):

TCE consumed = 5,178 mg + 1,259 mg + 468 mg = 6,906 mg

*cis*-DCE consumed = 1,259 mg + 468 mg = 1,729 mg

VC consumed = 468 mg

#### Methane Production:

Amount methane produced is equal to the cumulative mass of methane removed from the column = 240 mg

### Sulfate Reduction:

The mass of sulfate pumped through the column from day 89 to 572 was calculated. This time frame was chosen because before day 89 there was no measurable sulfate reduction and the sulfate exiting the column was assumed to have not consumed any electron donor.

Sulfate concentration in AGW = 1,000 mg/L

Experimental flow rate = 0.07 ml/min

Time = 572 – 89 = 483 days

Sulfate added (mg/column) =

1,000 mg/L X (0.07mL/min/1000mL/L) X 483 days X 24hours/day X 60 min/hour = 48,686

Mass of sulfate removed from day 89 to 572 is equal to the cumulative mass of sulfate exiting the column over this time frame = 8,190 mg

The amount of sulfate reduced is equal to the amount of sulfate pumped through the column minus the mass of sulfate removed in the effluent.

Mass sulfate reduced = 48,686 mg – 8,190 mg = 40,496 mg

### Acetate production:

Amount acetate produced is equal to the cumulative mass of acetate removed from the column = 19,472 mg

### Calculation for the amount of SRS<sup>TM</sup> consumed during reductive dechlorination, sulfate reduction and methane production:

The calculation shown here is for the entire operational period (up to day 572). The same method was used to calculate the amount of SRS<sup>TM</sup> consumed over the shorter time period up to day 404.

Balanced chemical equations for chlorinated ethenes, sulfate and methane with hydrogen were used to determine the molar ratio, hydrogen demand, and the amount of hydrogen consumed for each reaction (Table D1).

The initial amount of SRS<sup>TM</sup> added was = 42,200 mg

SRS<sup>TM</sup> is composed of 60% soybean oil and soybean oil is assumed to be 100% linoleic acid for the purposes of this calculation. Therefore the amount of linoleic acid added was = 42,200 mg X 60% = 25,320 mg



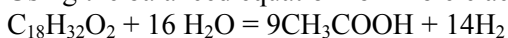
**Table D1:** Summary of Hydrogen Consumption Calculations

Reaction	Balanced Equation	Amount (mg)	Amount (mmol)	H <sub>2</sub> demand (mmol/mmol)	H <sub>2</sub> consumed (mmol)
TCE to cis-DCE	$C_2HCl_3 + H_2 = C_2H_2Cl_2 + H^+ + Cl^-$	6,906	53	1	53
cis-DCE to VC	$C_2H_2Cl_2 + H_2 = C_2H_3Cl + H^+ + Cl^-$	1,729	18	1	18
VC to ethene	$C_2H_3Cl + H_2 = C_2H_4 + H^+ + Cl^-$	468	7.5	1	7.5
Sulfate reduction	$SO_4^{2-} + 4H_2 + H^+ = HS^- + 4H_2O$	40,496	422	4	1,687
Methane production	$CO_2 + 4H_2 = CH_4 + 2H_2O$	240	15	4	60
Total hydrogen consumed					1,823

First the mmoles of linoleic acid consumed was calculated from the amount of acetate produced.

mmoles acetate produced = 19,472 mg / 59mg/mmmole = 330 mmoles.

Using the balanced equation for linoleic acid fermentation to acetate.



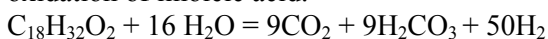
mmoles of linoleic acid consumed = 330 / 9 = 36.67 mmoles

and the mmoles of hydrogen produced = 36.67 X 14 = 514

The amount of hydrogen consumed is greater then the amount of hydrogen produced.

1,826-513 = 1,312 mmoles

Therefore the missing amount of hydrogen (1,312 mmoles) is assumed to come from complete oxidation of linoleic acid.



Hydrogen produced from complete oxidation of linoleic acid = 1,312 / 50 = 26 mmoles

The total mmoles of linoleic acid consumed is equal to the mmoles of linoleic acid consumed from acetate production plus the mmoles of linoleic acid consumed from complete oxidation = 37 +26 = 63 mmoles

The mass of linoleic acid consumed = 63 mmoles X 280 mg/mmol = 17,616 mg

The percent of linoleic acid consumed = 17,616 / 25,320 = 70%



## Anion Data for Control Column

Date	Day	Sample ID	Fow Rate (mL/Min)	Chloride (mg/L)	Chloride (µM)	Sum Chloride (µM)	Nitrite-N (mg/L)	Nitrate-N (mg/L)	Sulfate (mg/L)	Sulfate (µM)	Bromide (mg/L)	Phosphate (mg/L)
7/7/2006 14:10	0	7-July-06-2	0.00	--	--	--	--	--	--	--	--	--
7/10/2006 14:10	3	10-July-06-2	0.11	20	563	256	0.00	0.00	1146	12060	0.00	0.00
7/13/2006 9:00	6	13-July-06-2	0.11	22	620	507	0.00	0.00	848	8924	0.00	0.00
7/18/2006 14:10	11	18-July-06-2	0.09	22	627	915	0.00	0.00	914	9625	0.00	0.00
7/24/2006 15:15	17	24-July-06-2	0.11	21	592	1480	0.00	0.00	951	10016	0.00	0.00
7/31/2006 16:15	24	31-July-06-2	0.10	22	614	2101	0.00	0.00	994	10467	0.00	0.00
8/8/2006 16:30	32	8-August-06-2	0.09	21	580	2704	0.00	0.00	971	10219	0.00	0.00
8/14/2006 13:45	38	14-August-06-2	0.08	19	537	3094	0.00	0.00	926	9747	0.00	0.00
8/21/2006 14:45	45	21-August-06-2	0.07	16	455	3461	0.00	0.00	897	9444	0.00	0.00
8/28/2006 15:30	52	28-August-06-2	0.09	58	1634	4398	0.00	0.00	1012	10653	0.00	0.00
9/5/2006 14:15	60	5-Sept-06-2	0.09	15	414	5407	0.00	0.00	1029	10829	0.00	0.00
9/11/2006 10:15	66	11-Sept-06-2	0.08	13	372	5668	0.00	0.00	1004	10571	0.00	0.00
9/18/2006 10:30	73	18-Sept-06-2	0.14	11	317	6155	0.00	0.00	887	9333	0.00	0.00
9/25/2006 13:00	80	25-Sept-06-2	0.08	11	305	6415	0.00	0.00	901	9480	0.00	0.00
9/29/2006 14:00	84	29-Sept-06-2	0.09	11	322	6581	0.00	0.00	879	9250	0.00	0.00
10/2/2006 11:00	87	02-Oct-06-2	0.09	15	429	6721	0.00	0.00	889	9353	0.00	0.00
10/10/2006 11:40	95	10-Oct-06-2	0.09	16	445	7190	0.00	0.00	956	10061	0.00	0.00
10/16/2006 13:30	101	16-Oct-06-2	0.09	21	578	7603	0.00	0.00	1121	11796	0.00	0.00
10/30/2006 10:15	115	30-Oct-06-2	0.08	17	484	8445	0.00	0.00	1123	11825	0.00	0.00
11/6/2006 15:00	122	06-Nov-06-2	0.07	13	353	8760	0.00	0.00	850	8947	0.00	0.00
11/13/2006 11:55	129	13-Nov-06-2	0.08	13	367	9041	0.00	0.00	939	9880	0.00	0.00
11/20/2006 14:15	136	20-Nov-06-2	0.07	14	389	9321	0.00	0.00	949	9980	0.00	0.00
11/27/2006 14:30	143	27-Nov-06-2	0.07	13	375	9608	0.00	0.00	955	10055	0.00	0.00
12/4/2006 12:30	150	04-DEC-06-2	0.08	14	389	9916	0.00	0.00	968	10192	0.00	0.00
12/11/2006 15:30	157	11-DEC-06-2	0.08	13	358	10228	0.00	0.00	963	10140	0.00	0.00
12/18/2006 14:15	164	18-DEC-06-2	0.07	16	445	10523	0.00	0.00	1117	11759	0.00	0.00
12/27/2006 11:00	173	27-DEC-06-2	0.08	12	338	10939	0.00	0.00	858	9036	0.00	0.00
1/2/2007 11:30	179	2-JAN-07-2	0.08	11	300	11171	0.00	0.00	759	7992	0.00	0.00
1/8/2007 15:45	185	8-JAN-07-2	0.09	11	310	11418	0.00	0.00	809	8515	0.00	0.00
1/15/2007 14:30	192	15-JAN-07-2	0.09	11	312	11685	0.00	0.00	797	8393	0.00	0.00
1/22/2007 14:30	199	22-JAN-07-2	0.11	11	312	12031	0.00	0.00	798	8404	0.00	0.00
1/29/2007 11:00	206	29-JAN-07-2	0.09	18	508	12392	0.00	0.00	1082	11393	0.00	0.00
2/5/2007 14:00	213	5-feb-07-2	0.09	22	609	12927	0.00	0.00	1093	11506	0.00	0.00
2/12/2007 16:00	220	12-feb-07-2	0.09	15	426	13408	0.00	0.00	583	6140	0.00	0.00
2/19/2007 16:30	227	19-feb-07-2	0.10	17	474	13849	0.00	0.00	922	9705	0.00	0.00
2/26/2007 12:00	234	26-feb-07-2	0.09	18	513	14289	0.00	0.00	912	9604	0.00	0.00
3/5/2007 11:30	241	5-March-07-2	0.09	21	594	14812	0.00	0.00	933	9823	0.00	0.00
3/12/2007 15:30	248	12-March-07-2	0.09	11	324	15216	0.00	0.00	998	10507	0.00	0.00
3/19/2007 16:30	255	19-March-07-2	0.09	17	478	15583	0.00	0.00	791	8331	0.00	0.00
3/26/2007 16:00	262	26-March-07-2	0.11	17	484	16118	0.00	0.00	792	8334	0.00	0.00
3/29/2007 16:00	265	29-March-07-2	0.09	19	523	16304	0.00	0.00	801	8431	0.00	0.00
4/2/2007 17:15	269	2-April-07-2	0.06	24	677	16503	0.00	0.00	802	8443	0.00	0.00
4/9/2007 16:50	276	9-April-07-2	0.04	16	464	16745	0.00	0.00	821	8637	0.00	0.00
4/16/2007 16:00	283	16-April-07-2	0.03	16	458	16882	0.00	0.00	792	8339	0.00	0.00
4/23/2007 16:00	290	23-April-07-2	0.07	18	517	17224	0.00	0.00	828	8719	0.00	0.00
4/30/2007 14:00	297	30-April-07-2	0.06	18	502	17536	0.00	0.00	826	8697	0.00	0.00
5/7/2007 11:00	304	7-May-07-2	0.07	18	516	17898	0.00	0.00	958	10083	0.00	0.00
5/14/2007 9:45	311	14-May-07-2	0.06	26	728	18292	0.00	0.00	955	10054	0.00	0.00
5/22/2007 11:30	319	22-May-07-2	0.07	17	467	18813	0.00	0.00	981	10331	0.00	0.00
5/28/2007 11:00	325	28-May-07-2	0.07	17	481	19098	0.00	0.00	982	10336	0.00	0.00
6/4/2007 15:00	332	04-June-07-2	0.06	16	448	19398	0.00	0.00	991	10431	0.00	0.00
6/11/2007 12:00	339	11-June-07-2	0.05	16	450	19618	0.00	0.00	973	10246	0.00	0.00
6/18/2007 12:00	346	18-June-07-2	0.07	13	364	19889	0.00	0.00	943	9928	0.00	0.00
6/25/2007 13:00	353	25-June-07-2	0.05	13	371	20091	0.00	0.00	816	8590	0.00	0.00
7/2/2007 11:00	360	2-July-07-2	0.21	14	381	20860	0.00	0.00	963	10136	0.00	0.00
7/9/2007 15:30	367	9-July-07-2	0.05	12	330	21058	0.00	0.00	791	8321	0.00	0.00
7/16/2007 10:00	374	16-July-07-2	0.04	13	377	21209	0.00	0.00	973	10247	0.00	0.00
7/23/2007 14:00	381	23-July-07-2	0.07	14	385	21470	0.00	0.00	980	10312	0.00	0.00
7/30/2007 10:00	388	30-July-07-2	0.06	16	437	21718	0.00	0.00	890	9372	0.00	0.00
8/6/2007 14:30	395	6-Aug-07-2	0.07	13	361	21996	0.00	0.00	908	9560	0.00	0.00
8/13/2007 12:00	402	13-Aug-07-2	0.05	22	633	22249	0.00	0.00	905	9523	0.00	0.00
8/20/2007 17:00	409	20-Aug-07-2	0.08	19	522	22699	0.00	0.00	990	10419	0.00	0.00
8/27/2007 16:00	416	27-Aug-07-2	0.07	17	473	23023	0.00	0.00	973	10240	0.00	0.00
9/10/2007 16:00	430	10-sep-07-2	0.08	15	416	23733	0.00	0.00	942	9915	0.00	0.00
9/17/2007 16:00	437	17-sep-07-2	0.08	15	420	24061	0.00	0.00	957	10069	0.00	0.00
9/24/2007 16:00	444	24-sep-07-2	0.07	15	415	24341	0.00	0.00	957	10078	0.00	0.00
10/1/2007 16:00	451	1-oct-07-2	0.06	14	405	24579	0.00	0.00	940	9891	0.00	0.00
10/9/2007 11:30	459	9-oct-07-2	0.05	16	444	24832	0.00	0.00	978	10299	0.00	0.00
10/15/2007 15:00	465	15-oct-07-2	0.07	13	375	25081	0.00	0.00	959	10095	0.00	0.00
10/22/2007 14:00	472	22-oct-07-2	0.05	14	406	25286	0.00	0.00	948	9983	0.00	0.00



Anion Data for nBA Column

Date	Day	Sample ID	Flow Rate (mL/Min)	Chloride (mg/L)	Chloride (µM)	Sum Chloride (µM)	Nitrite-N (mg/L)	Nitrate-N (mg/L)	Sulfate (mg/L)	Sulfate (µM)	Bromide (mg/L)	Phosphate (mg/L)
3/21/2006 15:15	0	21-Mar-06-1	0.080									
3/22/2006 14:40	1	22-Mar-06-1	0.083	20	571	68	0.00	0.04	682	7177	46	0.00
3/24/2006 12:00	3	24-Mar-06-1	0.081	24	889	205	0.00	0.04	597	6280	51	0.00
3/27/2006 11:00	6	27-Mar-06-1	0.075	29	818	446	0.00	0.04	0.00	0.00	26	0.00
3/29/2006 16:00	8	29-Mar-06-1	0.082	40	1131	699	0.00	0.06	812	8546	7.5	0.00
3/31/2006 14:00	10	31-Mar-06-1	0.084	40	1129	961	0.00	0.00	767	8072	0.9	0.00
4/3/2006 16:00	13	03-Apr-06-1	0.091	68	1907	1573	0.00	0.00	704	7409	2.2	0.00
4/6/2006 10:00	16	06-Apr-06-1	0.089	49	1393	2185	0.00	0.00	838	8822	2.4	0.00
4/10/2006 14:00	20	10-Apr-06-1	0.076	43	1198	2742	0.00	0.00	822	8652	0.00	0.00
4/12/2006 16:00	22	12-Apr-06-1	0.068	45	1274	2994	0.00	0.00	821	8639	0.00	0.00
4/19/2006 16:00	29	19-Apr-06-1	0.075	60	1681	4116	0.00	0.00	959	10093	0.00	0.00
5/1/2006 9:00	36	1-May-06-1	0.078	42	1172	5190	0.00	0.00	970	10215	0.00	0.00
5/5/2006 11:00	40	5-May-06-1	0.073	30	855	5628	0.00	0.00	999	10515	0.00	0.00
5/10/2006 9:00	45	10-May-06-1	0.080	25	710	6071	0.00	0.00	1007	10599	0.00	0.00
5/12/2006 14:30	47	12-May-06-1	0.079	25	707	6242	0.00	0.00	1041	10959	0.00	0.00
5/14/2006 14:30	49	14-May-06-1	0.067	27	760	6382	0.00	0.00	1006	10585	0.00	0.00
5/15/2006 14:30	50	15-May-06-1	0.075	28	786	6465	0.00	0.00	1020	10740	0.00	0.00
5/16/2006 15:00	51	16-May-06-1	0.082	25	699	6555	0.00	0.00	1034	10886	4.6	0.00
5/17/2006 14:00	52	17-May-06-1	0.085	22	621	6632	0.00	0.00	990	10420	33	0.00
5/19/2006 14:00	53	18-May-06-1	0.089	22	612	6711	0.00	0.00	1016	10696	87	0.00
5/19/2006 13:00	54	19-May-06-1	0.083	21	578	6814	0.00	0.00	1002	10552	87	0.00
5/24/2006 11:00	56	24-May-06-1	0.073	25	705	6909	0.00	0.00	983	10345	104	0.00
5/25/2006 11:00	57	25-May-06-1	0.075	24	686	6984	0.00	0.00	951	10015	103	0.00
5/26/2006 13:30	58	26-May-06-1	0.078	23	653	7067	0.00	0.00	949	9987	113	0.00
5/29/2006 12:00	61	29-May-06-1	0.081	25	712	7300	0.00	0.00	950	10004	79	0.00
5/31/2006 14:00	63	31-May-06-1	0.084	30	857	7497	0.00	0.00	972	10237	34	0.00
6/2/2006 11:30	65	02-June-06-1	0.081	36	1023	7705	0.00	0.00	1038	10923	1.0	0.00
6/5/2006 13:15	68	05-June-06-1	0.082	27	759	8028	0.00	0.00	1065	11208	1.9	0.00
6/7/2006 11:30	70	07-June-06-1	0.082	20	568	8178	0.00	0.00	964	10149	0.00	0.00
6/9/2006 14:00	72	09-June-06-1	0.083	20	574	8322	0.00	0.00	984	10358	0.00	0.00
6/12/2006 10:30	75	12-June-06-1	0.086	20	560	8522	0.00	0.00	998	10510	0.00	0.00
6/15/2006 14:20	78	15-June-06-1	0.090	19	541	8748	0.00	0.00	991	10432	0.00	0.00
6/16/2006 14:45	79	16-June-06-1	0.082	20	553	8814	0.00	0.00	1002	10552	0.00	0.00
6/19/2006 10:40	82	19-June-06-1	0.083	22	619	9012	0.00	0.00	1019	10731	0.00	0.00
6/21/2006 11:00	84	21-June-06-1	0.079	38	1066	9204	0.00	0.00	974	10248	0.00	0.00
6/23/2006 15:00	86	23-June-06-1	0.080	38	1078	9472	0.00	0.00	1026	10798	0.00	0.00
6/26/2006 14:00	89	26-June-06-1	0.085	28	795	9810	0.00	0.00	1014	10669	0.00	0.00
6/29/2006 14:00	92	29-June-06-1	0.082	18	520	10042	0.00	0.00	898	9451	0.00	0.00
7/4/2006 13:30	97	04-July-06-1	0.087	17	483	10354	0.00	0.00	871	9167	0.00	0.00
7/6/2006 13:30	99	06-July-06-1	0.085	19	537	10478	0.00	0.00	903	9509	0.00	0.00
7/10/2006 13:40	103	10-July-06-1	0.087	18	500	10738	0.00	0.00	919	9670	0.00	0.00
7/13/2006 9:00	106	13-July-06-1	0.083	19	522	10910	0.00	0.00	932	9810	0.00	0.00
7/18/2006 13:30	111	18-July-06-1	0.081	19	524	11225	0.00	0.00	924	9721	0.00	0.00
7/24/2006 14:30	117	24-July-06-1	0.077	17	492	11567	0.00	0.00	939	9884	0.00	0.00
7/27/2006 15:00	120	27-July-06-1	0.070	18	506	11719	0.00	0.00	951	10011	0.00	0.00
7/28/2006 15:00	121	28-July-06-1	0.070	19	518	11771	0.00	0.00	961	10023	0.00	0.00
7/29/2006 13:30	122	29-July-06-1	0.070	18	514	11820	0.00	0.00	962	10130	0.00	0.00
7/30/2006 19:15	123	30-July-06-1	0.070	20	563	11887	0.00	0.00	977	10281	0.00	0.00
7/31/2006 16:00	124	31-July-06-1	0.070	18	520	11934	0.00	0.00	918	9668	0.00	0.00
8/1/2006 14:00	125	1-August-06-1	0.070	20	552	11984	0.00	0.00	927	9762	0.00	0.00
8/2/2006 15:00	126	2-August-06-1	0.070	21	581	12043	0.00	0.00	942	9914	0.00	0.00
8/3/2006 13:30	127	3-August-06-1	0.070	22	620	12100	0.00	0.00	951	9999	0.00	0.00
8/4/2006 14:00	128	4-August-06-1	0.070	27	749	12171	0.00	0.00	946	9955	0.00	0.00
8/6/2006 15:30	130	6-August-06-1	0.070	37	1054	12358	0.00	0.00	1069	11257	0.00	0.00
8/8/2006 15:00	132	8-August-06-1	0.079	40	1132	12605	0.00	0.00	974	10249	0.00	0.00
8/10/2006 14:00	134	10-August-06-1	0.070	37	1049	12800	0.00	0.00	926	9748	0.00	0.00
8/14/2006 13:45	138	14-August-06-1	0.079	27	762	13230	0.00	0.00	927	9758	0.00	0.00
8/21/2006 14:45	145	21-August-06-1	0.066	17	489	13648	0.00	0.00	916	9641	0.00	0.00
8/28/2006 15:30	152	28-August-06-1	0.080	17	478	14037	0.00	0.00	975	10258	0.00	0.00
9/5/2006 14:15	160	5-Sept-06-1	0.084	14	382	14449	0.00	0.00	985	10365	0.00	0.00
9/11/2006 10:15	166	11-Sept-06-1	0.084	12	326	14698	0.00	0.00	979	10307	0.00	0.00
9/18/2006 10:30	173	18-Sept-06-1	0.070	11	305	14922	0.00	0.00	877	9234	0.00	0.00
9/25/2006 13:00	180	25-Sept-06-1	0.070	9.5	266	15125	0.00	0.00	861	9064	0.00	0.00
9/29/2006 14:00	184	29-Sept-06-1	0.087	11	323	15275	0.00	0.00	814	8572	0.00	0.00
10/2/2006 11:00	187	02-Oct-06-1	0.089	11	309	15391	0.00	0.00	891	9618	0.00	0.00
10/10/2006 11:40	195	10-Oct-06-1	0.079	20	575	15795	0.00	0.00	916	9638	0.00	0.00
10/16/2006 13:30	201	16-Oct-06-1	0.072	21	598	16164	0.00	0.00	944	9935	0.00	0.00
10/23/2006 0:00	207	23-Oct-06-1	0.075	21	578	16575	0.00	0.00	874	9205	0.00	0.00
10/25/2006 0:00	209	25-Oct-06-1	0.070	21	579	16691	0.00	0.00	890	9364	0.92	0.00
10/26/2006 0:00	210	26-Oct-06-1	0.070	21	589	16750	0.00	0.00	893	9396	0.49	0.00
10/30/2006 10:15	215	30-Oct-06-1	0.075	20	550	17020	0.00	0.00	808	8504	0.00	0.00
11/1/2006 0:00	216	1-Nov-06-1	0.070	20	563	17108	0.00	0.00	890	9363	64	0.00
11/6/2006 15:00	222	06-Nov-06-1	0.073	59	1650	17763	0.00	0.00	740	7788	44	0.00
11/8/2006 12:00	224	08-Nov-06-1	0.070	69	1956	18103	0.00	0.00	742	7806	73	0.00
11/9/2006 0:00	224	09-Nov-06-1	0.070	70	1974	18202	0.00	0.00	650	6845	90	0.00
11/10/2006 0:00	225	10-Nov-06-1	0.070	73	2063	18406	0.00	0.00	550	5795	94	0.00
11/13/2006 11:55	229	13-Nov-06-1	0.077	94	2645	19323	0.00	0.00	335	3531	97	0.00
11/14/2006 0:00	229	14-Nov-06-1	0.070	105	2970	19465	0.00	0.00	230	2423	95	0.00
11/15/2006 0:00	230	15-Nov-06-1	0.070	104	2930	19763	0.00	0.00	173	1819	93	0.00
11/16/2006 0:00	231	16-Nov-06-1	0.070	108	3053	20064	0.00	0.00	155	1636	97	0.00
11/17/2006 0:00	232	17-Nov-06-1	0.070	111	3121	20376	0.00	0.00	131	1379	94	0.00
11/20/2006 14:15	236	20-Nov-06-1	0.071	136	3836	21658	0.00	0.00	94	993	86	0.00
11/24/2006 0:00	239	24-Nov-06-1	0.070	137	3857	22979	0.00	0.00	54	568	94	0.00
11/27/2006 14:30	243	27-Nov-06-1	0.071	149	4207	24455	0.00	0.00	41	427	93	0.00
11/29/2006 0:00	244	29-Nov-06-1	0.070	141	3980	25031	0.00	0.00	32	340	91	0.00
12/4/2006 0:00	249	04-DEC-06-1	0.062	162	4559	26935	0.00	0.00	28	293	94	0.00
12/6/2006 0:00	251	06-DEC-06-1	0.070	156	4381	27836	0.00	0.00	43	457	84	0.00
12/8/2006 0:00	253	08-DEC-06-1	0.070	147	4143	28695	0.00	0.00	231	2431	47	0.00
12/11/2006 0:00	256	11-DEC-06-1	0.076	82	2319	29751	0.00	0.00	652	6861	8.6	0.00
12/18/2006 0:00	263	18-DEC-06-1	0.064	68	1924	31113	0.00	0.00	1065	11214	0.00	0.00
12/27/2006 0:00	272	27-DEC-06-1	0.077	34	853	32543	0.00	0.00	853	8979	0.00	0.00
1/2/2007 11:30	279	2-JAN-07-1	0.0									

VFA Data for nBA Column

Date	Day	Sample ID	Flow Rate (mL/Min)	VFA Data															
				Lactate (mg/L)	Lactate (µM)	Acetate (mg/L)	Acetate (µM)	Sum Acetate (µM)	Propionate (mg/L)	Propionate (µM)	Formate (mg/L)	Formate (µM)	Butyrate (mg/L)	Butyrate (µM)	Sum Butyrate (µM)	Pyruvate (mg/L)	Pyruvate (µM)		
22-Mar	1	22-Mar-06-1	0.08	0.00	0.00	0.00	0.00	0.00	0.00	0.00	0.00	0.00	0.00	0.00	0.00	0.00	0.00	0.00	0.00
11-May	46	11-May-06-1	0.09	0.00	0.00	0.00	0.00	0.00	0.00	0.00	0.00	0.00	0.00	0.00	0.00	0.00	0.00	0.00	0.00
14-May	49	14-May-06-1	0.07	0.00	0.00	1.9	33	4.7	0.00	0.00	0.00	0.00	0.00	0.00	0.00	0.00	0.00	0.00	0.00
16-May	51	16-May-06-1	0.08	0.00	0.00	3.2	54	15	0.00	0.00	0.00	0.00	0.5	5.6	0.67	0.00	0.00	0.00	0.00
18-May	53	18-May-06-1	0.09	0.00	0.00	0.00	0.00	22	0.00	0.00	0.00	0.00	0.00	0.00	1.4	0.00	0.00	0.00	0.00
24-May	56	24-May-06-1	0.07	0.00	0.00	5.6	94	37	0.57	7.8	0.00	0.00	9.1	104	18	0.00	0.00	0.00	0.00
26-May	58	26-May-06-1	0.08	0.00	0.00	65	1108	172	1.1	15	0.00	0.00	3.8	44	34	0.00	0.00	0.00	0.00
31-May	63	31-May-06-1	0.08	0.00	0.00	0.00	0.00	506	0.00	0.00	0.00	0.00	0.00	0.00	47	0.00	0.00	0.00	0.00
5-Jun	68	05-June-06-1	0.08	0.00	0.00	33	558	670	0.00	0.00	0.00	0.00	0.00	0.00	47	0.00	0.00	0.00	0.00
9-Jun	72	09-June-06-1	0.08	0.00	0.00	35	591	945	0.00	0.00	0.00	0.00	0.00	0.00	47	0.00	0.00	0.00	0.00
15-Jun	78	15-June-06-1	0.09	0.00	0.00	35	594	1406	0.00	0.00	0.00	0.00	0.00	0.00	47	0.00	0.00	0.00	0.00
19-Jun	82	19-June-06-1	0.08	0.00	0.00	33	557	1682	0.00	0.00	0.00	0.00	0.00	0.00	47	0.00	0.00	0.00	0.00
23-Jun	86	23-June-06-1	0.08	0.00	0.00	30	509	1927	0.00	0.00	0.00	0.00	0.00	0.00	47	0.00	0.00	0.00	0.00
29-Jun	92	29-June-06-1	0.08	0.00	0.00	24	400	2248	0.00	0.00	0.00	0.00	0.00	0.00	47	0.00	0.00	0.00	0.00
6-Jul	99	06-July-06-1	0.08	0.00	0.00	16	279	2539	0.00	0.00	0.00	0.00	0.00	0.00	47	0.00	0.00	0.00	0.00
13-Jul	106	13-July-06-1	0.08	0.00	0.00	10	177	2730	0.00	0.00	0.00	0.00	0.00	0.00	47	0.00	0.00	0.00	0.00
27-Jul	120	27-July-06-1	0.07	0.00	0.00	3.8	65	2901	0.00	0.00	0.00	0.00	0.00	0.00	47	0.00	0.00	0.00	0.00
28-Jul	121	28-July-06-1	0.07	0.00	0.00	5.4	91	2909	0.00	0.00	0.00	0.00	0.00	0.00	47	0.00	0.00	0.00	0.00
29-Jul	122	29-July-06-1	0.07	0.00	0.00	19	314	2929	0.00	0.00	0.00	0.00	0.00	0.00	47	0.00	0.00	0.00	0.00
30-Jul	123	30-July-06-1	0.07	0.00	0.00	68	1159	3004	0.00	0.00	0.00	0.00	1.1	12	48	0.00	0.00	0.00	0.00
31-Jul	124	31-July-06-1	0.07	0.00	0.00	145	2458	3186	0.00	0.00	0.00	0.00	2.9	33	50	0.00	0.00	0.00	0.00
1-Aug	125	1-August-06-1	0.07	0.00	0.00	216	3658	3494	0.00	0.00	0.00	0.00	4.8	55	55	0.00	0.00	0.00	0.00
2-Aug	126	2-August-06-1	0.07	0.00	0.00	299	5073	3934	0.00	0.00	0.00	0.00	8.4	95	62	0.00	0.00	0.00	0.00
3-Aug	127	3-August-06-1	0.07	0.00	0.00	432	7323	4559	0.00	0.00	0.00	0.00	11.5	130	74	0.00	0.00	0.00	0.00
4-Aug	128	4-August-06-1	0.07	0.00	0.00	360	6103	5236	0.00	0.00	0.00	0.00	11.3	128	87	0.00	0.00	0.00	0.00
6-Aug	130	6-August-06-1	0.07	0.00	0.00	353	5986	6454	0.00	0.00	0.00	0.00	14.2	161	116	0.00	0.00	0.00	0.00
8-Aug	132	8-August-06-1	0.09	0.00	0.00	257	4353	7632	0.00	0.00	0.00	0.00	8.3	94	145	0.00	0.00	0.00	0.00
10-Aug	134	10-August-06-1	0.07	0.00	0.00	140	2378	8310	0.00	0.00	0.00	0.00	4.2	47	159	0.00	0.00	0.00	0.00
11-Aug	135	11-Aug-06-1	0.08	0.00	0.00	64	1080	8506	0.00	0.00	0.00	0.00	0.66	7.5	162	0.00	0.00	0.00	0.00
14-Aug	138	14-August-06-1	0.08	0.00	0.00	37	628	8797	0.00	0.00	0.00	0.00	0.56	6.3	164	0.00	0.00	0.00	0.00
21-Aug	145	21-August-06-1	0.07	0.00	0.00	0.00	0.00	9006	0.00	0.00	0.00	0.00	0.00	0.00	167	0.00	0.00	0.00	0.00
28-Aug	152	28-August-06-1	0.08	0.00	0.00	5.7	97	9045	0.00	0.00	0.00	0.00	0.00	0.00	167	0.00	0.00	0.00	0.00
5-Sep	160	5-Sept-06-1	0.08	0.00	0.00	1.5	25	9104	0.00	0.00	0.00	0.00	0.00	0.00	167	0.00	0.00	0.00	0.00
11-Sep	166	11-Sept-06-1	0.08	0.00	0.00	1.3	22	9121	0.00	0.00	0.00	0.00	0.00	0.00	167	0.00	0.00	0.00	0.00
18-Sep	173	18-Sept-06-1	0.07	0.00	0.00	0.00	0.00	9128	0.00	0.00	0.00	0.00	0.00	0.00	167	0.00	0.00	0.00	0.00
25-Sep	180	25-Sept-06-1	0.07	0.00	0.00	1.4	24	9137	0.00	0.00	0.00	0.00	23	0.00	167	0.00	0.00	0.00	0.00
29-Sep	184	29-Sept-06-1	0.09	2.8	31	1.2	20	9148	0.00	0.00	1.2	26	0.00	0.00	167	0.00	0.00	0.00	0.00
2-Oct	187	02-Oct-06-1	0.09	4.9	55	0.00	0.00	9152	0.00	0.00	1.8	40	0.00	0.00	167	0.00	0.00	0.00	0.00
10-Oct	195	10-Oct-06-1	0.08	3.1	34	1.4	24	9163	0.00	0.00	1.2	27	0.00	0.00	167	0.00	0.00	0.00	0.00
16-Oct	201	16-Oct-06-1	0.07	4.0	44	0.00	0.00	9170	0.00	0.00	0.91	20	0.00	0.00	167	0.00	0.00	0.00	0.00
23-Oct	208	23-Oct-06-1	0.08	3.8	42	0.00	0.00	9170	0.00	0.00	2.1	46	0.00	0.00	167	0.00	0.00	0.00	0.00
25-Oct	210	25-Oct-06-1	0.07	2.6	29	0.00	0.00	9170	0.00	0.00	1.0	21	0.00	0.00	167	0.00	0.00	0.00	0.00
26-Oct	211	26-Oct-06-1	0.07	3.0	34	0.00	0.00	9170	0.00	0.00	1.2	28	0.00	0.00	167	0.00	0.00	0.00	0.00
30-Oct	215	30-Oct-06-1	0.07	0.00	0.00	3.3	56	9182	1.5	21	4.7	105	0.00	0.00	167	0.00	0.00	0.00	0.00
1-Nov	217	1-Nov-06-1	0.07	0.00	0.00	23	398	9228	2.2	30	20	455	0.00	0.00	167	0.00	0.00	0.00	0.00
6-Nov	222	06-Nov-06-1	0.07	1.5	17	504	8545	11577	2.3	51	1.1	24	142	1607	589	0.00	0.00	0.00	0.00
8-Nov	224	08-Nov-06-1	0.07	2.3	26	726	12299	13678	3.2	43	2.0	44	158	1799	932	0.00	0.00	0.00	0.00
9-Nov	225	09-Nov-06-1	0.07	0.00	0.00	880	14910	15050	2.0	27	84	1859	190	2153	1131	0.00	0.00	0.00	0.00
10-Nov	226	10-Nov-06-1	0.07	0.00	0.00	997	16904	16653	3.9	54	94	2098	214	2430	1362	0.00	0.00	0.00	0.00
13-Nov	229	13-Nov-06-1	0.08	0.00	0.00	1185	20078	22834	5.8	80	2.4	52	265	3004	2270	0.00	0.00	0.00	0.00
14-Nov	230	14-Nov-06-1	0.07	4.4	50	1190	20164	24862	5.9	81	2.9	64	203	2299	2538	0.00	0.00	0.00	0.00
15-Nov	231	15-Nov-06-1	0.07	0.00	0.00	1171	19847	26879	5.9	81	68	1510	154	1749	2742	0.00	0.00	0.00	0.00
16-Nov	232	16-Nov-06-1	0.07	2.9	33	1163	19717	28873	5.5	76	2.7	59	148	1674	2914	0.00	0.00	0.00	0.00
17-Nov	233	17-Nov-06-1	0.07	2.3	26	1216	20606	30905	6.4	87	2.8	62	181	2050	3102	0.00	0.00	0.00	0.00
20-Nov	236	20-Nov-06-1	0.07	0.00	0.00	1344	22786	37584	7.4	101	162	3604	368	4174	4060	0.00	0.00	0.00	0.00
24-Nov	240	24-Nov-06-1	0.07	1.9	22	1430	24244	47066	8.1	111	4.0	90	484	5463	6003	0.00	0.00	0.00	0.00
29-Nov	245	29-Nov-06-1	0.07	5.6	63	1420	24073	59241	13	173	3.4	75	484	5494	8764	0.00	0.00	0.00	0.00
6-Dec	252	06-DEC-06-1	0.06	1.9	21	1400	23734	74163	7.8	106	2.9	65	381	4328	11830	0.00	0.00	0.00	0.00
8-Dec	254	08-DEC-06-1	0.07	2.2	25	1026	17384	78308	8.9	122	2.9	64	150	1699	12437	0.00	0.00	0.00	0.00
11-Dec	257	11-DEC-06-1	0.08	3.2	36	241	4077	81816	1.8	25	1.5	33	25	280	12761	2.3	0.03		
14-Dec	260	14-Dec-06-1	0.06	3.6	41	21	359	82426	0.00	0.00	4.0	89	2.6	30	12803	1.8	0.02		
18-Dec	264	18-DEC-06-1	0.06	1.8	21	5.6	94	82509	0.00	0.00	1.3	29	0.00	0.00	12809	0.00	0.00	0.00	0.00
27-Dec	273	27-DEC-06-1	0.08	2.9	32	1.4	24	82588	0.00	0.00	1.6	35	0.00	0.00	12809	0.00	0.00	0.00	0.00
2-Jan	279	2-JAN-07-1	0.08	0.00	0.00	0.00	0.00	82576	1.5	21	0.00	0.00	0.00	0.00	12809	0.00	0.00	0.00	0.00
9-Jan	286	9-JAN-07-1	0.07	0.00	0.00	24	398	82716	0.00	0.00	0.00	0.00	0.00	0.00	12809	0.00	0.00	0.00	0.00
10-Jan	287	10-JAN-07-1	0.07	2.3	26	269	4564	82966	0.00	0.00	22	488	0.00	0.00	12809	0.00	0.00	0.00	0.00
11-Jan	288	11-JAN-07-1	0.07	0.00	0.00	539	9132	83657	30	416	0.00	0.00							

VOC and DHG Data for SRS™ Column

Date	Day	Sample ID	Flow Rate (mL/Min)	Total TCE Mass (mg)	Total cis-DCE Mass (mg)	Total VC Mass (mg)	Total Ethene Mass (mg)	Total Methane Mass (mg)	Total Ethane Mass (mg)	PCE (µg/L)	TCE (µg/L)	cis-DCE (µg/L)	VC (µg/L)	Ethene (µg/L)	Ethane (µg/L)	Methane (µg/L)	Total Ethenes (µg/L)
3/21/2006 16:15	1	21-Mar-06-3	0.07	0.00	0.00	0.00	0.00	0.00	0.00	0.00	0.00	0.00	0.00	0.00	0.00	0.00	0.00
3/22/2006 14:40	1	22-Mar-06-3	0.07	0.02	0.61	0.00	0.00	0.00	0.00	0.06	0.32	2.4	6.3	0.09	1.5	0.00	0.00
3/22/2006 12:00	3	24-Mar-06-3	0.08	1.9	3.2	0.05	0.00	0.00	0.02	0.1	144	20	210	0.38	6.1	0.03	0.95
3/27/2006 11:00	6	27-Mar-06-3	0.06	3.5	3.2	0.15	0.01	0.00	0.08	0.16	67	49	412	0.30	4.7	0.05	1.7
3/29/2006 18:00	8	29-Mar-06-3	0.07	32	22	0.21	0.03	0.00	0.15	0.03	16	95	724	50	519	0.20	3.2
3/31/2006 14:00	10	31-Mar-06-3	0.07	62	34	0.24	0.04	0.00	0.49	0.00	193	147	69	710	0.12	1.8	0.00
4/3/2006 16:00	13	03-Apr-06-3	0.07	20	57	0.26	0.06	0.00	1.4	0.00	165	225	73	755	0.11	1.8	0.00
4/6/2006 10:00	16	06-Apr-06-3	0.08	185	82	0.29	0.08	0.00	2.2	0.00	268	206	92	948	0.00	0.07	2.5
4/10/2006 14:00	20	10-Apr-06-3	0.07	333	118	0.29	0.11	0.00	3.5	0.00	393	291	68	700	0.00	0.07	2.5
4/12/2006 16:00	22	12-Apr-06-3	0.07	422	132	0.29	0.12	0.00	4.1	0.00	457	342	62	645	0.00	0.00	0.00
4/17/2006 9:00	27	17-Apr-06-3	0.08	688	156	0.29	0.13	0.00	5.5	0.00	577	430	32	326	0.00	0.05	1.8
4/19/2006 16:00	29	19-Apr-06-3	0.08	825	163	0.29	0.14	0.00	6.1	0.00	673	480	26	265	0.00	0.04	1.4
4/27/2006 10:00	37	27-Apr-06-3	0.08	989	171	0.29	0.16	0.00	6.8	0.00	817	492	41	427	0.00	0.06	2.1
5/1/2006 9:00	36	1-May-06-3	0.05	1179	183	0.29	0.18	0.00	7.6	0.00	731	552	36	373	0.00	0.07	2.5
5/5/2006 11:00	40	5-May-06-3	0.08	1456	191	0.29	0.20	0.00	8.7	0.00	729	550	6.9	72	0.00	0.05	1.8
5/8/2006 11:30	43	8-May-06-3	0.07	1674	193	0.31	0.21	0.00	10.0	0.00	808	611	0.9	1.4	0.04	1.4	0.00
5/10/2006 9:00	45	10-May-06-3	0.07	1812	193	0.32	0.22	0.00	10.0	0.00	785	597	2.0	2.1	0.00	0.04	1.4
5/12/2006 14:30	47	12-May-06-3	0.08	1972	194	0.32	0.23	0.00	11.0	0.00	661	427	3.4	3.5	0.00	0.04	1.4
5/14/2006 14:30	49	14-May-06-3	0.05	2082	184	0.32	0.24	0.00	11.0	0.00	689	524	3.5	3.8	0.00	0.05	1.8
5/15/2006 14:30	50	15-May-06-3	0.06	2134	184	0.32	0.24	0.00	11.0	0.00	703	509	3.7	3.9	0.00	0.05	1.6
5/16/2006 15:00	51	16-May-06-3	0.06	2192	195	0.32	0.25	0.00	12.0	0.00	623	470	6.1	6.3	0.00	0.05	1.8
5/17/2006 14:00	52	17-May-06-3	0.07	2255	196	0.32	0.25	0.00	12.0	0.00	603	610	7.8	80	0.00	0.04	1.4
5/18/2006 14:00	53	18-May-06-3	0.11	2355	206	0.32	0.25	0.00	12.0	0.00	796	656	8.0	61	0.00	0.04	1.4
5/18/2006 13:00	54	18-May-06-3	0.13	2495	197	0.32	0.26	0.00	13.0	0.00	669	611	5.6	5.8	0.00	0.04	1.4
5/24/2006 11:00	56	24-May-06-3	0.07	2728	199	0.32	0.27	0.00	13.0	0.00	767	590	4.2	43	0.00	0.03	1.1
5/25/2006 11:00	57	25-May-06-3	0.08	2908	200	0.32	0.29	0.00	14.0	0.00	720	547	1.3	13	0.00	0.04	1.4
5/26/2006 13:30	58	26-May-06-3	0.08	2984	200	0.32	0.29	0.00	14.0	0.00	640	487	6.7	6.9	0.00	0.13	4.6
5/29/2006 12:00	61	29-May-06-3	0.08	3135	202	0.32	0.31	0.00	15.0	0.00	835	631	7.3	7.6	0.00	0.03	1.1
5/31/2006 14:00	64	31-May-06-3	0.08	3309	223	0.32	0.36	0.00	16.0	0.00	723	544	6.6	6.8	0.00	0.00	0.00
6/2/2006 11:30	65	02-June-06-3	0.09	3446	205	0.32	0.33	0.00	16.0	0.00	828	630	13	131	0.00	0.04	1.4
6/5/2006 13:15	68	05-June-06-3	0.08	3786	208	0.32	0.34	0.00	18.0	0.00	845	643	5.2	53	0.00	0.00	0.03
6/7/2006 11:30	70	07-June-06-3	0.07	4351	219	0.32	0.36	0.00	21.0	0.00	904	683	10	101	0.00	0.00	0.00
6/8/2006 14:00	72	08-June-06-3	0.08	4168	214	0.32	0.35	0.00	19.0	0.00	844	642	14	144	0.00	0.00	0.03
6/12/2006 10:30	75	12-June-06-3	0.08	4427	217	0.36	0.36	0.00	20.0	0.00	780	593	6.9	71	0.00	0.00	0.00
6/15/2006 14:00	82	15-June-06-3	0.08	4745	219	0.36	0.36	0.00	22.0	0.00	710	547	15	176	0.00	0.00	0.00
6/16/2006 14:45	79	16-June-06-3	0.08	4863	220	0.36	0.36	0.00	21.0	0.00	991	750	4.5	4.6	0.00	0.00	0.00
6/16/2006 14:00	82	16-June-06-3	0.09	5191	221	0.36	0.36	0.00	22.0	0.00	931	703	4.4	4.6	0.00	0.00	0.00
6/21/2006 11:00	84	21-June-06-3	0.08	5309	223	0.36	0.36	0.00	23.0	0.00	723	544	6.6	6.8	0.00	0.00	0.00
6/23/2006 15:00	86	23-June-06-3	0.08	5620	224	0.36	0.36	0.00	24.0	0.00	955	721	4.1	4.3	0.00	0.00	0.00
6/26/2006 14:00	89	26-June-06-3	0.09	5996	226	0.36	0.36	0.00	25.0	0.00	1089	812	5.6	5.8	0.00	0.00	0.00
6/28/2006 14:00	92	28-June-06-3	0.08	6291	229	0.36	0.36	0.00	25.0	0.00	904	683	10	101	0.00	0.00	0.00
7/4/2006 13:30	97	04-July-06-3	0.08	6851	238	0.36	0.36	0.00	27.0	0.00	862	650	23	233	0.00	0.00	0.00
7/6/2006 13:30	99	06-July-06-3	0.09	7071	241	0.36	0.36	0.00	28.0	0.00	842	712	31	317	0.00	0.00	0.00
7/16/2006 15:40	103	16-July-06-3	0.07	7520	263	0.36	0.36	0.00	28.0	0.00	665	521	39	398	0.19	3.0	0.00
7/17/2006 13:00	106	17-July-06-3	0.08	7800	266	0.36	0.36	0.00	28.0	0.00	802	603	3.5	4.4	0.00	0.00	0.00
7/18/2006 13:30	111	18-July-06-3	0.08	8293	319	0.36	0.36	0.00	32.0	0.00	806	619	69	712	0.00	0.00	0.00
7/24/2006 14:45	117	24-July-06-3	0.08	8651	376	0.37	0.37	0.00	36.0	0.00	729	554	69	914	0.04	0.04	1.4
7/31/2006 16:15	124	31-July-06-3	0.08	9491	411	0.38	0.38	0.00	40.0	0.00	769	574	109	112	0.00	0.00	0.00
8/4/2006 15:15	132	8-August-06-3	0.08	10217	474	0.42	0.42	0.00	44.0	0.00	741	567	125	1286	0.00	0.00	0.00
8/14/2006 13:45	138	14-August-06-3	0.09	10784	666	0.41	0.48	0.00	48.0	0.00	767	560	128	1324	0.00	0.07	2.5
8/21/2006 14:45	145	21-August-06-3	0.08	11413	779	0.41	0.48	0.00	48.0	0.00	718	540	131	1348	0.00	0.11	3.9
8/26/2006 15:30	152	26-August-06-3	0.09	12082	897	0.41	0.47	0.00	45.0	0.00	752	572	129	1334	0.00	0.07	2.5
8/29/2006 14:15	160	5-Sept-06-1	0.08	12820	1031	2.3	0.8	0.00	48.0	0.00	718	543	137	1413	0.52	8.3	0.19
8/31/2006 15:15	166	11-Sept-06-1	0.08	13392	1133	1.0	1.0	0.00	48.0	0.00	718	541	146	1542	0.95	15.2	0.28
9/18/2006 10:30	173	18-Sept-06-3	0.08	13945	1284	3.3	1.2	0.00	54.0	0.00	690	524	157	1618	0.82	13.0	0.28
9/25/2006 13:00	180	25-Sept-06-3	0.08	14433	1387	4.5	1.4	0.00	57.0	0.00	472	355	135	1398	2.0	32.0	9.1
9/28/2006 14:00	201	28-Sept-06-3	0.08	14691	1459	5.3	1.6	0.00	59.0	0.00	461	351	148	1468	2.0	32.0	9.1
10/2/2006 11:00	187	02-Oct-06-3	0.09	14903	1516	5.6	1.7	0.00	60.0	0.00	546	414	148	1489	0.96	15.0	0.00
10/10/2006 14:00	195	10-Oct-06-3	0.08	15466	1666	6.5	1.9	0.00	64.0	0.00	574	437	154	1593	0.74	12.0	0.26
10/16/2006 13:30	201	16-Oct-06-3	0.08	16097	1782	7.1	2.1	0.00	66.0	0.00	541	416	162	1627	0.78	12.0	0.26
10/23/2006 10:15	208	23-Oct-06-3	0.09	16418	1932	8.0	2.4	0.00	72.0	0.00	582	442	179	1846	0.80	13.0	0.00
10/30/2006 10:15	215	30-Oct-06-3	0.08	16926	2109	8.4	2.6	0.00	77.0	0.00	581	443	227	2344	0.89	6.2	0.34
11/6/2006 15:00	223	06-Nov-06-3															

Anion Data for SRS™ Column

Date	Day	Sample ID	Flow Rate (ml/min)	Chloride (mg/L)	Chloride (µM)	Sum Chloride (µM)	Nitrite-N (mg/L)	Nitrate-N (mg/L)	Sulfate (mg/L)	Sulfate (µM)	Sum Sulfate (µM)	Bromide (mg/L)	Phosphate (mg/L)
3/21/2006 15:15	0												
3/22/2006 14:40	1	22-Mar-06-3	0.07	20	563	53	0.00	0.00	684	7203	679	62	0.00
3/24/2006 12:00	3	24-Mar-06-3	0.08	30	843	202	0.00	0.00	579	6096	2084	45	0.00
3/27/2006 11:00	6	27-Mar-06-3	0.06	32	900	429	0.00	0.00	--	--	3676	14	0.00
3/29/2006 16:00	8	29-Mar-06-3	0.07	48	1346	697	0.00	0.00	831	8745	5761	10	0.00
3/31/2006 14:00	10	31-Mar-06-3	0.07	36	1017	938	0.00	0.00	745	7841	7450	1.0	0.00
4/3/2006 16:00	13	03-Apr-06-3	0.07	41	1164	1293	0.00	0.00	753	7929	10017	1.5	0.00
4/6/2006 10:00	16	06-Apr-06-3	0.08	48	1352	1685	0.00	0.00	831	8752	12618	2.3	0.00
4/10/2006 14:00	20	10-Apr-06-3	0.07	39	1093	2197	0.00	0.00	796	8382	16208	0.00	0.00
4/12/2006 16:00	22	12-Apr-06-3	0.07	36	1025	2422	0.00	0.00	829	8730	18025	0.00	0.00
4/17/2006 9:00	27	17-Apr-06-3	0.08	--	--	2983	0.00	0.00	--	--	22797	0.00	0.00
4/19/2006 16:00	29	19-Apr-06-3	0.06	25	706	3133	0.00	0.00	904	9520	24827	0.00	0.00
4/27/2006 10:00	32	27-Apr-06-3	0.06	20	563	3276	0.00	0.00	--	--	26975	0.00	0.00
5/1/2006 9:00	36	1-May-06-3	0.05	27	772	3476	0.00	0.00	999	10421	30988	1.8	0.00
5/3/2006 11:30	38	5-May-06-3	0.08	25	699	3809	0.00	0.00	1001	10535	34835	0.00	0.00
5/8/2006 11:30	43	8-May-06-3	0.07	--	--	4031	0.00	0.00	--	--	38177	0.00	0.00
5/10/2006 9:00	45	10-May-06-3	0.07	23	638	4157	0.00	0.00	1015	10684	40283	0.00	0.00
5/12/2006 14:30	47	12-May-06-3	0.08	24	679	4317	0.00	0.00	1015	10688	42883	0.00	0.00
5/14/2006 14:30	49	14-May-06-3	0.05	27	765	4412	0.00	0.00	1049	11039	44314	0.00	0.00
5/15/2006 14:30	50	15-May-06-3	0.06	29	829	4483	0.00	0.00	1050	11053	45300	1.6	0.00
5/16/2006 15:00	51	16-May-06-3	0.06	27	924	4511	0.00	0.00	1029	10830	46289	1.4	0.00
5/17/2006 14:00	52	17-May-06-3	0.07	23	639	4623	0.00	0.00	1025	10790	47254	36	0.00
5/18/2006 14:00	53	18-May-06-3	0.11	23	637	4723	0.00	0.00	1020	10736	48944	57	0.00
5/19/2006 13:00	54	19-May-06-3	0.13	20	562	4835	0.00	0.00	1045	11003	50964	84	0.00
5/24/2006 11:00	56	24-May-06-3	0.07	23	646	4954	0.00	0.00	1029	10837	53119	94	0.00
5/25/2006 11:00	57	25-May-06-3	0.08	20	577	5023	0.00	0.00	954	10044	54302	92	0.00
5/26/2006 13:30	58	26-May-06-3	0.08	19	539	5094	0.00	0.00	957	10074	72424	0.00	0.00
5/29/2006 12:00	61	29-May-06-3	0.08	18	516	5261	0.00	0.00	973	10246	58804	75	0.00
5/31/2006 14:00	63	31-May-06-3	0.08	17	474	5383	0.00	0.00	959	10090	61307	24	0.00
6/2/2006 11:30	65	02-June-06-3	0.09	23	654	5523	0.00	0.00	1030	10847	63907	6.4	0.00
6/5/2006 13:15	68	05-June-06-3	0.08	20	577	5745	0.00	0.00	1029	10828	67817	2.6	0.00
6/7/2006 11:30	70	07-June-06-3	0.08	18	502	5860	0.00	0.00	999	10092	70048	0.00	0.00
6/9/2006 14:00	72	09-June-06-3	0.08	19	525	5981	0.00	0.00	957	10074	72424	0.00	0.00
6/12/2006 10:30	75	12-June-06-3	0.08	19	522	6148	0.00	0.00	965	10154	75649	0.00	0.00
6/15/2006 14:20	78	15-June-06-3	0.09	18	521	6356	0.00	0.00	966	10173	79693	0.00	0.00
6/16/2006 14:45	79	16-June-06-3	0.08	20	555	6420	0.00	0.00	992	10443	80920	0.00	0.00
6/19/2006 10:40	82	19-June-06-3	0.09	30	855	6668	0.00	0.00	982	10340	84586	0.00	0.00
6/21/2006 11:00	84	21-June-06-3	0.09	32	909	6890	0.00	0.00	958	10088	87159	0.00	0.00
6/23/2006 15:00	86	23-June-06-3	0.08	42	1176	7158	0.00	0.00	964	10164	89794	0.00	0.00
6/25/2006 14:00	89	25-June-06-3	0.09	43	1345	7624	0.00	0.00	954	10045	93200	0.00	0.00
6/29/2006 14:00	92	29-June-06-3	0.07	48	1340	8032	0.00	0.00	866	9112	96894	0.00	0.00
7/4/2006 13:30	97	04-July-06-3	0.08	64	1790	8970	0.00	0.00	826	8699	102234	0.00	0.00
7/6/2006 13:30	99	06-July-06-3	0.09	65	1829	9418	0.00	0.00	793	8346	104341	0.00	0.00
7/10/2006 13:40	103	10-July-06-3	0.11	72	2033	10592	0.00	0.00	801	8437	109443	0.00	0.00
7/13/2006 9:00	106	13-July-06-3	0.08	69	1956	11257	0.00	0.00	796	8273	112230	0.00	0.00
7/18/2006 13:30	111	18-July-06-3	0.08	67	1931	12439	0.00	0.00	726	7638	117126	0.00	0.00
7/24/2006 14:45	117	24-July-06-3	0.08	57	1614	13729	0.00	0.00	642	6759	122434	0.00	0.00
7/31/2006 16:15	124	31-July-06-3	0.08	56	1569	15101	0.00	0.00	548	5768	127832	0.00	0.00
8/8/2006 15:15	132	8-August-06-3	0.08	54	1524	16591	0.00	0.00	419	4413	132739	0.00	0.00
8/14/2006 13:45	138	14-August-06-3	0.09	51	1438	17678	0.00	0.00	272	2868	135411	0.00	0.00
8/21/2006 14:45	145	21-August-06-3	0.09	49	1379	18913	0.00	0.00	175	1846	137477	0.00	0.00
8/28/2006 15:30	152	28-August-06-3	0.09	69	1952	20485	0.00	0.00	119	1254	138941	0.00	0.00
9/5/2006 14:15	160	5-Sept-06-1	0.08	59	1666	22189	0.00	0.00	82	866	139939	0.00	0.00
9/11/2006 10:15	166	11-Sept-06-3	0.09	56	1573	23395	0.00	0.00	50	529	140459	0.00	0.00
9/18/2006 10:30	173	18-Sept-06-3	0.08	57	1602	24732	0.00	0.00	31	326	140819	0.00	0.00
9/25/2006 13:00	180	25-Sept-06-3	0.08	57	1608	26056	0.00	0.00	196	2067	141806	0.00	0.00
9/29/2006 14:00	184	29-Sept-06-3	0.09	62	1743	26960	0.00	0.00	200	2101	142931	0.00	0.00
10/2/2006 11:30	187	02-Oct-06-3	0.09	57	1594	27611	0.00	0.00	224	2361	143801	0.00	0.00
10/10/2006 11:40	195	10-Oct-06-3	0.08	60	1694	29126	0.00	0.00	242	2548	146064	0.00	0.00
10/16/2006 13:30	201	16-Oct-06-3	0.08	86	2422	30608	0.00	0.00	199	2095	147735	0.00	0.00
10/30/2006 10:15	215	30-Oct-06-3	0.08	113	3193	35240	0.00	0.00	50	525	149896	0.00	0.00
11/6/2006 15:00	222	06-Nov-06-3	0.08	100	2827	37680	0.00	0.00	46	484	150305	0.00	0.00
11/13/2006 11:55	229	13-Nov-06-3	0.08	126	3545	40177	0.00	0.00	49	514	150696	8.1	0.00
11/20/2006 14:15	236	20-Nov-06-3	0.07	167	4714	43142	0.00	0.00	40	420	151031	5.6	0.00
11/27/2006 14:30	243	27-Nov-06-3	0.07	193	5445	46620	0.00	0.00	18	189	151539	6.2	0.00
12/4/2006 12:30	250	04-DEC-06-3	0.07	157	4413	50054	0.00	0.00	11	112	151344	6.6	0.00
12/11/2006 15:30	257	11-DEC-06-3	0.07	69	1941	52337	0.00	0.00	36	37	151998	3.6	0.00
12/18/2006 14:00	264	18-DEC-06-3	0.06	70	1965	53526	0.00	0.00	1.8	19	151415	0.00	0.00
12/27/2006 11:00	273	27-DEC-06-3	0.00	34	953	53526	0.00	0.00	10	106	151415	0.00	0.00
1/2/2007 11:30	279	2-JAN-07-3	0.06	20	550	53920	0.00	0.00	5.2	54	151457	0.00	0.00
1/8/2007 15:00	285	8-JAN-07-3	0.06	16	285	54203	0.00	0.00	4.5	87	151488	0.00	0.00
1/15/2007 14:00	292	15-JAN-07-3	0.07	16	441	54538	0.00	0.00	5.4	56	151530	0.00	0.00
1/22/2007 17:00	299	22-JAN-07-3	0.08	17	469	54916	0.00	0.00	6.4	67	151581	0.00	0.00
1/29/2007 11:45	306	29-JAN-07-3	0.08	24	673	55366	0.00	0.00	1.8	19	151615	1.6	0.00
2/5/2007 14:00	313	5-Feb-07-3	0.08	24	674	55941	0.00	0.00	11	120	151674	5.1	0.00
2/12/2007 16:00	320	12-Feb-07-3	0.08	12	335	56370	0.00	0.00	3.4	36	151740	15	0.00
2/19/2007 16:30	327	19-Feb-07-3	0.08	19	592	56987	0.00	0.00	3.5	37	151769	0.54	0.00
2/26/2007 12:00	334	26-Feb-07-3	0.07	20	560	57083	0.00	0.00	5.0	53	151801	0.00	0.00
3/5/2007 11:30	341	5-March-07-3	0.06	29	813	57531	0.00	0.00	4.6	49	151834	0.00	0.00
3/8/2007 15:30	344	8-March-07-3	0.07	26	746	57786	0.00	0.00	2.4	25	151846	0.00	0.00
3/12/2007 15:00	348	12-March-07-3	0.07	25	692	58075	0.00	0.00	6.5	68	151865	0.00	0.00
3/15/2007 15:00	351	15-March-07-3	0.08	23	661	58310	0.00	0.00	16	168	151906	0.00	0.00
3/19/2007 15:30	355	19-March-07-3	0.07	28	782	58602	0.00	0.00	38	405	152022	0.00	0.00
3/26/2007 15:00	362	26-March-07-3	0.06	33	938	59115	0.00	0.00	27	288	152228	0.00	0.00
3/29/2007 15:30	365	29-March-07-3	0.07	59	1659	59524	0.00	0.00	82	863	152410	0.00	0.00
4/2/2007 15:30	369	2-April-07-3	0.08	102	2886	60603	0.00	0.00	232	2442	153194	0.00	





## pH and ORP Data for Control Column

Date	Time (Days)	pH	ORP
7/13/2006	6	6.53	95
7/18/2006	11	6.55	87
7/24/2006	17	6.52	78
7/31/2006	24	6.75	58
8/8/2006	32	6.65	57
8/14/2006	38	6.62	70
8/21/2006	45	6.65	80
8/28/2006	52	6.75	112
9/5/2006	60	6.89	83
9/11/2006	66	6.98	37
9/18/2006	73	7.10	35
9/25/2006	80	7.04	50
9/29/2006	84	6.89	62
10/2/2006	87	7.08	64
10/10/2006	95	6.98	55
10/16/2006	101	7.17	60
10/23/2006	108	7.10	40
10/30/2006	115	7.37	34
11/6/2006	122	7.22	31
11/13/2006	129	6.92	45
11/20/2006	136	6.94	21
11/27/2006	143	6.91	18
12/4/2006	150	6.93	21
12/11/2006	157	6.84	26
12/18/2006	164	6.53	28
12/27/2006	173	6.87	-9
1/2/2007	179	6.96	-1
1/8/2007	185	6.82	176
1/15/2007	192	6.96	9
1/22/2007	199	6.69	-22
1/29/2007	206	6.95	0
2/5/2007	213	6.99	-29
2/12/2007	220	6.86	-24
2/19/2007	227	6.97	145
2/26/2007	234	7.07	125
3/5/2007	241	7.13	161
3/12/2007	248	7.12	157
3/19/2007	255	6.74	160
3/26/2007	262	6.72	151
4/2/2007	269	7.05	-14
4/9/2007	276	7.02	-19
4/16/2007	283	6.98	-5
4/23/2007	290	6.70	63
4/30/2007	297	6.75	156
5/7/2007	304	6.74	156
5/14/2007	311	6.80	183
5/22/2007	319	6.87	178
5/28/2007	325	6.78	163
6/4/2007	332	6.64	165
6/11/2007	339	6.75	168
6/18/2007	346	6.67	161
6/25/2007	353	6.76	94
7/2/2007	360	6.87	162
7/9/2007	367	6.96	172
7/16/2007	374	6.78	148
7/23/2007	381	6.89	163
7/30/2007	388	6.85	167
8/6/2007	395	6.79	177
8/13/2007	402	6.69	156
8/20/2007	409	6.61	175
8/27/2007	416	6.63	37
9/10/2007	430	6.59	11
9/17/2007	437	6.69	68
9/24/2007	444	6.73	80
10/1/2007	451	6.61	161
10/9/2007	459	6.62	78
10/15/2007	465	6.55	54
10/22/2007	472	6.54	100

## pH and ORP Data for nBA Column

Date	Time (Days)	pH	ORP
3/22/2006	1	6.99	--
3/24/2006	3	7.13	--
3/27/2006	6	6.87	--
3/29/2006	8	7.08	--
3/31/2006	10	7.05	--
4/6/2006	16	7.06	--
4/10/2006	20	7.03	--
4/19/2006	29	7.23	--
5/1/2006	36	6.88	--
5/10/2006	45	6.72	--
5/12/2006	47	6.88	--
5/15/2006	50	6.75	--
5/17/2006	52	6.86	--
5/18/2006	53	6.92	--
5/19/2006	54	7.03	--
5/24/2006	56	6.93	--
5/26/2006	58	6.84	--
5/29/2006	61	6.92	--
5/31/2006	63	6.70	--
6/2/2006	65	6.78	--
6/5/2006	68	6.58	--
6/7/2006	70	6.60	--
6/9/2006	72	6.37	--
6/12/2006	75	6.42	--
6/15/2006	78	6.32	--
6/16/2006	79	6.34	--
6/19/2006	82	6.47	--
6/21/2006	84	6.54	--
6/23/2006	86	6.47	--
6/26/2006	89	6.64	--
6/29/2006	92	6.57	--
7/4/2006	97	6.49	--
7/6/2006	99	6.67	3
7/10/2006	103	6.57	66
7/13/2006	106	6.70	71
7/18/2006	111	6.67	44
7/24/2006	117	6.63	35
7/31/2006	124	6.83	13
8/8/2006	132	6.73	15
8/14/2006	138	6.72	50
8/21/2006	145	6.66	72
8/28/2006	152	6.83	112
9/5/2006	160	6.98	72
9/11/2006	166	7.07	4
9/18/2006	173	7.07	11
9/25/2006	180	7.09	41
9/29/2006	184	6.92	57
10/2/2006	187	7.18	60
10/10/2006	195	7.01	42
10/16/2006	201	7.23	45
10/23/2006	208	7.12	22
10/30/2006	215	7.43	1
11/6/2006	222	7.15	4
11/13/2006	229	6.75	11
11/20/2006	236	6.65	-24
11/27/2006	243	6.67	-12
12/4/2006	250	6.56	-9
12/11/2006	257	6.51	10
12/18/2006	264	6.30	1
12/27/2006	273	6.75	-46
1/2/2007	279	6.80	-14
1/8/2007	285	6.88	--
1/15/2007	292	6.62	-11
1/22/2007	299	6.42	-27
1/29/2007	306	6.60	-27

## pH and ORP Data for SRS™ Column

Date	Time (Days)	pH	ORP
3/21/2007	0	7.37	--
3/22/2006	1	6.97	--
3/24/2006	3	7.01	--
3/27/2006	6	6.81	--
3/29/2006	8	7.02	--
3/31/2006	10	6.99	--
4/6/2006	16	6.99	--
4/10/2006	20	7.00	--
5/1/2006	36	6.98	--
5/8/2006	43	6.99	--
5/12/2006	47	6.93	--
5/15/2006	50	6.90	--
5/17/2006	52	6.98	--
5/18/2006	53	6.93	--
5/19/2006	54	7.01	--
5/24/2006	56	6.93	--
5/26/2006	58	7.04	--
5/29/2006	61	6.91	--
5/31/2006	63	6.85	--
6/2/2006	65	6.81	--
6/5/2006	68	6.73	--
6/7/2006	70	6.69	--
6/9/2006	72	6.51	--
6/12/2006	75	6.58	--
6/15/2006	78	6.52	--
6/16/2006	79	6.59	--
6/19/2006	82	6.60	--
6/21/2006	84	6.61	--
6/23/2006	86	6.56	--
6/26/2006	89	6.68	--
6/29/2006	92	6.72	--
7/4/2006	97	6.65	--
7/6/2006	99	6.72	-18
7/10/2006	103	6.60	27
7/13/2006	106	6.64	32
7/18/2006	111	6.65	50
7/24/2006	117	6.64	9
7/31/2006	124	6.68	10
8/8/2006	132	6.69	-7
8/14/2006	138	6.69	8
8/21/2006	145	6.59	-10
8/28/2006	152	6.66	1
9/5/2006	160	6.77	-12
9/11/2006	166	6.70	-24
9/18/2006	173	6.77	-17
9/25/2006	180	6.65	-29
9/29/2006	184	6.56	-34
10/10/2006	195	6.61	-31
10/16/2006	201	6.81	-23
10/23/2006	208	6.64	-22
10/30/2006	215	6.72	-26
11/6/2006	222	6.54	-40
11/13/2006	229	6.39	-32
11/20/2006	236	6.30	-30
11/27/2006	243	6.32	-25
12/4/2006	250	6.38	-53
12/11/2006	257	6.59	-147
12/18/2006	264	6.51	-183
12/27/2006	273	6.77	-160
1/2/2007	279	6.83	-162
1/8/2007	285	6.73	-164
1/15/2007	292	6.89	-186
1/22/2007	299	6.72	-172
1/29/2007	306	6.89	-179
2/5/2007	313	6.89	-177
2/12/2007	320	7.00	-118
2/19/2007	327	6.78	-202
2/26/2007	334	6.95	-172
3/5/2007	341	6.97	-203
3/12/2007	348	6.93	-211
3/19/2007	355	6.90	-163
3/26/2007	362	6.87	-166
4/2/2007	369	6.97	-151
4/9/2007	376	6.89	-139
4/16/2007	383	6.84	-123
4/23/2007	390	6.92	-123
4/30/2007	397	6.97	-141
5/7/2007	404	7.04	-80
5/14/2007	411	7.06	-88
5/22/2007	419	7.01	-80
5/28/2007	425	7.08	-87
6/4/2007	432	7.04	-281
6/11/2007	439	7.15	-230
6/18/2007	446	6.97	-220
6/25/2007	453	7.11	-224
7/2/2007	460	7.17	-266
7/9/2007	467	7.28	-262
7/16/2007	474	7.10	-232
7/23/2007	481	7.03	-251
7/30/2007	488	7.00	-273
8/6/2007	495	7.03	-235
8/13/2007	502	7.00	-266
8/20/2007	509	6.92	-196
8/27/2007	516	6.95	-194
9/10/2007	530	6.99	-110
9/17/2007	537	6.96	-101
9/24/2007	544	6.99	-118
10/1/2007	551	6.79	161
10/9/2007	559	6.96	71
10/15/2007	565	6.89	43
10/22/2007	572	6.81	-9

## Microbial Data for All Columns

Control Column						
Sample ID	Sample Date	DAY	<i>Dehalococcoides</i> 16S Gene Copies/L	QL	<i>Dehalococcoides</i> VC Reductase Gene Copies/L	QL
Control Column	October 2 to October 12 2006	97	Not detected	1 x 10 <sup>4</sup> /liter	--	--
Control Column	January 22 to February 2 2007	210	Not detected	4 x 10 <sup>2</sup> /liter	--	--
Control Column	April 19 to May 2 2007	300	6 x 10 <sup>3</sup> /liter <sup>(1)</sup>	2 x 10 <sup>4</sup> /liter	2 x 10 <sup>4</sup> /liter	2 x 10 <sup>4</sup> /liter
Control Column	July 27 to August 13 2007	402	2 x 10 <sup>4</sup> /liter <sup>(1)</sup>	2 x 10 <sup>4</sup> /liter	--	--
Control Column	Sept 24 07 - Oct 12 07	472	Not detected	4 x 10 <sup>3</sup> /liter	--	--

nBA Column						
Sample ID	Sample Date	DAY	<i>Dehalococcoides</i> 16S Gene Copies/L	QL	<i>Dehalococcoides</i> VC Reductase Gene Copies/L	QL
nBA Column	May 10, 2006	45	1 x 10 <sup>6</sup>	1 x 10 <sup>5</sup> /liter	4 x 10 <sup>5</sup> /liter	1 x 10 <sup>5</sup> /liter
nBA Column	June 2 - June 12 2006	75	Not detected	1 x 10 <sup>4</sup> /liter	--	--
nBA Column	June 28 to July 7 2006	100	Inconclusive	9 x 10 <sup>3</sup> /liter	2 x 10 <sup>5</sup> /liter	9 x 10 <sup>3</sup> /liter
nBA Column	October 2 to October 12 2006	197	Inconclusive	1 x 10 <sup>4</sup> /liter	--	--
nBA Column	December 4 to December 15 2006	261	9 x 10 <sup>3</sup> /liter	9 x 10 <sup>3</sup> /liter	1 x 10 <sup>4</sup> /liter	9 x 10 <sup>3</sup> /liter
nBA Column	January 22 to January 29 2007	306	5 x 10 <sup>2</sup> /liter <sup>(1)</sup>	9 x 10 <sup>2</sup> /liter	7 x 10 <sup>3</sup> /liter	9 x 10 <sup>3</sup> /liter

SRS™ Column						
Sample ID	Sample Date	DAY	<i>Dehalococcoides</i> 16S Gene Copies/L	QL	<i>Dehalococcoides</i> VC Reductase Gene Copies/L	QL
SRS™ Column	May 10, 2006	45	Inconclusive	1 x 10 <sup>4</sup> /liter	2 x 10 <sup>5</sup> /liter	1 x 10 <sup>5</sup> /liter
SRS™ Column	June 2 - June 12	75	Not detected	1 x 10 <sup>4</sup> /liter	--	--
SRS™ Column	June 28 to July 7	100	Not detected	1 x 10 <sup>4</sup> /liter	--	--
SRS™ Column	October 2 to October 12	197	2 x 10 <sup>4</sup> /liter	1 x 10 <sup>4</sup> /liter	--	--
SRS™ Column	November 3 to November 14 2006	230	1 x 10 <sup>3</sup> /liter	4 x 10 <sup>3</sup> /liter	4 x 10 <sup>2</sup> /liter	4 x 10 <sup>3</sup> /liter
SRS™ Column	December 22 2006 to January 2 2007	279	1 x 10 <sup>8</sup> /liter	4 x 10 <sup>3</sup> /liter	1 x 10 <sup>6</sup> /liter	4 x 10 <sup>3</sup> /liter
SRS™ Column	February 2 to February 16 2007	324	1 x 10 <sup>8</sup> /liter	4 x 10 <sup>3</sup> /liter	8 x 10 <sup>7</sup> /liter	4 x 10 <sup>3</sup> /liter
SRS™ Column	March 19 to April 5 2007	372	2 x 10 <sup>7</sup> /liter	2 x 10 <sup>4</sup> /liter	7 x 10 <sup>7</sup> /liter	2 x 10 <sup>4</sup> /liter
SRS™ Column	April 19 to May 2 2007	400	1 x 10 <sup>7</sup> /liter	2 x 10 <sup>4</sup> /liter	4 x 10 <sup>7</sup> /liter	2 x 10 <sup>4</sup> /liter
SRS™ Column	June 15 to June 29 2007	453	2 x 10 <sup>7</sup> /liter	2 x 10 <sup>4</sup> /liter	6 x 10 <sup>7</sup> /liter	2 x 10 <sup>4</sup> /liter
SRS™ Column	July 27 to August 13 2007	502	8 x 10 <sup>7</sup> /liter	2 x 10 <sup>4</sup> /liter	1 x 10 <sup>8</sup> /liter	2 x 10 <sup>4</sup> /liter
SRS™ Column	October 12 to October 25 2007	572	6 x 10 <sup>8</sup> /liter	4 x 10 <sup>3</sup> /liter	7 x 10 <sup>8</sup> /liter	4 x 10 <sup>3</sup> /liter